

# **SOIL, PLANT AND WATER REFERENCE METHODS FOR THE WESTERN REGION<sup>1</sup>**

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## PREFACE

This manual has been developed as a listing of recommended analytical methods for the analysis of soil, plant and irrigation water samples by agricultural laboratories for use in the Western Region, by the Western Coordinating Committee on Nutrient Management (WERA-103). This publication is an update of *Plant, Soil and Water Reference Methods for the Western Region*, 1994, (WREP 125) written by Dr. Ray Gavlak formerly of the University of Alaska, Dr. Donald Horneck of Oregon State University, and Dr. Robert O. Miller of Colorado State University.

Methods outlined in the appendix are for use as additional resources for the user and WERA-103 does not support or condone emerging methods. We would like to thank all those individuals who have contributed to this manual. We encourage all suggestions and comments from participating laboratories for improving this manual for future publication.

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## QUALITY ASSURANCE IN THE AGRICULTURAL LABORATORY

A Quality Assurance (QA) program Quality Control (QC) is essential for demonstrating long term performance of accuracy and precision to the laboratory clientele. By developing and implementing a QC program that not only monitors the process but provides a mechanism for improvement, the agricultural laboratory can only enhance its credibility and that of the industry.

Quality Assurance is a set of operating principles when strictly followed in the analytical laboratory will produce analytical results of known and defensible quality. It is composed of two main sub groups quality control and quality assessment. Quality control (QC) consists of analytical appraisal tools which the laboratory utilizes to verify the analytical process. A few of these are: use of control samples, recovery of known additions, analysis of external standards, use of analytical duplicates and use of maintenance or control sample charts. The process of using QC measures in the analytical laboratory is quality assessment and includes those items such as performance sample evaluations and performance audits.

The first step towards the development a QC program involves the establishment of quality control measures. The purpose of these are to monitor the analytical process(s), document statistically the precision and accuracy, and establish limits of analytical control of the method in the laboratory. The following list describes several terms used in a QC control program:

Accuracy: A combination of bias and precision of an analytical procedure, which reflects the closeness of an individual measured value to a true, correct, or assumed value.

Bias: A consistent deviation of measured values from the true value, caused by systematic errors in a procedure. Bias is assessed by measuring the recovery of known additions (spiked samples) and the recovery of internal standards and laboratory control or reference standards.

Detection Limits: The common term that encompasses various analytical detection limits. Some of the common detection limits (in increasing order of concentration detected) include the instrument detection limit (IDL), the lower limit of detection (LLD), the method detection limit (MDL), and Practical Quantitation Limit (PQL). The MDL, is the minimum concentration of a analyte that can be measured and reported with 99% confidence that has gone through the entire analytical process, including sample preparation and instrument analysis. MDL is a more useful indicator of the reported detection limit and is always larger than the IDL or LLD because it includes recovery efficiencies and concentration factors in the sample preparation. The MDL is the lowest level that can be achieved by an experienced analyst averaging at least seven trials and operating a well-calibrated instrument on one day. MDL is often estimated as three to five times the standard deviation (99.6%-99.9% confidence level) of the sample preparation blank concentration. The relation among these limits is about IDL:LLD:MDL:PQL = 1:2:4:10 . Most of the studies report one or more of the detection limits. PQL is the lowest level that can be quantified (measured) accurately (within  $\pm 10\%$ ) and reliable day-in and day-out in the lab. The PQL is routinely reported at 10 times the MDL which assures that any reported value is reliable. In certain instances it is referred as Limit of Quantitation (LOQ) or Reported Detection Limit (RDL).

Instrument Calibration Standards: Those standards prepared for the expressed purpose of calibration of an instrument for the determination of an known analyte in a unknown sample.

Laboratory Reference Standards: -A standard, usually certified, by an outside agency, used to measure the bias and precision in a procedure. Examples include the National Institute of Standards and Technology (NIST) Standard Reference Materials and the National Research Council of Canada (NRCC) reference materials (See Appendix C). Semiquantitative analytes cannot be certified.

Matrix Duplicate: An intra-laboratory split sample which is used to document the precision of a method in a given sample matrix. Duplicate analysis can also be used to detect calibration errors or "drift" and to detect sample to sample analyte carry-over contamination.

Method Blank: an analyte-free matrix to which all reagents are added in the same volumes or proportions as used in sample processing (labware, filter paper, reagents, instrumentation etc.). The method blank should be carried through the complete sample preparation and analytical procedure and is used to document contamination resulting from the analytical process. Blanks analyzed before, after, or randomly during a sample run can detect carry-over contamination.

Precision: a measure of the degree of agreement among replicate analyses of a sample (e.g., standard deviation, percent difference, or percent relative standard deviation).

Quality Control Reference Sample: a matrix matched internal control standard routinely used in the laboratory to evaluate long term accuracy and precision. An internal control sample should be randomly placed amongst each batch of unknown samples representing a minimum of 5% of the total samples.

Random Error: the deviation of any analytical value that can't be ascertained by standard statistical techniques.

Replicate: - a repeated operation occurring within an analytical procedure (sample, extraction etc.). Two or more analyses for the same constituent in an extract of a single sample constitute replicate extract analyses.

Systematic Error: is the difference between the value obtained for a characteristic and the true or conventional value which cannot be attributed to random error.

Surrogate Standard Addition: a pure element/compound added to a sample prior to analysis at varying levels for the purpose of evaluating overall efficiency of a method.

Uncertainty: all analyses have some level of inexactness in the reported value, reports may have an "estimated uncertainty" associated with reported measurements. Estimated uncertainty should be expressed in relative or absolute terms, and not as a range. In addition, reported uncertainties should reflect all knowledge of the measurement that might add to the value reported, (e.g., matrix interferences, homogeneity, blanks, etc).

The following outlines key components to the operational structure of a QA program:

- Staff organization and responsibilities
- Sample control and documentation procedures
- Standard operating procedures for each analytical method
- Analyst training requirements
- Equipment maintenance procedures
- Instrument calibration procedures
- Internal quality control activities
- Performance testing audits
- Data assessment procedures
- Validation and Reporting

From the operational plan both internal and external QC components are defined. Specific elements of an internal QC program: operator proficiency certification; laboratory calibration standard checks; analysis of reagent blanks; recovery of known standard additions, analysis of external supplied standards; analysis of replicates and control charts. Operator proficiency identifies the competence of the analyst performing the method. Calibration checks verify the instrument performance and quality of the instrument calibration standards. Analysis of reagent blanks document background contamination levels and the LOD and MDL. The use of recovery of known standards and analysis of external standards verify performance of the method for the analyte of interest, while duplicates provide data on precision. External QC components involve the use of external evaluation samples of known concentration through proficiency testing programs and purchased reference standards.

This brief overview is to familiarize the laboratory analyst/manager with the major components of a QA program. For the reader who is interested in developing or upgrading a QA program in the agricultural laboratory a number of articles and reference sources are available. The following is a brief list.

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## Scope and Application

This method quantifies the soil water content of a saturated soil. At saturation all soil pore space is occupied by water and no free water collects on the surface. Salinity crop tolerance data; the relationships between cation solution concentrations and soil exchangeable cations (i.e. SAR); and soluble soil boron, are based on the saturation paste extract (U.S. Salinity Laboratory Staff, 1954 and Robbins, 1990). From the saturation paste, soil pH may be determined directly on the paste (Method S - 1.10). By extracting the liquid phase of the saturation paste under partial vacuum estimates of: electrical conductivity,  $EC_e$  (soluble salts); solution concentrations of  $Na^+$ ,  $K^+$ ,  $Ca^{2+}$ ,  $Mg^{2+}$ ,  $Cl^-$ ,  $HBO_3^-$ ,  $NO_3^-$ ,  $SO_4^{2-}$ ,  $Mn^{2+}$ ,  $SeO_4^{2-}$ ,  $HCO_3^-$ ,  $CO_3^{2-}$ ; and SAR can be determined. Estimates of soil water holding capacity, wilting point and texture can be made from the saturated moisture content. The method is generally reproducible within  $\pm 12\%$ , dependent on the soil textural class (Klages, 1984).

## Equipment

1. Analytical balance: 500.0 g capacity, resolution  $\pm 0.1$  g
2. 500 mL container and cap (polypropylene container or 16 oz waxed paper cups).
3. Spatula, Blade 17.5 mm x 100 mm length.
4. Buchner filter assembly (preferably plastic) and vacuum system (capable of - 90 kPa).
5. Whatman No. 5 filter paper, or equivalent highly retentive filter paper.
6. Test tube or vial, 50 mL, polypropylene with cap.

## Reagents

1. Deionized water, ASTM Type I grade.  $EC < 10^{-4}$  dS  $m^{-1}$

## Procedure

1. Weigh  $200.0 \pm 0.5$  g air-dry soil pulverized to pass 10 mesh sieve ( $< 2.0$  mm) of known water content ( $P_w$ , %), into a 500 mL container and record total weight (See Comments #1 and #2).
2. Gradually add deionized water and mix uniformly (free of partially wetted clumps) until a saturated paste is obtained (See Comments #3 and #4). At saturation, the soil paste:
  - i. Does not have free standing water on the surface of the paste.
  - ii. Soil paste slides freely and cleanly off a spatula (does not apply to high clay soils,  $> 40\%$  clay).
  - iii. Paste will flow slightly when the container is tipped to a 45 degree angle from horizontal.
  - iv. Soil surface glistens as it reflects light.
  - v. Consolidates easily by tapping after a trench is formed in the paste with the flat side of a spatula (may not apply to sandy soils  $> 70\%$  sand).
3. Record weight, cap container and let stand for four (4) hours. Check saturation characteristics again and add soil or water as needed to obtain the desired characteristics (See Comment #5).
4. Record the mass of the soil (g) and total water (g) added.
5. After equilibration, thoroughly remix samples and determine soil pH, Method S - 1.10 (See Comment #6).
6. Transfer soil saturation paste to buchner funnel filter paper and spread evenly over surface. Apply -80 kPa vacuum and collect filtrate in test tube. Discontinue vacuum when cracks appear in soil paste. Refilter if filtrate is turbid. Determine  $EC_e$ ,  $HCO_3^-$  and  $CO_3^{2-}$  with in five (5) minutes (Methods S - 1.20 and S - 1.30). Cap and retain filtrate for additional analysis (See Comments #7, #8 and #9).

## Calculations

$$\text{SP \%} = \frac{(\text{Amount of water (g), added}) \times 100}{(\text{mass of air dry soil (g)} \times ((100 - P_w)/100))} \quad [\text{equ. S-1.0-1}]$$

Report saturation percentage (SP) to the nearest 0.1% (See Comment #10 and #11).

## Comments

1. Soil samples should not be oven-dried above 70 °C prior to extracting for soluble salts.
2. For organic soils (>16% organic matter) it is advisable to start with a 150 mL of water and add soil material.
3. Fine textured soils (> 40% clay) may puddle easily. To minimize puddling and obtain a more definite endpoint with fine-textured soils, water should be added with a minimum amount of stirring, especially in the early stages of wetting. Peat soils (>16% organic matter) will require soaking for twenty-four (24) hours. The method can be used assess greenhouse potting media.
4. Some fine textured soils swell considerably upon addition of water. In these cases, steps 2 and 3 must be repeated until the paste characteristics are stable. For salinity appraisal the paste can be extracted after four (4) hours; however, for sodic soil samples it should stand sixteen (16) or more hours. For the assessment of soil soluble boron, twenty-four (24) hours of paste equilibration is required.
5. Coarse textured soils, sandy loam and loamy sand with less than 15% clay, may not exhibit saturated paste characteristics of finer textured soils. For these soil types the relative accuracy of the method declines and should be noted when making soil comparisons.
6. If calcium carbonate precipitates are noted in the extract, dilute paste extract 1:1 with deionized water and note dilution in subsequent analysis. Samples may be refrigerated (4 °C) for storage (do not allow to freeze) for 30 days. Small quantities (200 uL) of thymol or toluene may be added to minimize the influence of microbial activity while samples are refrigerated (Carlson et al., 1971).
7. Determining saturated paste percentage alternative: take a 30 - 50 g sub sample of the paste, weigh, oven dry at 105 °C for four (4) hours, reweigh and calculate saturation percentage. Oven dry moisture values will be slightly higher than the direct method as air dry soil will retain 3-5% moisture, dependent on clay and salt content.
8. Extraction consistency is best achieved using a vacuum of -60 to -80 KPa (-0.6 to -0.8 bars) applied for thirty (30) minutes (Jacobson and Sandoval, 1970). Soils maybe centrifuged.
9. Approximately one-quarter to one-third of the water added in making the saturated paste can be recovered as extract (Loveday, 1974).
10. Soil Field Capacity (FC, 33 kPa) can be estimated from the saturation percentage as follows:  $\text{SP} \times 0.5 \approx \text{FC}$ . Soil water potential Permanent Wilting Point (PWP, 1500 kPa) can be estimated as follows:  $\text{SP} \times 0.25 \approx \text{PWP}$ . Saturation percentage is related to soil texture see Table S-1.00-A (based on organic matter contents less than 3%).
11. For fine-textured soils and those high in sodium ( $\text{SAR} > 10$ ), SP cannot be used to estimate FC and PWP values (Reeves et al. 1954).

Table S-1.00-A. Saturated paste moisture as it relates to soil texture. (Robbins, 1990).

SP (%)	Soil Texture
< 20	sand or loamy sand
20 - 35	sandy loam
35 - 50	loam or silt loam
50 - 65	clay loam
65 -	clay
> 81	organic soils

## Literature

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## Scope and Application

This method semi-quantifies the soil pH of a soil saturated paste (Method S - 1.00). Soil pH is a measure of the relative acidity or alkalinity of the soil solution that is in equilibrium with the solid particles. It is a measure of the intensity of acidity or alkalinity, but does not indicate the relative buffering capacity of the soil. It is most applicable to salt-affected soils with a pH ranging from 6.0 to 9.0 (Robbins et. al. 1990). Soil pH is measured to access soil chemical properties, crop suitability, lime needs and relative nutrient availability. The method is generally reproducible within  $\pm 0.10$  pH units.

## Equipment

1. pH meter, equipped with pH electrodes (indicating and reference).
2. Primary standard buffers, pH 4.00, 7.00, and 10.0.

## Procedure

1. Prepare a saturation paste, as outlined in Method S - 1.00.
2. Standardize / Calibrate the pH meter: (1) rinse electrode with deionized water and place in pH 7.00 primary standard buffer and adjust as necessary; (2) rinse electrode and place in pH 4.00 primary standard buffer; (3) adjust the slope until response is  $\pm 0.05$  units of expected response; and (4) check pH 7.00 primary standard buffer and adjust as necessary (See Comment #1). For high pH soils ( $>7.00$ ) use pH buffers 7.00 and 10.0.
3. Insert electrode into soil paste and gently rotate the container to remove entrapped air. When the meter has stabilized record soil pH as  $pH_{sp}$  to the nearest 0.01 pH unit.
4. Remove electrode(s), rinse with deionized water and blot excess water with filter paper (See Comment #2).

## Comments

1. Follow manufacturer's guidelines if meter does not read within 0.05 units of primary standards. Maintenance of combination electrodes differs from that of separate reference and glass electrodes; refer to manufacturer's instructions.
2. Store pH electrodes according to manufacturer's instructions (recommended practice is to store the electrodes in a primary standard buffer).

## Literature

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### Scope and Application

This method quantifies the amount of dissolved salts ( $\text{mg L}^{-1}$ ) by measurement of the electrical conductivity ( $\text{EC}_e$ ) of the soil saturated paste extract (Method S - 1.00). The relationship between  $\text{EC}_e$  and soluble salts is approximate due to differences in equivalent weights, ion equivalent conductivities, and relative proportions of major solutes in the paste extracts (Robbins, 1990). The  $\text{EC}_e$  measurement is sensitive to temperature and increases approximately 1.9% per °C (range 15 - 35 °C) (Rhoades, 1996). All  $\text{EC}_e$  data is normalized to 25 °C. Salt tolerance crop data is generally expressed in terms of the ( $\text{EC}_e$ ) of the saturation paste extract and used to assess the potential of soluble soil salts which may limit crop productivity. The method detection limit is approximately  $0.01 \text{ dS m}^{-1}$  ( $\text{mmhos cm}^{-1}$ ) and is generally reproducible within  $\pm 7\%$ .

### Equipment

1. Conductance meter with dynamic range from  $0.01$  to  $100 \text{ dS m}^{-1}$  conductance, temperature compensating, 25 °C.
2. Conductance cell having a cell constant (K) appropriate to the EC of the sample being measured (see Table S -1.2 -1). Pipet-type or dip-type cell and it recommended that it be capable of measuring temperature.

### Reagents

1. Deionized water  $\text{CO}_2$ -free, ASTM Type I grade.  $\text{EC} < 10^{-4} \text{ dS m}^{-1}$ .
2. Standard Reference Calibration Solution. Dissolve  $0.7456 \text{ g KCl}$  (previously dried at  $110 \text{ °C}$  for 2 h) in  $\text{CO}_2$ -free deionized water and dilute to  $1.0 \text{ L}$ . At  $25 \pm 0.1 \text{ °C}$  a  $0.010 \text{ N KCl}$  solution will have an  $\text{EC}_e$  of  $1.412 \text{ dS m}^{-1}$  ( $\text{mmhos cm}^{-1}$ ). For a  $0.100 \text{ N KCl}$  solution ( $7.456 \text{ g KCl}$  diluted to  $1.0 \text{ L}$ ) will have an  $\text{EC}_e$  of  $12.900 \text{ dS m}^{-1}$ . Standard EC calibration solutions are listed in Table S-1.20-A and can be purchased from a scientific supply vendor.

### Procedure

1. Prepare a saturation paste, as described in Method S - 1.00, and retain extract for  $\text{EC}_e$  measurement (See Comment #1).
2. Calibrate conductance cell. Operate and adjust instrument in accordance with manufacturer's instructions (See Comments #2 and #3). Rinse conductance cell with three aliquots of  $0.01 \text{ N KCl}$ , adjust a fourth portion to  $25 \pm 0.1 \text{ °C}$ , measure R (where R is the measured resistance ohms) and temperature  $t$ . Repeat measurement of R until value is constant. Calculate cell constant K. Develop four point calibration curve.

$$K = (0.001413) R_{\text{KCl}} / [1 + 0.019(25 - t)]$$

3. Rinse conductance cell with deionized water. Draw approximately  $2.0 \text{ mL}$  of soil saturation paste extract solution into conductance cell rinse and replace with a second aliquot. When the meter has stabilized record instrument reading.

### Calculations

$$\text{EC}_{25} = C_x(1000)K[1 + 0.019(25 - t)]$$

Where:  $C_x$  is the instrument measured value of the sample and  $t$  is temperature

Report  $EC_e$  to the nearest  $0.01 \text{ dS m}^{-1}$  as  $EC_e 25^\circ\text{C}$ .

(See Comments #4, and #5)

Table S -1.20-A Conductivity of KCl solutions at  $25^\circ\text{C}$  (Rhoades, 1996).

Concentration N	Conductivity dS m
0.001	0.147
0.010	1.413
0.020	2.767
0.050	6.668
0.10	12.90
0.20	24.82
0.50	58.64

### Comments

1. Exposure of the sample to the atmosphere may cause changes in conductivity due to loss or gain of dissolved gasses:  $\text{CO}_2$  and  $\text{NH}_3\text{-N}$ . Freshly distilled water has a conductivity of  $0.005 - 0.002 \text{ dS m}^{-1}$  increasing after a few weeks to  $0.002 - 0.004 \text{ dS m}^{-1}$ . This of special concern on samples with very low  $EC_e$ .
2. Clean platinum electrodes that are new or that are providing erratic EC readings with acid-dichromate cleaning solution. Cleaning solution: 32 mL of saturated sodium dichromate ( $\text{Na}_2\text{Cr}_2\text{O}_7$ ) and 1L 16 M sulfuric acid. Soak electrodes 16 hours followed by three rinses of deionized water rinses. If platinum is flaked, recoat according to procedure of APHA (1985).
3. For highly saline soils ( $EC_e > 8.0 \text{ dS m}^{-1}$ ) calibrate using  $0.100 \text{ N}$  KCl solution,  $EC_e 12.90 \text{ dS m}^{-1}$ .
4. The relationship between conductivity and soluble salts is approximate due to differences in solutes, solute conductivities, and equivalent weights. The general relationship (for solutions with an  $EC_e$  range of  $0.10 - 2.0 \text{ dS m}^{-1}$ ) is:

$$\text{Dissolved salt concentration (mg L}^{-1}\text{)} \approx 640 \times EC_e, \text{ in dS m}^{-1}$$

$$\text{Total cations (or anions) (mmolc L}^{-1}\text{ or meq L}^{-1}\text{)} \approx 10 \times EC_e, \text{ in dS m}^{-1}$$

$$\text{Osmotic potential at } 25^\circ\text{C (KPa)} \approx 0.39 \times EC_e, \text{ in dS m}^{-1}$$

The factor for converting  $EC_e$  to total dissolved salts ( $\text{mg L}^{-1}$ ) ranges from 550 to 900 dependent on the specific anions present and their concentration. For estimating approximate total cations or anions, USDA Handbook #60, Figure 4, graphically shows this relationship for typical salt concentrations.

5. Plant tolerances to salinity ( $EC_e$ ) of the soil saturated paste extract shown in Table S -1.2 - B.

Table S -1.20-B Impact of saturated paste soil salinity (EC<sub>e</sub>) on plant growth.

dS m <sup>-1</sup>	Plant salinity effects, productivity reduced 25%
0 - 2	salinity effects negligible (field bean, carrot, onion, red clover, strawberry)
2 - 4	very sensitive crops affected (spinach, lettuce, citrus, grape, alfalfa)
4 - 8	moderately salt tolerant crops affected (tomato, beet, wheat)
8 - 16	only salt tolerant crops yield satisfactory (barley, wheatgrass cotton, asparagus)
> 16	few salt tolerant crops yield satisfactory

## Literature

APHA (1985). Part 205. In *Standard Methods for the Examination of Water and Wastewater*. 16<sup>th</sup> edn. American Public Health Association, Washington, DC.

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Helrich, K. (ed.) 1990. *Official methods of analysis*. 15<sup>th</sup> Ed. Association of Official Analytical Chemists, Inc. Arlington, VA.

Rayment, G.E. and F.R. Higginson. 1992. Electrical conductivity. p. 15-16. *In: Australian Laboratory Handbook of Soil and Water Chemical Methods*. Inkata Press, Melbourne.

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## SATURATION PASTE EXTRACT ALKALINITY

### Bicarbonate and Carbonate

S - 1.30

### Scope and Application

This method quantifies bicarbonate ( $\text{HCO}_3^{1-}$ ) and carbonate ( $\text{CO}_3^{2-}$ ) concentration in  $\text{mmolc L}^{-1}$  ( $\text{meq L}^{-1}$ ) in the soil saturation paste extract (Method S - 1.00). It is based on titration with 0.10 N hydrochloric acid. The determination of  $\text{HCO}_3^{1-}$  and  $\text{CO}_3^{2-}$  should be made immediately due to the potential of the extract being super saturated relative to calcium carbonate ( $\text{CaCO}_3$ ). The concentration of  $\text{HCO}_3^{1-}$  affects the solubility of calcium, the ionic strength of the extract solution and is used to calculate the adjusted SAR (Robbins, 1990 and Hanson et al. 1993). The method detection limit is approximately  $0.05 \text{ mmolc L}^{-1}$  ( $\text{meq L}^{-1}$ ) and is generally reproducible within  $\pm 10\%$ .

### Equipment

1. Titration burette  $50.0 \pm 0.2 \text{ mL}$ , or automatic titrator.
2. pH meter and combination pH electrode.
3. Pipette,  $2.0 \pm 0.05 \text{ mL}$  and  $5.0 \pm 0.05 \text{ mL}$ .
4. 50 mL beaker.
5. Magnetic stir plate and micro size (0.25 mm) Teflon coated magnetic stir bar.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Primary standard buffer solutions: pH 4.00, 7.00 and 10.0.
3. Standardized hydrochloric acid (HCl) solution, 0.020 N with respect to  $\text{H}^+$  (See Comment #1).

### Procedure

1. Prepare a soil saturated paste extract according to Method S - 1.00 and retain extract for carbonate and bicarbonate analysis.
2. Standardize / calibrate the pH meter: (1) rinse electrode with deionized water and place in pH 7.0 primary standard buffer and adjust as necessary; (2) rinse electrode and place in pH 4.0 primary standard buffer; (3) adjust the slope until response is  $\pm 0.05$  units of expected response; and (4) and recheck standard buffers (See Comments #2 and #3).
3. Place 1.0 to 50 mL aliquot of saturation paste extract in beaker, and bring to 50 mL volume with deionized water and add magnetic stirrer. Place on stir plate and insert pH electrode (See Comment #4). Record amount of titrant needed to reach a pH of 8.3 for  $\text{CO}_3^{2-}$  and 4.5 for  $\text{HCO}_3^{1-}$  to the nearest 0.2 mL.
4. Determine the amount of  $\text{HCO}_3^{1-}$  in deionized water blank solution.

### Calculations

$$\text{CO}_3^{2-} \text{ mmolc L}^{-1} = \frac{(2 \times P \times N) \times 1000}{\text{aliquot (mL)}} \quad \text{HCO}_3^{1-} \text{ mmolc L}^{-1} = \frac{(T - (2 \times P)) \times N \times 1000}{\text{aliquot (mL)}}$$

P = number of mL of HCl of normality N to reach  $\text{CO}_3^{2-}$  inflection point, pH 8.3;

T = number of mL of HCl of normality N to reach  $\text{HCO}_3^{1-}$  inflection point, pH 4.5;

aliquot = volume of saturation paste extract sample, mL.



## Comments

1. Standardized 0.020 N HCl solution can be prepared from dilution of 1.00 N HCl standard reference solution or standardized by titration of known bases (Horneck, 1989).
2. Follow manufacturer's guidelines if meter does not read within 0.05 units of primary standards. Maintenance of combination electrodes differs from that of separate reference and glass electrodes; refer to manufacturer's instructions.
3. Store pH electrodes according to manufacturer's instructions (usual recommended practice is to store the electrodes in a primary standard buffer).

## Literature

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U.S. Salinity Laboratory Staff. 1954. Saturated soil paste. Diagnosis and improvement of saline and alkali soils. Agr. Handbook 60, USDA, Washington, D.C.

## SATURATION PASTE EXTRACT SOLUBLE CHLORIDE Chloride

S - 1.40

### Scope and Application

This method quantifies the concentration of chloride ( $\text{mmolc L}^{-1}$  or  $\text{meq L}^{-1}$ ) in the saturation paste extract (Method S - 1.00). Chloride may be determined using an ion selective electrode (potentiometric), chloridometer or ion chromatography instrument methods. Plant tolerance to chloride can be related to its concentration in the soil saturation paste extract. The method detection limit is approximately  $0.1 \text{ mmolc L}^{-1}$  dependent on the method of analysis and is generally reproducible within  $\pm 10\%$ . The unit  $\text{mmolc L}^{-1}$  is the accepted scientific unit for reporting the concentration of anions and cations and is equivalent to  $\text{meq L}^{-1}$ .

### Equipment

1. Solid-state chloride electrode and double junction reference electrode, chloridometer or Cl titrator.
2. pH/ion meter or millivolt meter.

### Reagents

1. Deionized water, ASTM Type I grade.
3. Chloride standard,  $1.0 \text{ mmolc L}^{-1}$ : Dissolve 74.1 mg of KCl in 500 mL of deionized water and dilute to 1.0 L final volume.

### Procedure

1. Prepare a soil saturated paste extract according to Method S - 1.00 and retain for chloride analysis (See Comment #1).
2. Determine the chloride concentration by ion selective electrode, chloridometer or ion chromatography. The instrument chosen will determine specific matrix modifications and sample dilutions. Adjust and operate instruments in accordance with manufacturer's instructions. Calibrate instrument using calibration solutions and determine chloride concentration of a method blank and unknown samples (See Comments #2 and #3). Report chloride concentration in saturation paste extract to the nearest  $0.1 \text{ mmolc L}^{-1}$  (See Comments #4).

### Comments

1. Care must be taken to clean all labware prior to analysis. Wash all labware with  $0.2 \text{ N HNO}_3$  and deionized water.
2. To accurately determine saturation paste chloride concentrations less than  $2.0 \text{ mmolc L}^{-1}$ , it is advisable to use standard additions techniques and potentiometric analysis (Fixen et al., 1988)
3. Samples containing chloride concentrations greater than the highest standard will require dilution.
4. Tolerance of plants to soil chloride levels in the soil saturated extract is listed in Table S-1.40-A.

Table S-1.40-A Tolerance of some plants to chloride in the soil saturated paste extract.

Crop	Chloride (mmol <sub>c</sub> L <sup>-1</sup> )
Alfalfa	23
Barley	90
Beets	90
Citrus (root stock dependent)	10-25
Corn (2-8 leaf stage)	70
Cotton	50
Grapes (Thompson Seedless)	25
Tomato	39
Wheat (young)	25

### Literature

Fixen, P.W., R.H. Gelderman and J.L. Denning. 1988. Chloride tests. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bull. No. 499 (revised).

Horneck, D. A., J. M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. *Ag. Expt. Station SM 89:4.*

Rhoades, J.D. 1982. Soluble salts. p. 167-178. *In*: A. L. Page et al. (ed.) Methods of soil analysis: Part 2. Agronomy Monogr. 9. 2nd ed. ASA and SSSA, Madison, WI.

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### **Scope and Application**

This procedure quantitatively determines the boron concentration in the soil saturation paste extract (Method S - 1.00). It is based on determination by Inductively coupled plasma emission spectrometry (ICP-AES) using one of three wavelengths. The method quantifies soil boron concentrations which can limit crop yield or be toxic to plant growth. The method is not applicable for assessing potential soil boron deficiencies. The method detection limit is approximately  $0.10 \text{ mg L}^{-1}$  and is generally reproducible to within  $\pm 8\%$ .

### **Equipment**

1. Analytical balance: 250 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. 15 mL test tube or vial, polypropylene.
3. Pipette,  $2.0 \pm 0.05 \text{ mL}$  and  $3.0 \pm 0.05 \text{ mL}$ .
4. Vortex stirring device.
5. Spectrophotometer, wavelength 420 nm or ICP-AES 249.678, 249.773 or 208.959 nm.

### **Reagents**

1. Deionized water, ASTM Type I grade.
2. Standard Boron Calibration solutions. Prepare six boron calibration standards: concentration 0.05, 0.20, 0.50, 1.0, 2.0 and  $4.0 \text{ mg L}^{-1}$ , prepared in deionized water from a standard  $1000 \text{ mg L}^{-1}$  solution.

### **Procedure**

1. Prepare a soil saturated paste according to Method S - 1.00 and allow to equilibrate twenty-four (24) hours. Retain extract for boron analysis.
2. Adjust and operate ICP-AES instrument according to manufacture's instructions. Determine B using the 249.773 nm or 249.678 nm wavelength (see Appendix A) and calibrate standards of 0.02, 0.50, 1.0 and  $4.0 \text{ mg L}^{-1}$  in deionized water matrix. Determine boron concentration of a method blank and unknown saturation paste extracts (See Comments #4 - #7).

### **Calculations**

Calculate boron concentration of saturated paste extract. Report boron concentration to the nearest  $0.01 \text{ mg L}^{-1}$  of the saturation paste extract.

### **Comments**

1. Prepare all reagents and perform all analyses in polypropylene or teflon labware. Do not use borosilicate glassware.
2. For laboratories utilizing ICP-AES instrumentation it is suggested to use a rinse between samples with of  $0.10 \text{ M D-sorbitol}$  solution.
3. Plant sensitivity to saturation paste extract boron is as follows (USDA Salinity Lab., 1954):

Table S-1.50-A Tolerance of some plants to boron in the soil saturated extract.

B mg L <sup>-1</sup>	Plant Sensitivity
< 0.7	safe for sensitive plants (peach, pear, plum)
0.7 - 1.5	moderately tolerant (cotton, wheat, bell pepper)
1.5 - 4.0	Toxic to all but tolerant plants (alfalfa, lettuce, sugar beet)
> 4.0	Generally toxic to all plants

### Literature

Horneck, D.A., J.M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. Ag. Expt. Station SM 89:4.

Watson, Maurice 1988. Recommended Soil Boron Test. p. 23-25. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures. North Dakota Agricultural Experiment Station Bulletin No. 499 (revised).

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**SATURATION PASTE EXTRACT  
CALCIUM, MAGNESIUM, SODIUM, AND SAR  
AAS or ICP-AES Method**

**S - 1.60**

### Scope and Application

This method quantitatively determines the concentration ( $\text{mmolc L}^{-1}$ ,  $\text{meq L}^{-1}$ ) of dissolved Ca, Mg and Na in the soil saturation paste extract (Method S - 1.00) using atomic absorption spectrometry (AAS) or Inductively coupled plasma emission spectrometry (ICP-AES). A chemical interference solution is used to minimize chemical matrix effects. The Sodium Absorption Ratio (SAR) of saturation paste extract is calculated from the concentration of these cations. The relationship between cation solution concentrations and exchangeable cations in the soil, is used to estimate exchangeable sodium percentage (ESP) from the SAR (Robbins, 1990). The method detection limit for these cations is approximately  $0.02 \text{ mmolc L}^{-1}$  on a solution basis and it is generally reproducible within  $\pm 7\%$ . The unit  $\text{mmolc L}^{-1}$  is the accepted scientific unit for reporting the concentration of anions and cations and is equivalent to  $\text{meq L}^{-1}$ .

### Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Atomic Absorption Spectrophotometer (AAS) Inductively coupled plasma emission spectrometry (ICP-AES) instrument.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Chemical interference solution,  $5000 \text{ mg L}^{-1}$  lanthanum oxide ( $\text{La}_2\text{O}_3$ ) -  $2000 \text{ mg L}^{-1}$ , cesium chloride (CsCl) solution. Dissolve: 4.691 g  $\text{La}_2\text{O}_3$  and 5.071 g CsCl in 1500 mL of deionized water and add 25.0 mL of  $\text{HClO}_4$  and 25.0 mL of  $\text{HNO}_3$  and dilute to 2000 mL.
3. Standard calibration solutions of Ca, Mg, and Na: Prepare six calibration solutions containing 0.05 -  $1.3 \text{ mmolc L}^{-1}$  of Na, 0.05 -  $3.5 \text{ mmolc L}^{-1}$  of Ca, and 0.02 -  $1.6 \text{ mmolc L}^{-1}$  for Mg prepared from 1000  $\text{mg L}^{-1}$  standard reference solutions and dilute to volume with chemical interference solution.

### Procedure

1. Prepare a soil saturated paste extract according to Method S - 1.00 and retain extract for cation analysis.
2. Dilute an aliquot of the saturated paste extract 10:1 with chemical interference solution (See Comment #1 and #2). For analysis by ICP-AES no chemical interference solution is required.
3. Adjust AAS or ICP-AES instrument according to manufacturer's instructions. Calibrate instrument using calibration solutions and determine individually cation (Ca, Mg, and Na) concentrations of saturation paste extracts and record as  $\text{mg L}^{-1}$  of analyte.

### Calculations

$$[\text{Ca}] \text{ mmolc L}^{-1} = \frac{\text{Ca mg L}^{-1} \times 10}{20.0 \text{ mg mmolc}^{-1}}$$

$$[\text{Mg}] \text{ mmolc L}^{-1} = \frac{\text{Mg mg L}^{-1} \times 10}{12.15 \text{ mg mmolc}^{-1}}$$

$$[\text{Na}] \text{ mmolc L}^{-1} = \frac{\text{Na mg L}^{-1} \times 10}{23.0 \text{ mg mmolc}^{-1}}$$

$$\text{SAR} = \frac{[\text{Na}]}{([\text{Ca}] + [\text{Mg}])/2}^{1/2}$$

Report Ca, Mg, and Na concentrations to the nearest  $0.1 \text{ mmolc L}^{-1}$  and SAR to the nearest 0.1 (See Comments #3, #4, #5 and #6).

### Comments

1. Saturation paste extract solutions containing greater than  $750 \text{ mg L}^{-1}$  soluble salts ( $> 1.2 \text{ dS m}^{-1}$ , estimated from  $\text{EC}_e$  Method S - 1.20) will require additional dilution.

- Cations may also be determined on the saturation paste extract using ICP-AES or ion chromatography instrumentation.
- A measure of soil sodicity, molar proportion of cation-exchange sites occupied by sodium ( $\text{Na}_{\text{exch}}$ ) can be calculated from the SAR (U. S. Salinity Laboratory, 1954). CEC can be determined using Method S-10.1 or S-10.2.

$$\text{ESP} = \frac{\text{Na}_{\text{exch}}}{\text{CEC}} \times 100 \quad [\text{equ. S -1.6-1}]$$

$$\text{ESP} = \frac{100 \times (-0.0126 + 0.0147 \times \text{SAR})}{(10 + (0.036 + 0.1051 \times \text{SAR}))} \quad [\text{equ. S -1.6-2}]$$

- Soils having an SAR greater than 13 and/or ESP > 15% are considered sodic.
- For laboratories utilizing ICP-AES instrumentation calibrate use the 422. 673 nm wavelength for Ca, 285.213 nm for Mg, and 588.995 nm wavelength for sodium (see Appendix A) using the standards of the calibration ranges described above.
- For samples that the  $\text{HCO}_3$  constitutes more than 25% of the anions it may be necessary to determine the adjusted SAR. See water method W - 1.60 to calculate.

## Literature

Kamphorst, A. and H. Bolt. 1978. Saline and sodic soils. p. 171-191 *In*: G.H. Bolt and M. G. M. Bruggenwert, (ed.) Soil chemistry. A basic elements 2nd ed. Elsevier Scientific, Amsterdam.

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U.S. Salinity Lab. Staff. 1954. Saturated soil paste. Diagnosis and Improvement of saline and alkali soils. Agr. Handbook 60, USDA, Washington, D.C.

### Scope and Application

This method quantifies the concentration of sulfate ( $\text{SO}_4^{2-}$  mmolc L<sup>-1</sup> or meq L<sup>-1</sup>) in the soil saturated paste extract (Method S - 1.00). The unit mmolc L<sup>-1</sup> is the new accepted scientific unit for reporting the concentration of anions and cations and is equivalent to meq L<sup>-1</sup>. Sulfate may be determined using turbidimetric, ion chromatography, or ICP-AES instrument methods. This method outlines the turbidimetric analysis which closely follows that described in 1992 Standard Method of the Examination of Waste Water. Sulfate is determined to evaluate anion balance in the soil saturated paste extract and estimate gypsum content. It has a method detection limit is approximately 0.02 mmolc L<sup>-1</sup> and is generally reproducible within  $\pm 7\%$ .

### Equipment

1. Magnetic stirrer.
2. Repipette dispenser calibrated to  $2.0 \pm 0.05$  mL
3. Pipette 10.0 mL.
4. Magnetic stir plate and Teflon stir bar.
5. Nephelometer (preferred), Turbidimeter or Spectrophotometer 340 nm.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Turbidimetric solution. Dissolve 30.0 g of  $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ ; 5.0 g  $\text{CH}_3\text{COONa} \cdot 3\text{H}_2\text{O}$ ; 1.0 g  $\text{KNO}_3$ ; 20 mL acetic acid,  $\text{CH}_3\text{COOH}$  (99%) and 0.111 g  $\text{Na}_2\text{SO}_4$ , in 500 mL deionized water and add 5.0 g of powered gum acacia, or gelatin (See Comment #1) suspension agent. Dilute to 1000 mL final volume.
3. Barium chloride crystals. Parr turbidimetric grade,  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  crystals 20 - 30 mesh. Use high purity  $\text{BaCl}_2$ , as low purity may result in low recovery of  $\text{SO}_4^{2-}$  (See Comment #2).
4. Standard sulfate-sulfur calibration solutions. Prepare 5.0 mmolc L<sup>-1</sup>  $\text{SO}_4^{2-}$  calibration stock solution, dissolve 0.4353 g of oven dry  $\text{K}_2\text{SO}_4$  in 500 mL of deionized water and dilute to one 1000 mL. Prepare six 100 mL calibration solutions of: 0.05, 0.1, 0.2, 0.4, 0.8, and 1.0 mmolc L<sup>-1</sup>  $\text{SO}_4^{2-}$  from a 5.0 mmolc L<sup>-1</sup>  $\text{SO}_4^{2-}$  solution and bring to final volume with deionized water.

### Procedure

1. Prepare a soil saturated paste extract according to Method S - 1.00 and retain for sulfate analysis (See Comment #3). If the aliquot is turbid, filter prior to analysis.
2. Dilute a 10.0 mL aliquot with 10.0 mL of deionized water. Repeat using sulfate standards and method blank.
3. Add 2.0 mL of turbidimetric solution using a repipette (See Comment #4). Add magnetic stir bar and beginning stirring.
4. While stirring add 0.2 g of  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  crystals with measuring spoon.
5. Stir for sixty ( $60 \pm 3$ ) seconds, then remove from stirrer and after five ( $5 \pm 0.5$ ) minutes read absorbance with nephelometer or spectrophotometer at 340 nm (See Comment #5 and #6). Repeat with sulfate calibration solutions and method blank. Using standard calibration solutions and determine sulfate concentration of saturate paste extracts and method blank. Record as mmolc L<sup>-1</sup>  $\text{SO}_4^{2-}$  of analyte in extract solution to two significant digits.



## Calculations

Report soil saturated paste extract:

$$\text{mmolc L}^{-1} \text{SO}_4^{2-} = (\text{mmolc L}^{-1} \text{SO}_4^{2-} \text{ saturated paste extract} - \text{method blank}) \times (2)$$

$$(1.0 \text{ mmolc L}^{-1} \text{SO}_4^{2-} = 48.03 \text{ mg L}^{-1} \text{SO}_4^{2-})$$

## Comments

1. A number of suspension agents have been reported in the literature which include: gum acacia, gelatin, glycerol, PVP-K30 (polyvinylpyrrolidone), and Tween 80 which have proven effective in turbidimetric analysis. Each of these will require experimentation and practice using  $\text{SO}_4$ -S spiking to fully refine the technique.
2. Use  $\text{BaCl}_2$  specifically designated for turbidimetric determination of sulfate-sulfur. Sources: J.T. Baker Cat. Parr Turbidimetric  $\text{BaCl}_2$ , JT0974-5; VWR JT0974-5; and GFS Chemicals, Reagent Grade ACS #602.
3. Care must be taken to clean all labware prior to analysis. Pre-rinse all extraction flasks, turbidimetric and spectrometer cuvette in hot water followed by 0.5 N HCl rinse with deionized water.
4. Check repipette volume, calibrate using an analytical balance.
5. Samples containing  $\text{SO}_4^{2-}$  concentrations greater than the highest standard will require dilution.
6. For laboratories utilizing ICP-AES instrumentation calibrate use the 182.669 nm wavelength and calibration standards of 0.05, 0.50, 1.0, and 5.0 mmolc  $\text{L}^{-1} \text{SO}_4^{2-}$  (see Appendix A-1).

## Literature

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- Standard methods for the examination of waste water. 1992. pp.131 - 132. Arnold Greendery, Lenore S. Clescerl and Andrew D. Eaton (ed.) 18th ed. American Public Health Association, American Water Works Association, and Water Pollution Control Federation.

### Scope and Application

This method quantifies the concentration of nitrate (NO<sub>3</sub><sup>-</sup>) (mmolc L<sup>-1</sup> or meq L<sup>-1</sup>) in the saturation paste extract (Method S - 1.00). Nitrate may be determined using an ion selective electrode (ISE, see Method S-3.20), ion chromatography or cadmium reduction spectrophotometric methods. This method outlines the use of the cadmium reduction spectrophotometric method (automated) outlined by (Keeney, 1982). The method detection limit is approximately 0.04 mmolc L<sup>-1</sup> dependent on the method of analysis and is generally reproducible within ± 10%. Nitrate is determined to for anion balance and crop nitrogen nutrient status. The unit mmolc L<sup>-1</sup> is the accepted scientific unit for reporting the concentration of anions and cations and is equivalent to meq L<sup>-1</sup>.

### Equipment

1. Spectrophotometer, autoanalyzer, or flow injection analyzer (FIA) instrument.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Standard calibration solutions of NO<sub>3</sub>-N. Prepare six calibration standards ranging from 0.05 to 1.5 mmolc L<sup>-1</sup> concentration, diluted in 0.05 N CaCl<sub>2</sub> solution prepared from 16.1 mmolc L<sup>-1</sup>(1000 mg L<sup>-1</sup>) NO<sub>3</sub><sup>-</sup> standard solution.

### Procedure

1. Prepare a soil saturated paste extract according to Method S - 1.00 and retain for nitrate analysis (See Comment #1).
2. Nitrate (NO<sub>3</sub><sup>-</sup>) content of the extract is determined using a spectrophotometer, automated flow analyzer (Technicon Method No. 329-74W/A) or FIA instrument. Calibrate using standard calibration solutions and operate instrument in accordance with manufacturer instructions. Determine nitrate concentration of saturated paste extract, method blank, unknown samples and record results as mg L<sup>-1</sup> of nitrate in extract solution (See Comment #2)

### Calculations

Report soil saturated paste extract:

$$\text{mmolc L}^{-1} \text{ NO}_3^- = (\text{mmolc L}^{-1} \text{ NO}_3^- \text{ saturated paste extract} - \text{method blank})$$

$$(1.0 \text{ mmolc L}^{-1} \text{ NO}_3^- = 62.0 \text{ mg L}^{-1} \text{ NO}_3^-)$$

### Comments

1. Care must be taken to clean all labware prior to analysis. Wash all labware with 0.1 N HCl and deionized water.
2. Samples containing nitrate concentrations greater than the highest standard will require dilution.

## Literature

Bremner, J. M. and D.R. Keeney. 1965. Determination and isotopic ratio analysis of different forms of nitrogen in soils: I. Apparatus and procedure for distillation for and determination of ammonium. *Soil Sci. Soc. Am. Proc.* 29:504-507.

Dahnke, W.C. 1990. Testing soils for available nitrogen. p. 120-140. *In: R.L. Westerman (ed.) Soil testing and plant analysis. Soil Sci. Soc. Am. Book series 3 ASA Madison WI.*

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Rhoades, J.D. 1982. Soluble salts. p. 167-178. *In: A. L. Page et al. (ed.) Methods of soil analysis: Part 2. Agronomy Monogr. 9. 2nd ed. ASA and SSSA, Madison, WI.*

## Scope and Application

This method semi-quantifies the pH of soil, using a 1:1 soil:water extract of the soil using deionized water. Soil pH is a measure of the relative acidity or alkalinity of the soil solution that is in equilibrium with the solid particles. It is a measure of the intensity of acidity or alkalinity, but does not indicate the relative buffering capacity of the soil. It is most applicable to soils with a pH ranging from 4.0 to 9.0. This method differs from the saturation paste soil pH (Method S - 1.10) and soil pH 1:2 (Method S - 2.20) in that pH is determined using a 1:1 soil:water ratio and determines the pH of the supernate for a soil slurry. Generally soil pH obtained using the 1:1 ratio is 0.15 to 0.25 pH units higher than that obtained using the saturated paste extract but lower than that obtained by the 1:2 dilution. Soil pH is measured to assess soil chemical properties, crop suitability, lime needs and relative nutrient availability. The method is generally reproducible within  $\pm 0.12$  pH units.

## Equipment

1. Paper cups 3 oz or 100 mL beaker.
2. Analytical balance, 250 g capacity, resolution  $\pm 0.1$  g.
3. pH meter, equipped with pH glass electrodes (indicating and reference).

## Reagents

1. Deionized water, ASTM Type I grade.
2. Primary standard buffers, pH 4.0, 7.0, and 10.0

## Procedure

1. Weigh  $5.0 \pm 0.1$  g of air dry soil pulverized to pass 10 mesh sieve ( $< 2.0$  mm) into a 100 mL beaker or 3 oz plastic cup (See Comment #1).
2. Add 5.0 mL of deionized water and stir thoroughly for 5 seconds.
3. Let stand fifteen (15) minutes, allow suspended soil particles to settle before reading pH.
4. Standardize / calibrate the pH meter: (1) rinse electrode with deionized water and place in pH 7.0 primary standard buffer and adjust as necessary; (2) rinse electrode and place in pH 4.0 primary standard buffer; (3) adjust the slope until response is  $\pm 0.05$  units of expected response; and (4) Check pH 7.0 primary standard buffer and adjust as necessary (See Comment #2). For high pH soils ( $> 7.0$ ) use pH buffers 7.0 and 10.0.
5. Read the pH by placing the electrodes in the slurry, swirling gently and read the pH immediately. Ensure the electrode tips are in the slurry and not the supernate. Report soil pH to the nearest 0.01 unit.
6. Rinse the electrodes with deionized water and blot dry between pH determinations. Do not wipe the electrode (See Comment #3, #4 and #5).
7. When the meter is not in use, immerse the electrodes in pH 4.00 buffer.

## Calculations

Record soil pH as  $\text{pH}_{1:1}$

## Comments

1. This follows the procedure outlined by Eckert (1989) in Recommended Chemical Tests Procedures for the North Central Region.
2. Follow manufacturer's guidelines if meter does not read within 0.05 units of primary standards. Maintenance of combination electrodes differs from that of separate reference and glass electrodes; refer to manufacturer's instructions.
3. pH electrodes will equilibrate faster in fine textured soils (clay constant > 20%, See Method S-14.1) and soils high in soluble salts (> 1.0 dS m<sup>-1</sup>. See Method 1.10).
4. Store pH electrodes according to manufacturer's instructions (recommended practice is to store the electrodes in a primary standard buffer). Combination electrodes should be stored in a pH 4.0 buffered solution containing 5.0 g L<sup>-1</sup> potassium chloride.
5. To determine soil buffer pH add SMP (See Method S-2.50) or Woodruff buffer (See Method S-2.60) reagent to the soil slurry.

## Literature

Eckert, D. J. 1988. Recommended pH and Lime Requirement Tests. p. 6-8. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bulletin No. 499 (revised)..

Kalra, Y. 1995. Determination of pH of soils by different methods: collaborative study. *J. of AOAC International*. 78:310-320.

McLean, E.O. 1982. Soil pH and lime requirement. p. 199-223. *In* A.L. Page et al. (ed.) Methods of soil analysis, part 2. Agronomy Monogr. 9, 2nd ed. ASA and SSSA, Madison, WI.

Rhoades, J.D. 1982. Soluble salts. p. 167-178. *In*: A. L. Page et al. (ed.) Methods of soil analysis: Part 2. Agronomy Monogr. 9. 2nd ed. ASA and SSSA, Madison, WI.

Rhoades, J.D. and S. Miyamoto. 1990. Testing soils for salinity. p. 299-336. *In*: R.L. Westerman (ed.) Soil testing and plant analysis. 3rd ed. SSSA, Madison, WI.

U.S. Salinity Lab. Staff. 1954. Saturated Soil Paste. Diagnosis and Improvement of saline and alkali Soils. Agr. Handbook 60, USDA, Washington, D.C.

### Scope and Application

This method semi-quantifies the pH of soil, using a 1:1 soil:water extract of the soil using a 0.01 M CaCl<sub>2</sub> solution. Soil pH is a measure of the relative acidity or alkalinity of the soil solution that is in equilibrium with the solid particles. It is a measure of the intensity of acidity or alkalinity, but does not indicate the relative buffering capacity of the soil. It is most applicable to soils with a pH ranging from 4.0 to 9.0. This method differs from the saturation paste soil pH (Method S - 1.10) and soil pH 1:1 (Method S - 2.20) in that pH is determined using a salt suspension of 0.01 M CaCl<sub>2</sub>. This method has advantages in that the pH is determined independent of the soluble salt concentration of soils, as the suspension remains flocculated errors associated with liquid junction potential are minimized. The pH is more reproducible than the 1:1 water method on soils low in soluble salts ( $EC_e < 0.4 \text{ dS m}^{-1}$ ). Soil pH is measured to access soil chemical properties, crop suitability, lime needs and relative nutrient availability. The method is generally reproducible within  $\pm 0.06$  pH units. Soil pH is measured to access soil chemical properties, crop suitability, lime needs and relative nutrient availability. The method is generally reproducible within  $\pm 0.12$  pH units.

### Equipment

1. Paper cups 3 oz or 100 mL beaker.
2. Analytical balance, 250 g capacity, resolution  $\pm 0.1$  g.
3. pH meter, equipped with pH glass electrodes (indicating and reference).

### Reagents

1. Calcium chloride, 0.01 M CaCl<sub>2</sub>: Dissolve 2.940 g of calcium chloride dihydrate (CaCl<sub>2</sub> · 2H<sub>2</sub>O) with deionized water in a 2 L volumetric flask and dilute to volume.  $EC_e$  of solution should be between 2.24 and 2.40  $\text{dS m}^{-1}$  at 25° C.
2. Primary standard buffers, pH 4.0, 7.0, and 10.0.

### Procedure

1. Weigh  $5.0 \pm 0.1$  g of air dry soil pulverized to pass 10 mesh sieve (< 2.0 mm) into a 100 mL beaker or 3 oz plastic cup (See Comment #1).
2. Add 5.0 mL of 0.01 M CaCl<sub>2</sub> solution and stir thoroughly for 5 seconds.
3. Let stand fifteen (15) minutes, allow suspended soil particles to settle before reading pH.
4. Standardize / calibrate the pH meter: (1) rinse electrode with deionized water and place in pH 7.0 primary standard buffer and adjust as necessary; (2) rinse electrode and place in pH 4.0 primary standard buffer; (3) adjust the slope until response is  $\pm 0.05$  units of expected response; and (4) Check pH 7.0 primary standard buffer and adjust as necessary (See Comment #2). For high pH soils (> 7.0) use pH buffers 7.0 and 10.0.
5. Read the pH by placing the electrodes in the slurry, swirling gently and read the pH immediately. Ensure the electrode tips are in the slurry and not the supernate. Report soil pH to the nearest 0.01 unit.
6. Rinse the electrodes with deionized water and blot dry between pH determinations. Do not wipe the electrode (See Comment #2, #3 and #4).
7. When the meter is not in use, immerse the electrodes in pH 4.00 buffer.

### Calculations

Record soil pH as pH<sub>1:1 CaCl</sub>

## Comments

1. Follow manufacturer's guidelines if meter does not read within 0.05 units of primary standards. Maintenance of combination electrodes differs from that of separate reference and glass electrodes; refer to manufacturer's instructions.
2. pH electrodes will equilibrate faster in fine textured soils (clay constant > 20%, See Method S-14.1) and soils high in soluble salts (> 1.0 dS m<sup>-1</sup>. See Method 1.10).
3. Store pH electrodes according to manufacturer's instructions (recommended practice is to store the electrodes in a primary standard buffer). Combination electrodes should be stored in a pH 4.0 buffered solution containing 5.0 g L<sup>-1</sup> potassium chloride.
4. To determine soil buffer pH add SMP (See Method S-2.50) or Woodruff buffer (See Method S-2.60) reagent to the soil slurry.

## Literature

- Eckert, D. J. 1988. Recommended pH and Lime Requirement Tests. p. 6-8. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bulletin No. 499 (revised)..
- Kalra, Y. 1995. Determination of pH of soils by different methods: collaborative study. *J. of AOAC International*. 78:310-320.
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- Rhoades, J.D. 1982. Soluble salts. p. 167-178. *In*: A. L. Page et al. (ed.) Methods of soil analysis: Part 2. Agronomy Monogr. 9. 2nd ed. ASA and SSSA, Madison, WI.
- Rhoades, J.D. and S. Miyamoto. 1990. Testing soils for salinity. p. 299-336. *In*: R.L. Westerman (ed.) Soil testing and plant analysis. 3rd ed. SSSA, Madison, WI.
- U.S. Salinity Lab. Staff. 1954. Saturated Soil Paste. Diagnosis and Improvement of saline and alkali Soils. Agr. Handbook 60, USDA, Washington, D.C.

### Scope and Application

This method semi-quantifies the electrical conductivity (  $EC_{1:1}$  ) of soil, using a 1:1 soil:water extract (volume to volume) of the soil using deionized water. EC is measured using a conductivity probe. This method differs from the saturation paste  $EC_e$  (Method S - 1.20), and is used as a simplified method for determining soluble salts but is less well related to field soil water composition and content. This method is not recommended for gypsiferous soils and will lead to bias high results. The  $EC_e$  measurement is sensitive to temperature and increases approximately 1.9% per °C (range 15 - 35 °C) (Rhoades, 1996). All  $EC_e$  data is normalized to 25 °C. Soil  $EC_{1:1}$  is measured to access soil chemical properties, crop suitability, lime needs and relative nutrient availability. The method is generally reproducible within  $\pm 5\%$ .

### Equipment

1. Paper cups 3 oz or 100 mL beaker.
2. 10 cm<sup>3</sup> scoop, volumetric.
3. Conductance cell and conductance meter with dynamic range from 0.01 to 100 dS m<sup>-1</sup> conductance, temperature compensating, 25 °C.

### Reagents

1. Deionized water CO<sub>2</sub>-free, ASTM Type I grade.  $EC < 10^{-4}$  dS m<sup>-1</sup>.
2. Standard Reference Calibration Solution. Dissolve 0.7456 g KCl (previously dried at 110 °C for 2 h) in CO<sub>2</sub>-free deionized water and dilute to 1.0 L. At 25  $\pm 0.1$  °C a 0.010 N KCl solution will have an  $EC_e$  of 1.412 dS m<sup>-1</sup> (mmhos cm<sup>-1</sup>). For a 0.100 N KCl solution (7.456 g KCl diluted to 1.0 L) will have an  $EC_e$  of 12.900 dS m<sup>-1</sup>. Standard EC calibration solutions are listed in Table S-2.30-A and can be purchased from a scientific supply vendor.

### Procedure

1. Using a 10 cm<sup>3</sup> scoop measure two scoops of air dry soil pulverized to pass 10 mesh sieve (< 2.0 mm) into a 100 mL beaker or 3 oz plastic cup.
2. Add 20 mL of deionized water.
3. Let stand fifteen (15) minutes, allow suspended soil particles to settle before reading  $EC_{1:1}$  (See Comment #1).
4. Calibrate conductance cell. Operate and adjust instrument in accordance with manufacturer's instructions (See Comments #2, #3 and #4). Rinse conductance cell with three aliquots of 0.01 N KCl, adjust a fourth portion to 25  $\pm 0.1$  °C, measure R (where R is the measured resistance ohms) and temperature  $t$ . Repeat measurement of R until value is constant. Calculate cell constant K.

$$K = (0.001413) (R_{KCl}) / [1 + 0.019(25 - t)]$$

5. Rinse conductance cell with deionized water. Stir conductance probe into extract solution. When the meter has stabilized record instrument reading.

### Calculations

$$EC_{1:1} = C_x(1000)K[1 + 0.019(25 - t)]$$

Where:  $C_x$  is the measured C of the sample and  $t$  is temperature



Report  $EC_{1:1}$  to the nearest  $0.01 \text{ dS m}^{-1}$  as  $EC_e$   $25^\circ\text{C}$ . (See Comments #4, and #5)

Table S - 2.30 -A Conductivity of KCl solutions at  $25^\circ\text{C}$  (Rhoades, 1996).

Concentration N	Conductivity dS m
0.001	0.147
0.010	1.413
0.020	2.767
0.050	6.668
0.10	12.90
0.20	24.82
0.50	58.64

### Comments

1. Exposure of the sample to the atmosphere may cause changes in conductivity due to loss or gain of dissolved gasses:  $\text{CO}_2$  and  $\text{NH}_3\text{-N}$ . Freshly distilled water has a conductivity of  $0.005 - 0.002 \text{ dS m}^{-1}$  increasing after a few weeks to  $0.002 - 0.004 \text{ dS m}^{-1}$ . This of special concern on samples with very low  $EC_{1:1}$ .
2. Clean platinum electrodes that are new or that are providing erratic EC readings with acid-dichromate cleaning solution. Cleaning solution: 32 mL of saturated sodium dichromate ( $\text{Na}_2\text{Cr}_2\text{O}_7$ ) and 1L 16 M sulfuric acid. Soak electrodes 16 hours followed by three rinses of deionized water rinses. If platinum is flaked, recoat according to procedure of APHA (1985).
3. For highly saline soils ( $EC_e > 8.0 \text{ dS m}^{-1}$ ) calibrate using  $0.100 \text{ N}$  KCl solution,  $EC_e$   $12.90 \text{ dS m}^{-1}$ .
4. The conductance probe is affected by the wall of the container. This is called "field effect". Therefore it is important to calibrate and analyze the samples and standards at the same position from the bottom and side wall of the container.

### Literature

APHA (1985). Part 205. In Standard Methods for the Examination of Water and Wastewater'. 16<sup>th</sup> edn. American Public Health Association, Washington, DC.

Dahnke, W.C. and D.A. Whitney. 1988. Recommended pH and lime requirement tests. p. 32-34. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bulletin No. 499 (revised).

Helrich, K. (ed.) 1990. Official methods of analysis. 15<sup>th</sup> Ed. Association of Official Analytical Chemists, Inc. Arlington, VA.

Kalra, Y and D.G. Maynard. 1991. Methods manual for forest soil and plant analysis. P. 32-36. Forestry Canada, Northwest Region Information Report NOR-X-319.

Rayment, G.E. and F.R. Higginson. 1992. Electrical conductivity. p. 15-16. *In*: Australian Laboratory Handbook of Soil and Water Chemical Methods. Inkata Press, Melbourne

Rhoades, J.D. 1996. Lime requirement. p. 417-435. *In*: J. M. Bartels et al. (ed.) Methods of soil analysis: Part 3. Chemical methods. 3rd ed. ASA and SSSA, Madison, WI. Book Series

### Scope and Application

This method semi-quantifies the electrical conductivity (  $EC_{1:2}$  ) of soil, using a 1:2 soil:water extract (volume to volume) of the soil using deionized water. EC is measured using a conductivity probe. This method differs from the saturation paste  $EC_e$  (Method S - 1.20), and is used as a simplified method for determining soluble salts but is less well related to field soil water composition and content. This method is not recommended for gypsiferous soils and will lead to bias high results. The  $EC_e$  measurement is sensitive to temperature and increases approximately 1.9% per °C (range 15 - 35 °C) (Rhoades, 1996). All  $EC_e$  data is normalized to 25 °C. Soil  $EC_{1:2}$  is measured to access soil chemical properties, crop suitability, lime needs and relative nutrient availability. The method is generally reproducible within  $\pm 8\%$ .

### Equipment

1. Paper cups 3 oz or 100 mL beaker.
2. 10 cm<sup>3</sup> scoop, volumetric.
3. Conductance cell and conductance meter with dynamic range from 0.01 to 100 dS m<sup>-1</sup> conductance, temperature compensating, 25 °C.

### Reagents

1. Deionized water CO<sub>2</sub>-free, ASTM Type I grade.  $EC < 10^{-4}$  dS m<sup>-1</sup>.
2. Standard Reference Calibration Solution. Dissolve 0.7456 g KCl (previously dried at 110 °C for 2 h) in CO<sub>2</sub>-free deionized water and dilute to 1.0 L. At 25  $\pm$  0.1 °C a 0.010 N KCl solution will have an  $EC_e$  of 1.412 dS m<sup>-1</sup> (mmhos cm<sup>-1</sup>). For a 0.100 N KCl solution (7.456 g KCl diluted to 1.0 L) will have an  $EC_e$  of 12.900 dS m<sup>-1</sup>. Standard EC calibration solutions are listed in Table S-2.30-A and can be purchased from a scientific supply vendor.

### Procedure

1. Using a 10 cm<sup>3</sup> scoop measure two scoops of air dry soil pulverized to pass 10 mesh sieve (< 2.0 mm) into a 100 mL beaker or 3 oz plastic cup.
2. Add 40 mL of deionized water.
3. Let stand fifteen (15) minutes, allow suspended soil particles to settle before reading  $EC_{1:2}$  (See Comment #1).
4. Calibrate conductance cell. Operate and adjust instrument in accordance with manufacturer's instructions (See Comments #2, #3 and #4). Rinse conductance cell with three aliquots of 0.01 N KCl, adjust a fourth portion to 25  $\pm$  0.1 °C, measure R (where R is the measured resistance ohms) and temperature  $t$ . Repeat measurement of R until value is constant. Calculate cell constant K.

$$K = (0.001413) (R_{KCl}) / [1 + 0.019(25 - t)]$$

5. Rinse conductance cell with deionized water. Stir conductance probe into extract solution. When the meter has stabilized record instrument reading.

### Calculations

$$EC_{1:2} = C_x(1000)K[1 + 0.019(25 - t)]$$

Where:  $C_x$  is the measured C of the sample and  $t$  is temperature

Report  $EC_{1:2}$  to the nearest  $0.01 \text{ dS m}^{-1}$  as  $EC_e$   $25^\circ\text{C}$ .

(See Comments #4, and #5)

Table S - 2.30 -A Conductivity of KCl solutions at  $25^\circ\text{C}$  (Rhoades, 1996).

Concentration N	Conductivity dS m
0.001	0.147
0.010	1.413
0.020	2.767
0.050	6.668
0.10	12.90
0.20	24.82
0.50	58.64

### Comments

1. Exposure of the sample to the atmosphere may cause changes in conductivity due to loss or gain of dissolved gasses:  $\text{CO}_2$  and  $\text{NH}_3\text{-N}$ . Freshly distilled water has a conductivity of  $0.005 - 0.002 \text{ dS m}^{-1}$  increasing after a few weeks to  $0.002 - 0.004 \text{ dS m}^{-1}$ . This of special concern on samples with very low  $EC_e$ .
2. Clean platinum electrodes that are new or that are providing erratic EC readings with acid-dichromate cleaning solution. Cleaning solution: 32 mL of saturated sodium dichromate ( $\text{Na}_2\text{Cr}_2\text{O}_7$ ) and 1L 16 M sulfuric acid. Soak electrodes 16 hours followed by three rinses of deionized water rinses. If platinum is flaked, recoat according to procedure of APHA (1985).
3. For highly saline soils ( $EC_e > 8.0 \text{ dS m}^{-1}$ ) calibrate using  $0.100 \text{ N}$  KCl solution,  $EC_e$   $12.90 \text{ dS m}^{-1}$ .
4. The conductance probe is affected by the wall of the container. This is called "field effect". Therefore it is important to calibrate and analyze the samples at the same position from the bottom and side wall of the container.
5. Plant/crop tolerances to salinity of the soil 1:2 extract electrical conductivity ( $EC_{1:2}$ ) are shown in Table S -2.30 - B.

Table S - 2.30-B Impact of salinity of the soil 1:2 extract electrical conductivity (EC<sub>1:2</sub>) on plant growth.

dS m <sup>-1</sup>	Plant salinity effects
< 0.4	Most crops will grow well, no injury (Pear, peach apple plum vetch, beans)
0.4 - 0.8	Very slightly saline. Yields of crops of low salt tolerance maybe reduced by 50% (Ladino clover, red clover, red fox tail, soybeans, strawberry, and orange.
0.8 - 1.6	Slightly saline. Yields of fruit and vegetable crops of medium salt tolerance maybe reduced 50% (orchard grass, birdsfoot trefoil, sunflower, corn rice sorghum, oats, cucumber, onion, carrot lettuce, bell pepper, potato, broccoli, cantaloupe, grape, and olive).
1.6 - 2.4	Moderately saline. Yield of virtually all fruit crops significantly reduced. Yield reductions of 50% may occur in the most sensitive forage and field crops. (barley, wheatgrass, cotton rape, sugarbeet, spinach, asparagus, and beets).
2.4 - 3.2	Strongly saline. Only highly salt-tolerant forage and field crops will yield satisfactorily.
> 3.2	Very strongly saline. Only a few highly salt-tolerant grasses, herbaceous plants and certain shrubs and trees will grow (salt grass, alkali sacaton)

## Literature

APHA (1985). Part 205. In Standard Methods for the Examination of Water and Wastewater'. 16<sup>th</sup> edn. American Public Health Association, Washington, DC.

Dahnke, W.C. and D.A. Whitney. 1988. Recommended pH and lime requirement tests. p. 32-34. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bulletin No. 499 (revised).

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Rhoades, J.D. 1996. Lime requirement. p. 417-435. *In*: J. M. Bartels et al. (ed.) Methods of soil analysis: Part 3. Chemical methods. 3rd ed. ASA and SSSA, Madison, WI. Book Series

## Scope and Application

Buffer pH is the measure of a soil's active and reserve acidity, i.e. buffer capacity and is used to estimate lime recommendations. The method is based on the reaction of soil buffered acidity with a chemical buffer resulting in change in the pH of the buffer. Several tests have been developed to measure lime requirement including SMP Buffer, Woodruff (1967), Mehlich (1939) and Adams & Evans (1962). The SMP (Shoemaker, McLean & Pratt, 1961) buffer tests is one of the more popular lime requirement tests used for estimating exchange acidity including that associated with exchangeable aluminum and is used predominately on soils of the northeast. Others such as the Woodruff and Mehlich method, are dependent on geographic region or preference. Standard calibration curves exist for liming based on a SMP value to a desired pH for soil groups in a geographic area. Local calibration of the method is desirable. The procedure is generally reproducible with in 0.1 pH units.

## Equipment

1. Analytical balance: 100.0 g capacity, resolution  $\pm 0.01$  g.
2. Repipette dispenser(s): calibrated to  $5.0 \pm 0.2$  mL and  $10.0 \pm 0.2$  mL.
3. pH meter, equipped with pH electrodes (indicating and reference).
4. Primary standard buffers, pH 4.0, 7.0, and 10.0.
5. Container 50 mL (polypropylene or waxed paper).
6. Glass stirring rod.

## Reagents

1. Deionized water, ASTM type I Grade.
2. Sodium Hydroxide, 0.1 N NaOH - Dissolve 4.0 g of NaOH pellets in about 500 mL deionized water. Allow to cool to room temperature and bring to 1000 mL volume.
3. SMP Buffer Solution, pH  $7.5 \pm 0.1$ : Using a 1000 mL volumetric flask, completely dissolve 1.8 g of ground para-nitrophenol in 500 mL deionized water. Add 2.8 g of triethanolamine TEA (weigh rather than pipette vicious liquid). Then dissolve 3.0 g potassium chromate ( $K_2CrO_4$ ), 2.0 g calcium acetate,  $(CH_3COO)_2Ca \cdot H_2O$  and 53.1 g calcium chloride dihydrate ( $CaCl_2 \cdot 2H_2O$ ) in the solution. Bring to 975 mL volume with deionized water and stir with magnetic stirrer (can take up to 12 hours). Adjust solution to pH 7.50 with 0.1 N NaOH or 4 M HCl if necessary and dilute to 1000 mL volume with deionized water. CAUTION: Triethanolamine and potassium chromate are hazardous materials, consult MSDS sheet before using. Verify buffer capacity by titrating 20 mL of SMP buffer from pH 7.50 to pH 5.00 with standardized 0.1 M HCl. Should require  $0.28 \pm 0.005$  cmol<sub>c</sub> of HCl / pH unit.
4. pH standard buffers, pH 4.0 and 7.0.

## Procedure

1. Weigh  $5.0 \pm 0.05$  g of air-dried soil pulverized to pass 10 mesh sieve ( $< 2.0$  mm) into 50 mL container. Generally, samples are placed in rows of six to accommodate continuous stirring and reading samples.
2. Add 5.0 mL of deionized water. Stir (leaving a stir rod in each sample) and allow to soak for thirty (30) minutes.
3. Add 10.0 mL of SMP Buffer Solution (See Comment #1) and stir every five (5) min during the ensuing twenty (20) minutes period.
4. Standardize / calibrate pH meter: (1) rinse electrode with deionized water and place in pH 7.0 primary standard buffer and adjust as necessary; (2) rinse electrode and place in pH 4.0 primary standard buffer; (3) adjust the slope until response is  $\pm 0.05$  units of expected response; and (4) Check pH 7.0 primary standard buffer and adjust as necessary (See Comment #2).

5. Immediately following the final stirring (twenty (20) minutes after addition of SMP buffer solution), insert the electrodes and observe the pH reading of the suspension, swirl gently and observe the subsequent reading. Continue until pH readings are constant, then record the pH reading to the nearest 0.1 unit as pH<sub>smp</sub> (See Comment #3 and #4).
6. Between readings, thoroughly rinse electrodes with deionized water and pat dry. Consult Table S - 2.50-A for lime requirement.

Table S - 2.50-A. Calibrations for lime requirement for the surface 20 cm of soil using the SMP buffer pH method.

Soil SMP buffer pH	Desired Soil pH			
	7	6.5	6	Organic Soil 5.2
	Amount of 100 % CaCO <sub>3</sub> required (tons ac <sup>-1</sup> )			
6.8	1.1	0.9	0.8	0.6
6.7	1.8	1.6	1.3	1.0
6.6	2.4	2	1.7	1.3
6.5	3.1	2.6	2.1	1.7
6.4	4.0	3.4	2.8	2.1
6.3	4.7	4	3.3	2.5
6.2	5.4	4.6	3.7	2.9
6.1	6	5	4.1	3.2
6.0	5.8	5.7	4.7	3.6
5.9	7.7	6.5	5.3	4.1
5.8	8.3	7	5.7	4.4
5.7	9	7.6	6.2	4.7
5.6	9.7	8.2	6.7	5.2
5.5	10.4	8.8	7.2	5.5
5.4	11.3	9.6	7.8	6
5.3	11.9	10	8.2	6.3
5.2	12.7	10.7	8.7	6.7
5.1	13.6	11.5	9.2	7.1

Based on 8 inch furrow slice weighing 2.4 million pounds. From Eckert 1988.

### Comments

1. Check repipette dispensing volume calibration using an analytical balance.
2. Follow manufacturer's guidelines if meter does not read within 0.05 units of primary standards. Maintenance of combination electrodes differs from that of separate reference and glass electrodes; refer to manufacturer's instructions. Store pH electrodes according to manufacturer's instructions (recommended practice is to store the electrodes in a primary standard buffer).

3. Reading soil-buffer solution pH between 20 and 25 min after the addition of the SMP buffer is necessary because the pH of the suspension will continue to decrease over time. The electrodes should be rinsed occasionally with 0.1 N HCl and deionized water when making a series of determinations to eliminate increased pH readings caused by electrode contamination. The method outlined is a modification of the method described by McLean (1982).
4. SMP Buffer solution is classified as hazardous waste and must be disposed of in a suitable manner.

## Literature

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McLean, E.O. 1982. Soil pH and lime requirement. *In: A.L. Page, R.H. Miller, and D.R. Keeney (eds) Methods of soil analysis. Part 2. Agron. Monogr. 9, Am. Soc. Agron., Madison, WI. p. 199-223.*

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### Scope and Application

This method involves the quantitative extraction of nitrate ( $\text{NO}_3\text{-N}$ ) from soils using 2.0 N KCl. Nitrate is determined by reduction to nitrite ( $\text{NO}_2\text{-N}$ ) via a cadmium reactor, diazotized with sulfanilamide and is coupled to N-(1-Naphthyl)-ethylenediamine dihydrochloride to form an azochromophore (red-purple in color) measured spectrophotometrically at 520 nm. The method is readily adapted to manual or automated techniques. The procedure outlined follows that outlined by Keeney and Nelson (1982) for determining nitrate nitrogen with a modification in which 25 mL of KCl and 5.0 g of soil are used instead of 100 mL and 10 g soil. Extending the shaking period to thirty minutes with 2.0 N KCl (Bremner et.al. 1965), permits the simultaneous extraction of ammonium and nitrate. Care must be taken to avoid contamination from filter paper and operator handling. Cadmium is a hazardous material, follow manufacturers recommendations in handling this material. Soil nitrate-nitrogen can be used to predict crop response to nitrogen fertilizers. The method detection limit is approximately  $0.5 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible  $\pm 6\%$ .

### Equipment

1. Analytical balance,; 100.0 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Repipette dispenser, calibrated to  $25.0 \pm 0.2 \text{ mL}$ .
3. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute (opm).
4. Extraction vessels and associated filtration vessel.
5. Whatman No. 42 or equivalent highly retentive filter paper.
6. spectrophotometer, autoanalyzer, or flow injection analyzer (FIA) instrument.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Potassium chloride extracting solution, 2.0 N KCl: Dissolve 150 g of reagent grade KCl in 500 mL deionized water and dilute to a 1000 mL (See Comment #1).
3. Standard calibration solutions of  $\text{NO}_3\text{-N}$ . Prepare six calibration standards ranging from 0.1 to 20.0  $\text{mg L}^{-1}$  concentration, diluted in 2.0 N KCl extraction solution prepared from 1000  $\text{mg L}^{-1}$   $\text{NO}_3\text{-N}$  standard solution.

### Procedure

1. Weigh  $5.0 \pm 0.05 \text{ g}$  of air-dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into extraction vessel. Add 25.0 mL of 2.0 N KCl extraction reagent using repipette dispenser (See Comment #2). Include a method blank.
2. Place extraction vessel(s) on reciprocating mechanical shaker for thirty (30) minutes.
3. Filter extract (See Comment #3), refilter if filtrate is cloudy (comment #4).
4. Nitrate-N content of the extract is determined using a spectrophotometer, automated flow analyzer (Technicon Method No. 329-74W/A) or FIA instrument. Calibrate using standard calibration solutions and operate instrument in accordance with manufacturer instructions. Determine nitrate concentration of KCl extract, method blank, unknown samples and record results as  $\text{mg L}^{-1}$  of nitrate in extract solution (See Comment #5).

### Calculation

$$\text{NO}_3\text{-N mg kg}^{-1} \text{ in soil} = (\text{NO}_3\text{-N mg L}^{-1} \text{ in filtrate} - \text{method blank}) \times 5$$

Report soil nitrate concentration to the nearest  $0.1 \text{ mg kg}^{-1}$  (See Comment #6)



## Comments

1. Soils may be extracted with 1.0 N KCl for the determination of nitrate only.
2. Check repipette dispensing volume calibration using an analytical balance.
3. Check filter paper supply for possible contamination of and  $\text{NO}_3\text{-N}$ . If contamination is greater than 0.2  $\text{mg L}^{-1}$  on a solution basis, rinse filter paper with 2.0 N KCl.
4. Soil KCl extract may be stored up to three weeks if stored at 4 °C and/or with 100  $\mu\text{L}$  of toluene or thymol.
5. Samples having nitrate concentrations exceeding the highest standard will require dilution and reanalysis.
6. Nitrate-nitrogen ( $\text{NO}_3\text{-N}$ ) results can be expressed on a volume basis. Assuming the sample represents a 0-6 inch (0-15 cm) depth of the soil, then:  $\text{NO}_3\text{-N mg kg}^{-1} \times 2.0 \cong \text{NO}_3\text{-N lbs ac}^{-1}$

## Literature

Bremner, J. M. and D.R. Keeney. 1965. Determination and isotopic ratio analysis of different forms of nitrogen in soils: I. Apparatus and procedure for distillation for and determination of ammonium. *Soil Sci. Soc. Am. Proc.* 29:504-507.

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Keeney, D.R. and D.W. Nelson. 1982. Nitrogen - inorganic forms. *In* A.L. Page (eds.) *Methods of soil analysis, part 2.* Agron. Monogr. 9, 2nd ed. ASA and SSSA, Madison, WI. p. 643-698.

### Scope and Application

This method involves the quantitative extraction of nitrate ( $\text{NO}_3\text{-N}$ ) from soils using 40 g (0.02  $\text{N}$ )  $\text{CaSO}_4$  solution. Nitrate is determined by reaction with chromotropic acid to form an azochromophore (yellow dye) which is measured spectrophotometrically at 420 nm (Kowalenko and Lowe, 1973). The method has an interference from chloride which can be masked with addition of antimony (Sb). The procedure outlined follows that outlined by Simms and Jackson (1971). If nitrite concentrations are high ( $> 2.0 \text{ mg kg}^{-1}$ ) an urea-sulfite solution may be added which converts nitrite ( $\text{NO}_2\text{-N}$ ) to nitrate ( $\text{NO}_3\text{-N}$ ). In addition there may be organic interferences on soils high in organic matter ( $> 6.0\%$ ). Care must be taken to avoid contamination from filter paper and handling. Soil nitrate-nitrogen can be used to predict plant response to nitrogen fertilizers. The method detection limit is approximately  $1.0 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible  $\pm 6\%$ .

### Equipment

1. Analytical balance: 100.0 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Repipette dispenser(s), calibrated to  $25.0 \pm 0.2$ ,  $2.5 \pm 0.05$ , and  $2.5 \pm 0.05 \text{ mL}$ .
3. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute (opm).
4. Extraction vessels and associated filtration vessel.
5. S&S #597 filter paper or equivalent highly retentive filter paper.
6. Water bath  $5^\circ\text{C}$ .
7. Vortex stirrer.
8. Spectrophotometer instrument.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Calcium Sulfate Extracting solution, 0.1  $\text{N}$   $\text{CaSO}_4$ : Dissolve 40 g of reagent grade  $\text{CaSO}_4$  in 9 L deionized water and dilute to 10 L final volume (See Comment #1).
3. Antimony Interference Solution: In a 2000 mL flask add 1600 mL  $\text{H}_2\text{SO}_4$  reagent grade. Using heat slowly dissolve 6.0 g of 100 mesh metal antimony (Sb) powder and stir until dissolved. Cool and bring to 2000 mL final volume with deionized water. Solution will require heating to  $50^\circ\text{C}$ , to redissolve antimony prior to usage.
4. Chromotropic Acid solution: Dissolve 0.4 g chromotropic acid (disodium salt) in 800 mL of  $\text{H}_2\text{SO}_4$ , mix and bring to 2000 mL final volume with  $\text{H}_2\text{SO}_4$ . Store in opaque glass bottle.
5. Urea-Sulfite solution: Add 5.0 g of urea to 4.0 gm sodium sulfite to 50 mL deionized water and bring to 100 mL final volume.
6. Standard calibration solutions of  $\text{NO}_3\text{-N}$ . Prepare six calibration standards ranging from 1.0 to 20.0  $\text{mg L}^{-1}$  concentration, diluted in calcium sulfate extraction solution prepared from 1000  $\text{mg L}^{-1}$   $\text{NO}_3\text{-N}$  standard solution.

### Procedure

1. Weigh  $5.0 \pm 0.05 \text{ g}$  of air-dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into extraction vessel. Add 25.0 mL of 0.10  $\text{N}$   $\text{CaSO}_4$  extraction reagent using repipette dispenser (See Comment #2). Include a method blank.
2. Place extraction vessel(s) on reciprocating mechanical shaker for fifteen (15) minutes.
3. Filter extract (See Comment #3), refilter if cloudy. Extract must be clear prior to analysis.
4. Place 1.0 mL of extract in test tube and place test tube in cooled water bath. Concurrently prepare  $\text{NO}_3\text{-N}$  standards with unknown samples.
5. Add 200  $\mu\text{L}$  of urea-sulfite solution, and vortex mix.
6. Add 1.0 mL antimony solution, vortex stir and stand for one (1) hour in a  $5^\circ\text{C}$  water bath.

7. Add 2.5 mL of chromotropic acid solution, vortex stir and under cover for thirty (30) minutes.
8. Determine absorbance on spectrophotometer at 420 nm wavelength. Calibrate using standard calibration solutions and operate instrument in accordance with manufacturer instructions. Determine nitrate concentration of calibration standards, method blank, unknown samples and record results as mg L<sup>-1</sup> of nitrate in extract solution (See Comment #4 and #5).

### Calculation

$$\text{NO}_3\text{-N mg kg}^{-1} \text{ in soil} = (\text{NO}_3\text{-N mg L}^{-1} \text{ in filtrate} - \text{method blank}) \times 5$$

Report soil nitrate concentration to the nearest 0.1 mg kg<sup>-1</sup> (See Comment #6)

### Comments

1. Resuspend to remove undissolved CaSO<sub>4</sub>. As an alternative 20g/L of Ca(OH)<sub>2</sub> can be substituted for CaSO<sub>4</sub>.
2. Check repipette dispensing volume calibration using an analytical balance.
3. Check filter paper supply for possible contamination of and NO<sub>3</sub>-N. If contamination is greater than 0.2 mg L<sup>-1</sup> on a soil extract basis, rinse filter paper with CaSO<sub>4</sub> extraction solution.
4. Samples having nitrate concentrations exceeding the highest standard will require dilution and reanalysis.
5. Soil extracts contains antimony are classified as a hazardous waste and must be disposed of in a suitable manner.
6. Nitrate-nitrogen (NO<sub>3</sub>-N) results can be expressed on a volume basis. Assuming the sample represents a 0-6 inch (0-15 cm) depth of the soil, then: NO<sub>3</sub>-N mg kg<sup>-1</sup> × 2.0 ≅ NO<sub>3</sub>-N lbs ac<sup>-1</sup>

### Literature

Dahnke, W.C. 1990. Testing soils for available nitrogen. p. 120-140. *In*: R.L. Westerman (ed.) Soil testing and plant analysis. Soil Sci. Soc. Am. Book series 3 ASA Madison WI.

Keeney, D.R. and D.W. Nelson. 1982. Nitrogen - inorganic forms. *In* A.L. Page (eds.) Methods of soil analysis, part 2. Agron. Monogr. 9, 2nd ed. ASA and SSSA, Madison, WI. p. 643-698.

Kowalenko, C.G. and L.E. Lowe. 1973. Determination of nitrates in soil extracts. Soil Sci. Soc. Am. Proc. 37:660.

Sims, J.R. and G.D. Jackson. 1971. Rapid analysis of soil nitrate with chromotropic acid. Soil Sci. Soc. Am. Proc. 35:603-606.

## Scope and Application

This method involves the semiquantitative extraction of ammonium ( $\text{NH}_4\text{-N}$ ) from soils using 2.0 N KCl. Ammonium is determined by spectrophotometric, diffusion-conductivity instruments or distillation techniques. The method doesn't quantitatively extract ammonium from mineral structures (i.e. nonexchangeable  $\text{NH}_4\text{-N}$ ) or bound to organic compounds. The method is readily adapted to manual or automated techniques. The procedure outlined follows that outlined by Keeney and Nelson (1982) for determining nitrate nitrogen with a modification in which 25 mL of KCl and 5.0 g of soil are used instead of 100 mL and 10 g soil. Care must be taken to avoid contamination from filter paper and operator handling. Soil ammonium concentrations are generally low in mineral soils ( $< 10 \text{ mg kg}^{-1}$ ). The method detection limit is approximately  $0.2 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible  $\pm 7\%$ .

## Equipment

1. Analytical balance: 100.0 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Repipette dispenser, calibrated to  $25.0 \pm 0.2 \text{ mL}$ .
3. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute (opm).
4. Extraction vessels and associated filtration vessel.
5. Whatman No. 42 or equivalent highly retentive filter paper.
6. Spectrophotometer, or flow injection analyzer (FIA), or distillation instruments.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Potassium chloride extracting solution, 2.0 N KCl: Dissolve 150 g of reagent grade KCl in 500 mL deionized water and dilute to a 1000 mL (See Comment #1).
3. Standard calibration solutions of  $\text{NH}_4\text{-N}$ . Prepare six calibration standards ranging from 0.1 to 20.0  $\text{mg L}^{-1}$  concentration, diluted in 2.0 N KCl extraction solution prepared from 1000  $\text{mg L}^{-1}$   $\text{NO}_3\text{-N}$  standard solution.

## Procedure

1. Weigh  $5.0 \pm 0.05 \text{ g}$  of air-dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into extraction vessel. Add 25.0 mL of 2.0 N KCl extraction reagent using repipette dispenser (See Comment #2). Include a method blank.
2. Place extraction vessel(s) on reciprocating mechanical shaker for thirty (30) minutes.
3. Filter extract (See Comment #3), refilter if filtrate is cloudy (comment #4).
4. Ammonium-N content of the extract is determined using a spectrophotometer, diffusion-conductivity instruments or distillation techniques using standard calibration solutions (See Comment #4 and #5). The ammonium nitrogen content of the digest solution can be determined with a rapid flow analyzer (Technicon Method No. 334-74A/A) or an flow injection analyzer (FIA). This determination can also be made using the Kjeldahl distillation method. Adjust and operate instruments in accordance with manufacturer's instructions. Determine ammonium concentration of a method blank and unknown samples.

## Calculation

$$\text{NH}_4\text{-N mg kg}^{-1} \text{ in soil} = (\text{NH}_4\text{-N mg L}^{-1} \text{ in filtrate} - \text{method blank}) \times 5$$

Report soil aluminum concentration to the nearest 0.1 mg kg<sup>-1</sup> (See Comment #6)

## Comments

1. Soils may be extracted with 2.0 N KCl for the simultaneous determination of nitrate (Method 3.10).
2. Check repipette dispensing volume calibration using an analytical balance.
3. Check filter paper supply for possible contamination of and NH<sub>4</sub>-N. If contamination is greater than 0.2 mg L<sup>-1</sup> on a solution basis, rinse filter paper with 2.0 N KCl.
4. It is recommended that soils extracted for aluminum be analyzed with in two (2) hours after extraction.
5. Samples having ammonium concentrations exceeding the highest standard will require dilution and reanalysis.
6. Ammonium-nitrogen (NH<sub>4</sub>-N) results can be expressed on a volume basis. Assuming the sample represents a 0-6 inch (0-15 cm) depth of the soil, then: NH<sub>4</sub>-N mg kg<sup>-1</sup> × 2.0 ≅ NH<sub>4</sub>-N lbs ac<sup>-1</sup>

## Literature

Bremner, J. M. and D.R. Keeney. 1965. Determination and isotopic ratio analysis of different forms of nitrogen in soils: I. Apparatus and procedure for distillation for and determination of ammonium. *Soil Sci. Soc. Am. Proc.* 29:504-507.

Dahnke, W.C. 1990. Testing soils for available nitrogen. p. 120-140. *In*: R.L. Westerman (ed.) *Soil testing and plant analysis*. Soil Sci. Soc. Am. Book series 3 ASA Madison WI.

Keeney, D.R. and D.W. Nelson. 1982. Nitrogen - inorganic forms. *In* A.L. Page (eds.) *Methods of soil analysis, part 2*. Agron. Monogr. 9, 2nd ed. ASA and SSSA, Madison, WI. p. 643-698.

## Scope and Application

This method involves the semi quantitative extraction of minerizable ammonium nitrogen based on an anaerobic incubation as described by Keeney (1982). Minerizable nitrogen is determined by spectrophotometric, diffusion-conductivity instruments or distillation techniques. Mineralizable nitrogen is used to predict available soil nitrogen for plant response to nitrogen fertilizers. The method detection limit is approximately 0.2 mg kg<sup>-1</sup> (on a dry soil basis) and is generally reproducible  $\pm$  7%.

## Equipment

1. Analytical balance: 100.0 g capacity, resolution  $\pm$  0.01 g.
2. Repipette dispenser, calibrated to 25.0  $\pm$  0.2 mL.
3. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute (opm).
4. Extraction vessels 125 mL and associated filtration vessel.
5. Whatman No. 42 or equivalent highly retentive filter paper.
6. Spectrophotometer, or flow injection analyzer (FIA), or distillation instruments.
7. Incubator capable of 40  $\pm$  0.5 °C.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Potassium chloride extracting solution, 2.0 N KCl: Dissolve 150 g of reagent grade KCl in 500 mL deionized water and dilute to a 1000 mL.
3. Standard calibration solutions of NH<sub>4</sub>-N. Prepare six calibration standards ranging from 0.1 to 20.0 mg L<sup>-1</sup> concentration, diluted in 2.0 N KCl extraction solution prepared from 1000 mg L<sup>-1</sup> NO<sub>3</sub>-N standard solution.

## Procedure

1. Weigh 20.0  $\pm$  0.05 g of air-dried soil pulverized to pass 10 mesh sieve (< 2.0 mm) into 125 mL extraction vessel. Add 25.0 mL of deionized water using repipette dispenser, stir well with a glass rod and add a 2<sup>nd</sup> aliquot of 25.0 mL of deionized water (See Comment #1 and #2).
2. Cover mouth of extraction vessel with parafilm and then plastic and tightly secure with lid.
3. Place extraction vessel(s) in incubator at 40.0 °C for 7 days (168 hours).
4. Remove from incubator and add 50.0 mL of 2.0 N KCl using repipette dispenser.
5. Place extraction vessel(s) on reciprocating mechanical shaker for sixty (60) minutes.
6. Filter extract (See Comment #3), refilter if filtrate is cloudy (See Comment #2).
7. Repeat using Method S-3.50 for extractable NH<sub>4</sub>-N utilizing the same soil(s), reference soil.
8. Ammonium-N content of the extract is determined using a spectrophotometer, diffusion-conductivity instruments or distillation techniques using standard calibration solutions (See Comment #4 and #5). The ammonium nitrogen content of the digest solution can be determined with a rapid flow analyzer (Technicon Method No. 334-74A/A) or an flow injection analyzer (FIA). This determination can also be made using the Kjeldahl distillation method. Adjust and operate instruments in accordance with manufacturer's instructions. Determine ammonium concentration of a method blank and unknown samples.

## Calculation

Mineralizable NH<sub>4</sub>-N mg kg<sup>-1</sup> in soil = (NH<sub>4</sub>-N mg L<sup>-1</sup> incubated - reference extract)  $\times$  5  
Report mineralizable soil nitrogen concentration to the nearest 0.1 mg kg<sup>-1</sup> (See Comment #6)

## Comments

1. It is recommended that samples be rapidly air dried at ambient temperature immediately after sampling to minimize mineralization to nitrate.
2. Check repipette dispensing volume calibration using an analytical balance.
3. Check filter paper supply for possible contamination of and  $\text{NH}_4\text{-N}$ . If contamination is greater than  $0.2 \text{ mg L}^{-1}$  on a solution basis, rinse filter paper with  $2.0 \text{ N KCl}$ .
4. Samples having ammonium concentrations exceeding the highest standard will require dilution and reanalysis.
5. It is recommended that soils extracted for ammonium be analyzed within two (2) hours after extraction.
6. Mineralizable N ( $\text{NH}_4\text{-N}$ ) results can be expressed on a volume basis. Assuming the sample represents a 0-6 inch (0-15 cm) depth of the soil, then:  $\text{NH}_4\text{-N mg kg}^{-1} \times 2.0 \cong \text{NH}_4\text{-N lbs ac}^{-1}$

## Literature

Dahnke, W.C. 1990. Testing soils for available nitrogen. p. 120-140. *In*: R.L. Westerman (ed.) Soil testing and plant analysis. Soil Sci. Soc. Am. Book series 3 ASA Madison WI.

Keeney, D.R. and D.W. Nelson. 1982. Nitrogen - inorganic forms. *In* A.L. Page (eds.) Methods of soil analysis, part 2. Agron. Monogr. 9, 2nd ed. ASA and SSSA, Madison, WI. p. 643-698.

Kowalenko, C.G. and L.E. Lowe. 1973. Determination of nitrates in soil extracts. Soil Sci. Soc. Am. Proc. 37:660.

Sims, J.R. and G.D. Jackson. 1971. Rapid analysis of soil nitrate with chromotropic acid. Soil Sci. Soc. Am. Proc. 35:603-606.

## Scope and Application

This method estimates the relative bioavailability of ortho-phosphate ( $\text{PO}_4\text{-P}$ ) using 0.5 N  $\text{NaHCO}_3$  adjusted to pH 8.50 for soils mildly acidic to alkaline pH and is based on the method developed by Olsen et al., 1954. In the process of extraction,  $\text{CO}_2$  from bicarbonate is driven off, pH increases and bicarbonate converts to carbonate. Thus there is lower calcium activity as calcium carbonate is formed increasing the quantity of phosphates in solution. Phosphorus content is determined spectrophotometrically at 882 nm at an acidity of 0.24 M  $\text{H}_2\text{SO}_4$  (Rodriguez et al., 1994) by reacting with ammonium molybdate using ascorbic acid as a reductant in the presence of antimony (Murphy and Riley, 1962) using manual or automated techniques. The method has shown to be well correlated to crop response phosphorus fertilization on neutral to alkaline soils. In the Pacific Northwest and in the Northern Great Plains the method is used for the simultaneous extraction of plant available potassium, nitrate and specific cases sulfur. The method has a phosphorus detection limit of approximately  $2.0 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible to within  $\pm 12\%$ .

## Equipment

1. Analytical balance, resolution  $\pm 0.01 \text{ g}$ .
2. Oscillating mechanical shaker, capable of 180 oscillations per minute (opm).
3. Repipette dispenser, calibrated to  $40.0 \pm 0.4 \text{ mL}$ ,  $9.0 \pm 0.1 \text{ mL}$ .
4. 125-mL plastic extraction erlenmeyer and associated filtration labware.
5. Whatman No. 1, No. 2, filter paper or equivalent.
6. Pipettes:  $0.250 \pm 0.005 \text{ mL}$ ,  $0.500 \pm 0.005 \text{ mL}$ ,  $1.00 \pm 0.01 \text{ mL}$ ,  $2.00 \pm 0.02 \text{ mL}$ ,  $3.00 \pm 0.03 \text{ mL}$ ,  $4.00 \pm 0.04 \text{ mL}$ .
7. Spectrophotometer wavelength 882 nm and 2.5 cm matching spectrophotometer cuvette or automated Flow Injection Analysis system instrumentation.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Sodium bicarbonate extracting solution (0.5 N  $\text{NaHCO}_3$  @ pH - 8.50). Dissolve 42.01 g of  $\text{NaHCO}_3$  in about 900 mL of deionized water. Adjust the pH to  $8.50 \pm 0.05$  with 2.0 N  $\text{NaOH}$  before diluting with deionized water to 1,000 mL. This solution is unstable with regard to pH and should be prepared as required (See Comments #1).
3. Modified Reagent A (Watanabe and Olsen, 1965).
  - Ammonium Molybdate: Dissolve 12.0 g. of A.R.  $[(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}]$  in 250 mL of deionized water.
  - Antimony Potassium Tartrate: Dissolve 0.291 g. of A.R. antimony potassium tartrate  $[\text{K}(\text{SbO}) \cdot \text{C}_4\text{H}_4\text{O}_6 \cdot \frac{1}{2} \text{H}_2\text{O}]$  in 100 mL of deionized water.
  - Add both of the dissolved reagents to 1,000 mL of 5.76 N  $\text{H}_2\text{SO}_4$  (160 mL of concentrated sulfuric acid per liter, Self and Rodriguez, 1996) mix thoroughly and make to 2,000 mL. Modified Reagent A (mixed reagent) last at least four months if it is stored in an opaque plastic bottle.
4. Reagent B, ascorbic/molybdate reagent (Watanabe and Olsen, 1965). Dissolve 1.32 g. of A.R. ascorbic acid ( $\text{C}_6\text{H}_4\text{O}_6$ ) in 250 mL of modified Reagent A and mix well. This reagent should be prepared as required.
5. Phosphorus Calibrations Standards. From a standard solution containing  $1,000 \text{ mg L}^{-1} \text{PO}_4\text{-P}$ , prepare 100 mL of standard in 0.5N  $\text{NaHCO}_3$  containing  $100 \text{ mg L}^{-1} \text{PO}_4\text{-P}$ . Then, using the  $100 \text{ mg L}^{-1} \text{PO}_4\text{-P}$  standard, prepare seven calibration solutions of 100 mL each in 0.5 N  $\text{NaHCO}_3$  with  $\text{PO}_4\text{-P}$  concentrations of 0.00, 0.25, 0.50, 0.75, 1.00, 2.00, and  $4.00 \text{ mg L}^{-1}$ .



## Extracting Procedure

1. Weigh  $2.00 \pm 0.02$  g of air dried soil pulverized to pass 10 mesh sieve ( $< 2.0$  mm) in a 125-mL plastic extraction erlenmeyer.
2. Add 40.0 mL of 0.5 N  $\text{NaHCO}_3$  extraction solution (See Comments #2, #3 and #4). Include a method blank and standard quality control samples.
3. Place extraction vessels on oscillating mechanical shaker for thirty (30) minutes.
4. Filter suspension immediately - within 1 minute (refilter if filtrate is cloudy).

## Phosphorus Analysis

1. Pipette a 3.0 mL aliquot of standard or soil extract (See Comment #5) into a 2.5 cm matching spectrometer tube.
2. Add 9.0 mL of deionized water.
3. Add 3.0 mL of Reagent B (ascorbic/molybdate reagent).
4. Adjust and operate spectrophotometer in accordance with manufacture's instructions. Read absorbance at a wavelength of 882 nm after 10 minutes of adding the Reagent B. Adjust the 0.000 absorbance using the 0.00 standard. Determine absorbance of a method blank, standards and unknown samples. Calculate phosphorus concentration for blank and unknown samples from standard curve and record phosphorus to the nearest  $0.01 \text{ mg L}^{-1}$   $\text{PO}_4\text{-P}$  in extract solution (See Comments #6 - #8).

## Calculations

Report soil bicarbonate available phosphorus to the nearest  $0.1 \text{ mg kg}^{-1}$  (See Comment #9):

$$\text{Soil PO}_4\text{-P mg kg}^{-1} = (\text{PO}_4\text{-P mg L}^{-1} \text{ in extract} - \text{blank}) \times 20$$

## Comments

1. Storage of 0.5N  $\text{NaHCO}_3$  solution can result in the reagent becoming more alkaline. Increased alkalinity of the extraction solution results in an increase of inorganic phosphorus (Olsen et al., 1954 and Cowling et al., 1987).
2. Clean all extraction and filtration labware with 0.5 N  $\text{HCl}$  and three deionized water rinses to removed potential ortho-phosphate contamination.
3. Check repipette dispensing volume, calibrate using an analytical balance.
4. Extractable phosphorus increases 0.43 expressed as  $\text{mg kg}^{-1}$ , for each degree rise in temperature between 20 and 30 °C for soils between 5 and 40  $\text{mg kg}^{-1}$  (Olsen et al., 1954).
5. For automated FIA analysis sample extracts will require neutralization of  $\text{NaHCO}_3$  and degassing to remove dissolved  $\text{CO}_2$  prior to analysis. Specific FIA instruments have capability to remove dissolved  $\text{CO}_2$ .
6. Samples exceeding highest standard will require dilution with the extracting solution for reanalysis.
7. Potassium in the extract is can be measured by flame emission spectrophotometry by either manual or an automated system. The  $\text{NaHCO}_3$  extract is combined with acidified lithium nitrate, degassed and analyzed by flame emission spectrophotometry (Schoenau and Karamanos, 1993).
8. Nitrate ( $\text{NO}_3\text{-N}$ ) may also be determined directly from the Olsen extract. Extracts will require neutralization of  $\text{NaHCO}_3$  and degassing to remove dissolved  $\text{CO}_2$  prior to analysis.
9. Generally, soils having a bicarbonate available phosphorus level below  $20 \text{ mg kg}^{-1}$  will have a response to applications of phosphorus fertilizers for most crops.

## Literature

Horneck, D. A., J. M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. Ag. Expt. Station SM 89:4.

Cowling, J.C., T.W. Speir and H.J. Percival. 1987. Potential problems with the determination of Olsen and microbial P of soils due to the instability of 0.5 M sodium bicarbonate. Commun. in Soil Sci. and Plant Anal. 18(6):637-652.

Knudsen, D. and D. Beegle. 1988. Recommended Phosphorus Tests. p. 12-14. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bull. No. 499 (revised).

Murphy, J. and J.P. Riley. 1962. A modified single solution method for determination of phosphates in natural waters. Anal. Chim. Acta 27:31-36.

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Olsen, S.R. and L.E. Sommers. 1982. Phosphorus. p 403-430. *In*: A.L. Page et al. (eds.) Methods of soil analysis, part 2. Agron. Mongr. 9. 2nd ed. ASA and SSSA, Madison, WI.

Rodriguez, J.B., J.R. Self, and P.N. Soltanpour. 1994. Optimal conditions for phosphorus analysis by the ascorbic acid-molybdenum blue method. Soil Sci.Soc. Am. J. 58:866-870.

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Watanabe, F.S., and S.R. Olsen. 1965. Test of an ascorbic acid method for determining phosphorus in water and NaHCO<sub>3</sub> extracts from soils. Soil Sci. Soc. Am. Proc. 29:677-678.

## Scope and Application

This method estimates the relative bioavailability of ortho-phosphate ( $\text{PO}_4\text{-P}$ ) in acid to neutral pH soils using a dilute acid solution of pH - 2.60 that is 0.025 M HCl and 0.03 M  $\text{NH}_4\text{F}$  based on a 1:7 extraction soil ratio. In the process of extraction, phosphorus is solubilize under two different mechanisms, the strong acid increases the solubility of phosphates by protonation, whereas the fluoride lower the activity of calcium as calcium fluoride increasing the quantity of phosphates in solution. Phosphorus content is determined spectrophotometrically at 882 nm at an acidity of 0.19 M  $\text{H}_2\text{SO}_4$  (Rodriguez et al., 1994) by reacting with ammonium molybdate using ascorbic acid as a reductant in the presence of antimony (Murphy and Riley, 1962). The method is shown to be well correlated to crop response to phosphorus fertilization on neutral to acid soils. The method has a phosphorus detection limit of approximately  $2.0 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible within  $\pm 10\%$ .

## Equipment

1. Analytical balance: resolution 0.01 g.
2. Reciprocating mechanical shaker, capable of 180 oscillations per minute (opm).
3. Repipette dispenser, calibrated to  $20.0 \pm 0.2 \text{ mL}$ ,  $12 \pm 0.1 \text{ mL}$ .
4. 125-mL plastic extraction erlenmeyer and associated filtration vessel.
5. Whatman No. 1 filter paper or equivalent.
6. Pipettes:  $0.250 \pm 0.005 \text{ mL}$ ,  $0.500 \pm 0.005 \text{ mL}$ ,  $1.00 \pm 0.1 \text{ mL}$ ,  $2.00 \pm 0.02 \text{ mL}$ ,  $3.00 \pm 0.03 \text{ mL}$ .
7. 2.5 cm matching spectrophotometer cuvette.
8. Spectrophotometer, wavelength 882 nm or automated Flow Injection Analysis system.

## Reagents

1. Deionized water, ASTM type I grade.
2. Hydrochloric acid 0.5 N HCL: Dilute 103 mL of concentrated HCL to a volume of 2,500 mL with deionized water.
3. Ammonium fluoride stock solution, 1.0 N  $\text{NH}_4\text{F}$ : dissolve 74.0 g of  $\text{NH}_4\text{F}$  in deionized water and dilute the solution to 2,000 mL.
4. Bray and Kurtz P-1 extracting solution (0.025 N HCl-0.03 N  $\text{NH}_4\text{F}$  @ pH - 2.60): Mix thoroughly 1,000 mL of 0.5 N HCl and 600 mL of 1.0 N  $\text{NH}_4\text{F}$  with about 18.0 L of deionized water. Adjust the pH to  $2.60 \pm 0.05$  with diluted HCl or  $\text{NH}_4\text{OH}$  before dilution to 20.0 L. This produces a solution of 0.03 N  $\text{NH}_4\text{F}$  and 0.025 N HCL. Store the solution in a polyethylene container. Check the pH before use.
5. Modified Reagent A (Watanabe and Olsen, 1965).
  - Ammonium Molybdate: Dissolve 12.0 g. of A.R.  $[(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}]$  in 250 mL of deionized water.
  - Antimony Potassium Tartrate: Dissolve 0.291 g. of A.R. antimony potassium tartrate  $[\text{K}(\text{SbO}) \cdot \text{C}_4\text{H}_4\text{O}_6 \cdot \frac{1}{2} \text{H}_2\text{O}]$  in 100 mL of deionized water.
  - Add both of the dissolved reagents to 1,000 mL of 5.76 N  $\text{H}_2\text{SO}_4$  (160 mL of concentrated sulfuric acid per liter, Self and Rodriguez, 1996) mix thoroughly and make to 2,000 mL. Modified Reagent A (mixed reagent) last at least four months if it is stored in an opaque plastic bottle.
6. Reagent B, ascorbic/molybdate reagent (Watanabe and Olsen, 1965). Dissolve 1.32 g. of A.R. ascorbic acid ( $\text{C}_6\text{H}_4\text{O}_6$ ) in 250 mL of modified Reagent A and mix well. This reagent should be prepared as required. Modified Reagent A (Watanabe and Olsen, 1965).
7. Phosphorus Calibrations Standards. From a standard solution containing  $1,000 \text{ mg L}^{-1} \text{PO}_4\text{-P}$  prepare 100 mL of standard in Bray and Kurtz P-1 extracting solution containing  $100 \text{ mg L}^{-1} \text{PO}_4\text{-P}$ . Then, using the  $100 \text{ mg L}^{-1} \text{PO}_4\text{-P}$  standard, prepare six calibration solutions of 100 mL each in Bray and Kurtz P-1 extracting solution with  $\text{PO}_4\text{-P}$  concentrations of 0.00, 0.75, 1.50, 3.00, 6.00, and  $12.00 \text{ mg L}^{-1}$ .

## Extracting Procedure

1. Weigh  $2.00 \pm 0.02$  g of air dried soil pulverized to pass 10 mesh sieve ( $< 2.0$  mm) in a 125-mL plastic extraction erlenmeyer.
2. Add 14.0 mL of Bray and Kurtz P-1 extracting solution (See Comment #2). Include a method blank.
3. Place extraction vessels on reciprocating mechanical shaker for 5 min.
4. Filter suspension immediately, refilter if filtrate is cloudy.

## Phosphorus Analysis

1. Pipette 1.0 mL aliquot of standard or soil extract into a 2.5 cm matching spectrometer tube (See Comment #3).
2. Add 12.0 mL of deionized water.
3. Add 2.0 mL of Reagent B (ascorbic/molybdate reagent developing reagent).
4. Adjust and operate spectrophotometer in accordance with manufacture's instructions. Read absorbance at a wavelength of 882 nm after ten (10) minutes of adding the Reagent B. Adjust the 0.000 absorbance using the 0.00 standard. Determine absorbance of a method blank, standards and unknown samples. Calculate phosphorus concentration for blank and unknown samples from standard curve and record phosphorus to the nearest  $0.01 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$  in extract solution (See Comment #4).

## Calculations

Report soil Bray and Kurtz P-1 available phosphorus to the nearest  $0.1 \text{ mg kg}^{-1}$  :

$$\text{soil PO}_4\text{-P mg kg}^{-1} = (\text{PO}_4\text{-P mg L}^{-1} \text{ in extract} - \text{method blank}) \times 10$$

(See Comment #5).

## Comments

1. This method for available P follows the procedure originally outlined by Olsen and Summers (1982). The original Bray-P1 method describes a soil extractant ratio of 1:7 and an extraction time of 60 seconds. To simplify the method a number of labs in the east and North Central United States have altered the procedure to a soil extractant ratio of 1:10 and an extraction time of five minutes.
2. Check repipette dispensing volume, calibrate using an analytical balance.
3. Extracts can be analyzed directly for phosphorus by automated FIA analysis.
4. Samples exceeding highest standard will require dilution with the extracting solution for reanalysis.
5. Generally, soils having less than  $25 \text{ mg kg}^{-1}$  will have a response to applications of phosphorus fertilizers for most crops.

## Literature

Bray, R.H. and L.T. Kurtz 1945. Determination of total, organic and available phosphorus in soil. *Soil Sci.* 59:39-45.

Murphy, J., and J.P. Riley. 1962. A modified single solution method for determination of phosphates in natural waters. *Anal. Chim. Acta* 27:31-36.

Olsen, S.R. and L.E. Sommers. 1982. Phosphorus. p 416-418. *In*: A.L. Page et al. (eds.) Methods of soil analysis, part 2. Agron. Mongr. 9. 2nd ed. ASA and SSSA, Madison, WI.

Rodriguez, J.B., J.R. Self, and P.N. Soltanpour. 1994. Optimal conditions for phosphorus analysis by the ascorbic acid-molybdenum blue method. *Soil Sci.Soc. Am. J.* 58:866-870.

Self, J.R. and J.B. Rodriguez. 1996. Laboratory manual for soil and plant chemical analysis. Soil, Water, and Plant Testing Laboratory, Colorado State University.

Watanabe, F.S., and S.R. Olsen. 1965. Test of an ascorbic acid method for determining phosphorus in water and  $\text{NaHCO}_3$  extracts from soils. *Soil Sci. Soc. Am. Proc.* 29:677-678.

### Scope and Application

This method estimates the relative bioavailability of  $\text{PO}_4\text{-P}$ ,  $\text{NO}_3\text{-N}$ , K, Ca, Mg, Zn, Fe, Mn, and Cu using a 1.0 M  $\text{NH}_4\text{HCO}_3$  - 0.005 M DTPA solution adjusted to pH - 7.60 for soils with neutral to alkaline pH. It is based on the method developed by Soltanpour and Schwab (1977) and is advantageous since it simultaneously extracts macro and micronutrients in a single extract with subsequent analysis using a spectrophotometer and ICP-AS instruments (Soltanpour, 1991). In the process of extraction,  $\text{CO}_2$  from bicarbonate is given off, solution pH increases and bicarbonate converts to carbonate which in turn reduces calcium activity (as calcium carbonate) increasing the quantity of  $\text{PO}_4\text{-P}$  in solution. Phosphorus content is determined spectrophotometrically at 882 nm at an acidity of 0.18 M  $\text{H}_2\text{SO}_4$  (Rodriguez et al., 1994) by reacting with ammonium molybdate using ascorbic acid as a reductant in the presence of antimony (Murphy and Riley, 1962). Nitrates are water soluble and can be analyzed from solution. Ammonium exchanges with potassium, calcium, and magnesium, and the original pH (7.60) of the AB-DTPA allows DTPA to extract and chelate iron, manganese, zinc, and copper and toxic metals. The method has shown to be well correlated to crop response to phosphorus fertilization. The method is well suited for screening mine spoils, sewage sludge amended soils and soils contaminated with potentially toxic elements (ie. As, Se, B, Cd, Ni, Pb, and Mo). The method has a phosphorus detection limit of approximately  $0.50 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible to within  $\pm 7\%$ .

### Equipment

1. Analytical balance: resolution 0.01 g.
2. Reciprocating mechanical shaker, capable of 180 oscillations per minute (opm).
3. Repipette dispenser, calibrated to  $20 \pm 0.2 \text{ mL}$ ,  $12 \pm 0.1 \text{ mL}$ .
4. 125-mL plastic extraction erlenmeyer and associated filtration vessel.
5. Whatman No. 1 filter paper or equivalent.
6. Pipettes:  $0.250 \pm 0.005 \text{ mL}$ ,  $0.500 \pm 0.005 \text{ mL}$ ,  $1.00 \pm 0.01 \text{ mL}$ ,  $2.00 \pm 0.02 \text{ mL}$ ,  $3.00 \pm 0.03 \text{ mL}$ .
7. 2.5 cm matching spectrophotometer cuvette.
8. Spectrophotometer, wavelength 520 and 882 nm, automated FIA and/or ICP-AES.

### Reagents

1. Deionized water, ASTM Type I grade.
2. 0.005 M DTPA solution. Dissolve 19.70 of diethylenetriamine pentaacetic acid (DTPA) in about 9.5 L of deionized water. Since DTPA is not very soluble in water it needs about eight hrs of constant agitation for total dissolution, dilute to 10 L with deionized water. The pH of the DTPA solution is  $2.40 \pm 0.10$  and it is very stable. Store the DTPA solution in a polyethylene container.
3. AB-DTPA extracting solution (1M  $\text{NH}_4\text{HCO}_3$  - 0.005 M DTPA @ pH - 7.60): Using a vortex stirrer, dissolve 79.06 g. of  $\text{NH}_4\text{HCO}_3$  with about 900 mL of 0.005 M DTPA solution, after dissolution of the ammonium bicarbonate dilute to 1.0 L with 0.005 M DTPA, the pH of this solution is about 7.5, adjust the pH to  $7.60 \pm 0.05$  by agitation. This AB-DTPA solution is unstable with regard to pH, prepare only the quantity that is needed on a daily basis (Self and Rodriguez, 1996). (see comment # 1).
4. 5 M NaOH. In a plastic container with about 750 mL of deionized water, carefully dissolve 200 g. of A.R. sodium hydroxide.

5. Modified Reagent A (Watanabe and Olsen, 1965).
  - Ammonium Molybdate: Dissolve 12.0 g. of A.R.  $[(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}]$  in 250 mL of deionized water.
  - Antimony Potassium Tartrate: Dissolve 0.291 g. of A.R. antimony potassium tartrate  $[\text{K}(\text{SbO}) \cdot \text{C}_4\text{H}_4\text{O}_6 \cdot \frac{1}{2} \text{H}_2\text{O}]$  in 100 mL of deionized water.
  - Add both of the dissolved reagents to 1,000 mL of 5.76 N  $\text{H}_2\text{SO}_4$  (160 mL of concentrated sulfuric acid per liter, Self and Rodriguez, 1996) mix thoroughly and make to 2,000 mL. Modified Reagent A (mixed reagent) last at least four months if it is stored in an opaque plastic bottle.
6. Reagent B, ascorbic/molybdate reagent (Watanabe and Olsen, 1965). Dissolve 1.32 g. of A.R. ascorbic acid ( $\text{C}_6\text{H}_4\text{O}_6$ ) in 250 mL of modified Reagent A and mix well. This reagent should be prepared as required.
7. Phosphorus Calibrations Standards. From a standard solution containing  $1,000 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$ , prepare 100 mL of standard in AB-DTPA extracting solution containing  $100 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$ . Then, using the  $100 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$  standard, prepare seven calibration solutions of 100 mL each in AB-DTPA extracting solution with  $\text{PO}_4\text{-P}$  concentrations of 0.00, 2.5, 5.0, 10.0, 20.0, and  $40.0 \text{ mg L}^{-1}$ .

### Extracting Procedure

1. Weigh  $10.0 \pm 0.1 \text{ g}$  of air dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) in a 125-mL plastic extraction erlenmeyer.
2. Add 20.0 mL of AB-DTPA extracting solution. Include a method blank (See Comment # 2 and # 3).
3. Place extraction vessel on mechanical shaker for fifteen (15) minutes.
4. Filter suspension immediately, refilter if filtrate is cloudy.

### Phosphorus Chemical Analysis

1. Pipette 0.250 mL aliquot of standard or soil extract into a 2.5 cm matching spectrometer tube (See Comment #4).
2. Add 0.250 mL of 5 M NaOH (Self and Rodriguez, 1996).
3. Let the reaction occurs for 10 min.
4. Add 10.0 mL of deionized water (See Comment #5).
5. Add 2.0 mL of Reagent B (ascorbic/molybdate reagent).
6. Adjust and operate spectrophotometer in accordance with manufacture's instructions. Read absorbance at a wavelength of 882 nm after ten (10) minutes of adding the Reagent B. Adjust the 0.000 absorbance using the 0.00 standard. Determine absorbance of a method blank, standards and unknown samples. Calculate phosphorus concentration for blank and unknown samples from standard curve and record phosphorus to the nearest  $0.01 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$  in extract solution (See Comment #6).

### Calculations

Report soil AB-DTPA extractable phosphorus to the nearest  $0.1 \text{ mg kg}^{-1}$  :

$$\text{soil } \text{PO}_4\text{-P } \text{mg kg}^{-1} = (\text{PO}_4\text{-P } \text{mg L}^{-1} \text{ in extract} - \text{method blank}) \times 2$$

(See Comment # 7).

### Comments

1. pH of the extraction solution is unstable and may become effervescence in automated dispensers.
2. Clean all extraction and filtration labware with 0.5 N HCl and three deionized water rinses to removed potential ortho-phosphate contamination.
3. Cover filtration completely vessels with a plastic sheet to diminish contamination of  $\text{NH}_4\text{-N}$  to surrounding soils and water samples.

4. Extracts can be analyzed directly for phosphorus by automated FIA analysis.
5. Check pipette dispensing volume, calibrate using an analytical balance.
6. Determine nitrate on aliquot of soil extract, method S - 3.10. Determine K, Ca, Mg, Fe, Mn, Zn and Cu by AAS or IAP-AS. The use of Legere teflon nebulizer (Burtec Instrument Corporation, Delmar, New York) facilitates the analysis of solutions high in dissolved solids by IAP-AS analysis.
7. Samples exceeding highest standard will require dilution with AB-DTPA for reanalysis.

## Literature

Murphy, J., and J.P. Riley. 1962. A modified single solution method for determination of phosphates in natural waters. *Anal. Chim. Acta* 27:31-36.

Rodriguez, J.B., J.R. Self, and P.N. Soltanpour. 1994. Optimal conditions for phosphorus analysis by the ascorbic acid-molybdenum blue method. *Soil Sci. Soc. Am. J.* 58:866-870.

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Watanabe, F.S., and S.R. Olsen. 1965. Test of an ascorbic acid method for determining phosphorus in water and  $\text{NaHCO}_3$  extracts from soils. *Soil Sci. Soc. Am. Proc.* 29:677-678.



## Scope and Application

This method estimates the relative bioavailability of ortho-phosphate ( $\text{PO}_4\text{-P}$ ) in acid to neutral soils with cation exchange capacities of less than 20 meq/100 g using a buffer solution of pH 4.80. This method, first proposed by Morgan (1941), was described in detail by Lunt et al. (1950) later by Greweling and Peech (1965). The Morgan extracting reagent is a well buffered solution of 0.52 M  $\text{CH}_3\text{-COOH}$ -0.73 M  $\text{CH}_3\text{-COONa}$  @ pH - 4.80. Phosphorus is determined spectrophotometrically at 882 nm at an acidity of 0.21 M  $\text{H}_2\text{SO}_4$  (Rodriguez et al., 1994) by reacting with ammonium molybdate using ascorbic acid as a reductant in the presence of antimony (Murphy and Riley, 1962). The method is used on soils of Northeast and Pacific Northwest and can be used for K and other cations. The method has a phosphorus detection limit of approximately  $1.0 \text{ mg kg}^{-1} \text{ PO}_4\text{-P}$  (on a dry soil basis) and is generally reproducible within  $\pm 8\%$ .

## Equipment

1. Analytical balance: resolution 0.01 g.
2. Reciprocating mechanical shaker, capable of 180 oscillations per minute (opm).
3. Repipette dispenser, calibrated to  $20 \pm 0.2 \text{ mL}$ ,  $12 \pm 0.1 \text{ mL}$ .
4. 125-mL plastic extraction erlenmeyer and associated filtration vessel.
5. Whatman No. 1 filter paper or equivalent.
6. Pipettes:  $0.250 \pm 0.005 \text{ mL}$ ,  $0.500 \pm 0.005 \text{ mL}$ ,  $1.00 \pm 0.01 \text{ mL}$ ,  $2.00 \pm 0.02 \text{ mL}$ ,  $3.00 \pm 0.03 \text{ mL}$ .
7. 2.5 cm matching spectrophotometer cuvette.
8. Spectrophotometer, wavelength 882 nm or automated Flow Injection Analysis (FIA) system.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Morgan Extracting Reagent ( 0.52 M  $\text{CH}_3\text{-COOH}$  - 0.73 M  $\text{CH}_3\text{-COONa}$  @ pH - 4.80). Dissolve 100 g sodium acetate ( $\text{CH}_3\text{-COONa} \cdot 3\text{H}_2\text{O}$ ) in about 900 mL pure water. Add 30 mL glacial acetic acid ( $\text{CH}_3\text{-COOH}$ ), adjust the pH to  $4.80 \pm 0.05$  with diluted acetic acid or NaOH, and dilute to 1,000 mL with deionized water.
3. Modified Reagent A (Watanabe and Olsen, 1965).
  - Ammonium Molybdate: Dissolve 12.0 g. of A.R.  $[(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}]$  in 250 mL of deionized water.
  - Antimony Potassium Tartrate: Dissolve 0.291 g. of A.R. antimony potassium tartrate  $[\text{K}(\text{SbO}) \cdot \text{C}_4\text{H}_4\text{O}_6 \cdot \frac{1}{2} \text{H}_2\text{O}]$  in 100 mL of deionized water.
  - Add both of the dissolved reagents to 1,000 mL of 5.76 N  $\text{H}_2\text{SO}_4$  (160 mL of concentrated sulfuric acid per liter, Self and Rodriguez, 1996) mix thoroughly and make to 2,000 mL. Modified Reagent A (mixed reagent) last at least four months if it is stored in an opaque plastic bottle.
4. Reagent B, ascorbic/molybdate reagent (Watanabe and Olsen, 1965). Dissolve 1.32 g. of A.R. ascorbic acid ( $\text{C}_6\text{H}_4\text{O}_6$ ) in 250 mL of modified Reagent A and mix well. This reagent should be prepared as required.
5. Phosphorus Calibrations Standards. From a standard solution containing  $1,000 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$ , prepare 100 mL of standard in Morgan extracting solution containing  $100 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$ . Then, using the  $100 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$  standard, prepare six calibration solutions of 100 mL each in 0.5 N Morgan solution with  $\text{PO}_4\text{-P}$  concentrations of 0.00, 0.75, 1.50, 3.00, 6.00, and  $12.00 \text{ mg L}^{-1}$ .

## Extracting Procedure

1. Weigh  $5.00 \pm 0.05 \text{ g}$  of air dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) in a 125 mL extraction vessel.
2. Add 25.0 mL of Morgan extraction solution (See comment #1 and #2). Include a method blank.

3. Place extraction vessel(s) on mechanical shaker for thirty (30) minutes.
4. Filter suspension immediately, refilter if filtrate is cloudy.

### Phosphorus Chemical Analysis

1. Pipette 1.0 mL aliquot of standard or soil extract into a 2.5 cm matching spectrometer tube (see comment # 3).
2. Add 12.0 mL of deionized water (see comment # 4).
3. Add 2.0 mL of Reagent B (ascorbic/molybdate reagent).
4. Adjust and operate spectrophotometer in accordance with manufacture's instructions. Read absorbance at a wavelength of 882 nm after ten (10) minutes of adding the Reagent B. Adjust the 0.000 absorbance using the 0.00 standard. Determine absorbance of a method blank, standards and unknown samples. Calculate phosphorus concentration for blank and unknown samples from standard curve and record phosphorus to the nearest 0.01 mg L<sup>-1</sup> PO<sub>4</sub>-P in extract solution (See Comment #5).

### Calculations

Report modified Morgan extractable soil phosphorus to the nearest 0.1 mg kg<sup>-1</sup> :

soil PO<sub>4</sub>-P mg kg<sup>-1</sup> = (PO<sub>4</sub>-P mg L<sup>-1</sup> in extract - method blank) × 5  
(See Comment # 6).

### Comments

1. Soils may be stored in an air-dry condition for several months with no effect on extractable P.
2. The soil extracts should not be stored for more than 24 hours after extraction.
3. Extracts can be analyzed directly for phosphorus by automated FIA analysis.
4. Check repipette dispensing volume, calibrate using an analytical balance.
5. Samples exceeding highest standard will require dilution with Morgan extractant for reanalysis.
6. Accurate fertilizer recommendations for phosphorus must be based on field response data conducted under local soil-climate-crop conditions. Interpretations will vary, depending on soil characteristics crops and yield potential.

### Literature

- Greweling, T., and M. Peech. 1965. chemical soil tests. Cornell Agr. Exp. Sta. Bull. 960. Jackson, M.L. 1958. Soil Chemical Analysis. Prentice-Hall, Inc., Englewood Cliffs, NJ.
- Lunt, H.A., C.L.W. Swanson, and H.G.M. Jacobson. 1950. The Morgan Soil Testing System. Conn. Agr. Exp. Sta. (New Haven) Sta. Bull. 541.
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- Murphy, J., and J.P. Riley. 1962. A modified single solution method for determination of phosphates in natural waters. Anal. Chim. Acta 27:31-36.
- Rodriguez, J.B., J.R. Self, and P.N. Soltanpour. 1994. Optimal conditions for phosphorus analysis by the ascorbic acid-molybdenum blue method. Soil Sci.Soc. Am. J. 58:866-870.
- Self, J.R. and J.B. Rodriguez. 1996. Laboratory manual for soil and plant chemical analysis. Soil, Water, and Plant Testing Laboratory, Colorado State University.

Watanabe, F.S., and S.R. Olsen. 1965. Test of an ascorbic acid method for determining phosphorus in water and NaHCO<sub>3</sub> extracts from soil. *Soil Sci. Soc. Amer. Proc.* 29:677-678.

## Scope and Application

This method estimates the relative bioavailability of ortho-phosphate (PO<sub>4</sub>-P) in mildly acid to alkaline soils. This method was developed in Alberta Canada by Ashworth and Mrazek (1989) and is a modification of the procedure described by van Lierop, 1988. The original bases of the method is on the work of Bray and Kurtz (1945). The Modified Kewlona extracting reagent is a well buffered solution of ( 0.015 N NH<sub>4</sub>F - 0.50 M CH<sub>3</sub>-COOH - 1.0 M CH<sub>3</sub>-COO NH<sub>4</sub> . Phosphorus is determined spectrophotometrically at 882 nm at an acidity of 0.19 M H<sub>2</sub>SO<sub>4</sub> (Rodriguez et al., 1994) by reacting with ammonium molybdate using ascorbic acid as a reductant in the presence of antimony (Murphy and Riley, 1962). The method is used on soils of Western Canada and the Northern Rocky Mountain States and can be used for the simultaneous determination of K . The method has a phosphorus detection limit of approximately 1.0 mg kg<sup>-1</sup> (on a dry soil basis) and is generally reproducible within ±10%.

## Equipment

1. Analytical balance 100.0 resolution 0.01 g.
2. Reciprocating mechanical shaker, capable of 180 oscillations per minute (opm).
3. Repipette dispenser, calibrated to, 12.0 ± 0.1 mL, 50.0 ± 0.2 mL.
4. 125-mL plastic extraction erlenmeyer and associated filtration vessel.
5. Whatman No. 1 filter paper or equivalent.
6. Pipettes:, 1.00 ±0.01 mL, 2.00 ±0.02 mL.
7. 2.5 cm matching spectrophotometer cuvette.
8. Spectrophotometer, wavelength 882 nm or automated Flow Injection Analysis (FIA) system.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Modified Kewlona extracting reagent ( 0.015 N NH<sub>4</sub>F - 0.50 M CH<sub>3</sub>-COOH - 1.0 M (CH<sub>3</sub>-COO NH<sub>4</sub>) . Ammonium fluoride stock solution (1.00 N NH<sub>4</sub>F): Dissolve 5.54 g of NH<sub>4</sub>F in 1500 mL of deionized water. Add 192g ammonium Acetate (CH<sub>3</sub>-COO NH<sub>4</sub>) and 143.7 mL Glacial Acetic Acid (CH<sub>3</sub>-COOH) and dilute to 10.0 L final volume.
3. Modified Reagent A (Watanabe and Olsen, 1965).
  - Ammonium Molybdate: Dissolve 12.0 g. of A.R. [(NH<sub>4</sub>)<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub> · 4H<sub>2</sub>O] in 250 mL of deionized water.
  - Antimony Potassium Tartrate: Dissolve 0.291 g. of A.R. antimony potassium tartrate [K(SbO) · C<sub>4</sub>H<sub>4</sub>O<sub>6</sub> · ½ H<sub>2</sub>O] in 100 mL of deionized water.
  - Add both of the dissolved reagents to 1,000 mL of 5.76 N H<sub>2</sub>SO<sub>4</sub> (160 mL of concentrated sulfuric acid per liter, Self and Rodriguez, 1996) mix thoroughly and make to 2,000 mL. Modified Reagent A (mixed reagent) last at least four months if it is stored in an opaque plastic bottle.
4. Reagent B, ascorbic/molybdate reagent (Watanabe and Olsen, 1965). Dissolve 1.32 g. of A.R. ascorbic acid (C<sub>6</sub>H<sub>4</sub>O<sub>6</sub>) in 250 mL of modified Reagent A and mix well. This reagent should be prepared as required.
5. Phosphorus Calibrations Standards. From a standard solution containing 1,000 mg L<sup>-1</sup> PO<sub>4</sub>-P, prepare 100 mL of standard in Morgan extracting solution containing 100 mg L<sup>-1</sup> PO<sub>4</sub>-P. Then, using the 100 mg L<sup>-1</sup> PO<sub>4</sub>-P standard, prepare six calibration solutions of 100 mL each in Modified Kewlona solution with PO<sub>4</sub>-P concentrations of 0.00, 0.75, 1.50, 3.00, 6.00, and 12.00 mg L<sup>-1</sup>.

## Extracting Procedure

1. Weigh 5.00 ± 0.05 g of air dried soil pulverized to pass 10 mesh sieve (< 2.0 mm) in a 125 mL extraction vessel.
2. Add 50.0 mL of Morgan extraction solution (See comment #1 and #2). Include a method blank.

3. Place extraction vessel(s) on a reciprocating mechanical shaker for thirty (30) minutes.
4. Filter suspension immediately, refilter if filtrate is cloudy.

### Phosphorus Chemical Analysis

1. Pipette 1.0 mL aliquot of standard or soil extract into a 2.5 cm matching spectrometer tube (see comment # 3).
2. Add 12.0 mL of deionized water (see comment # 4).
3. Add 2.0 mL of Reagent B (ascorbic/molybdate reagent).
4. Adjust and operate spectrophotometer in accordance with manufacture's instructions. Read absorbance at a wavelength of 882 nm after ten (10) minutes of adding the Reagent B. Adjust the 0.000 absorbance using the 0.00 standard. Determine absorbance of a method blank, standards and unknown samples. Calculate phosphorus concentration for blank and unknown samples from standard curve and record phosphorus to the nearest 0.01 mg L<sup>-1</sup> PO<sub>4</sub>-P in extract solution (See Comment #5).

### Calculations

Report Modified Kewlona extractable soil phosphorus to the nearest 0.1 mg kg<sup>-1</sup> :

$$\text{soil PO}_4\text{-P mg kg}^{-1} = (\text{PO}_4\text{-P mg L}^{-1} \text{ in extract} - \text{method blank}) \times 10 \quad (\text{See Comment \# 6}).$$

### Comments

1. Soils may be stored in an air-dry condition for several months with no effect on extractable P.
2. The soil extracts should not be stored for more than 24 hours after extraction.
3. Extracts can be analyzed directly for phosphorus by automated FIA analysis.
4. Check repipette dispensing volume, calibrate using an analytical balance.
5. Samples exceeding highest standard will require dilution with Morgan extractant for reanalysis.
6. Accurate fertilizer recommendations for phosphorus must be based on field response data conducted under local soil-climate-crop conditions. Interpretations will vary, depending on soil characteristics crops and yield potential. Potassium extracted by the Modified kewlona extract is equivalent to that extracted by the ammonium acetate method (S - 5.10).

### Literature

Ashworth J. And K. Mrazek. 1989. The acetic-fluoride method for available phosphorus and potassium. Proceedings Alberta Soil Science Workshop 26:178-182.

Bray, R.H. and L.T. Kurtz 1945. Determination of total, organic and available phosphorus in soil. Soil Sci. 59:39-45.

Murphy, J., and J.P. Riley. 1962. A modified single solution method for determination of phosphates in natural waters. Anal. Chim. Acta 27:31-36.

Rodriguez, J.B., J.R. Self, and P.N. Soltanpour. 1994. Optimal conditions for phosphorus analysis by the ascorbic acid-molybdenum blue method. Soil Sci.Soc. Am. J. 58:866-870.

Self, J.R. and J.B. Rodriguez. 1996. Laboratory manual for soil and plant chemical analysis. Soil, Water, and Plant Testing Laboratory, Colorado State University.

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Watanabe, F.S., and S.R. Olsen. 1965. Test of an ascorbic acid method for determining phosphorus in water and NaHCO<sub>3</sub> extracts from soil. *Soil Sci. Soc. Amer. Proc.* 29:677-678.

## Scope and Application

This method semiquantitatively determines the amount of soil plant available K, Ca, Mg, and Na residing on the soil colloid exchange sites by displacement with ammonium acetate solution buffered to pH 7.0. Cation concentrations are determined using atomic emission (AES), absorption spectrometry (AAS) or ICP-AES instrumentation. A chemical interference solution is used to minimize chemical matrix effects. It is based on a modification of the procedure outlined by Knudsen et al. (1982) for exchangeable K. Generally, these cations are associated with the exchange sites. The exception are soils that have high soluble salts and are saline, which requires a special preanalysis treatment. The method doesn't correct for calcium and magnesium extracted as free carbonates or gypsum. In the northern Great Plains the method has been used to determine available sulfur. The method detection limit is approximately of 25 mg kg<sup>-1</sup> (on a dry soil basis) and is generally reproducible ± 7%.

## Equipment

1. Analytical balance: 250 g capacity, resolution ± 0.01 g.
2. Reciprocating horizontal mechanical shaker, 180 oscillations per minute (opm).
3. Repipette dispenser, calibrated to 25.0 ± 0.2 mL.
3. Extraction vessel and associated filtration vessel.
4. Whatman No. 42, No. 2 or equivalent highly retentive filter paper.
5. Atomic Emission/Absorption Spectrophotometer (AAE) (AAS) or ICP-AES.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Ammonium acetate (1.0 N) extraction solution neutral @ pH 7.0: Add 570 mL of glacial acetic acid CH<sub>3</sub>COOH (99%) to 8000 mL of deionized water. Add 680 mL of concentrated ammonium hydroxide adjust pH to 7.0 with 3.0 N glacial acetic acid or 3.0 N ammonium hydroxide and dilute to 10 L final volume. Check solution for possible contamination of K, Na, Ca, and Mg (see comment #1).
3. Chemical interference solution: Dissolve 5000 mg L<sup>-1</sup> lanthanum oxide (La<sub>2</sub>O<sub>3</sub>) and 2000 mg L<sup>-1</sup> cesium chloride solution: Dissolve 4.691 g La<sub>2</sub>O<sub>3</sub> and 5.071 g CsCl in 1.5 L of deionized water and add 25 mL of HClO<sub>4</sub> and 25 mL of HNO<sub>3</sub> and dilute to 2000 mL. Check solution for contamination of K, Na and/or other cations.
4. Standard calibration solutions of K, Ca, Mg and Na. Prepare six calibration standard solutions of 1.0 - 20 mg L<sup>-1</sup> of K, 1.0 - 70 mg L<sup>-1</sup> of Ca, and 0.50 - 20 mg L<sup>-1</sup> for Mg and Na prepared from 1000 mg L<sup>-1</sup> stock solutions. Dilute calibration solutions with a solution containing the chemical interference solution and 0.04 N ammonium acetate solution.

## Procedure

1. Weigh 2.50 ± 0.05 g of air dried soil pulverized to pass 10 mesh sieve (< 2.0 mm) into a 120 mL extraction vessel (See Comment #2). Add 25.0 mL of NH<sub>4</sub>OAc extraction solution and place on orbital mechanical shaker for thirty (30) minutes (See Comment #3). Include a method blank.
2. Filter, refilter if filtrate is cloudy (See Comment #4 and #5).
3. Dilute 1:25 an aliquot of the soil extract with the chemical interference solution.
4. Adjust and operate AAE, AAS or ICP-AES instrument with six standards in accordance with manufacturer's instructions. Calibrate using prepared calibration solutions. Determine K, Mg, Ca and Na concentration of extract, of method blank, and record results as mg L<sup>-1</sup> of cation in solution.

## Calculations

Report as mg kg<sup>-1</sup> of K, Ca, Mg and Na to the nearest 1 mg kg<sup>-1</sup> (See Comment #7 and #8):

$\text{mg kg}^{-1}$  cations in soil =  $(\text{mg L}^{-1}$  cation in solution - method blank)  $\times 25 \times 10$

### Comments

1. Extraction solution should be prepared monthly, subject to microbial growth.
2. Soils high in soluble salts ( $\text{EC}_e > 1 \text{ dS m}^{-1}$ ) should be washed with deionized water before adding extraction with ammonium acetate to reduce potential errors of soluble salts. This can be accomplished by adding 50 mL of deionized water to 2.5 g of soil, placed in a 50 mL centrifuge tube and centrifuge at 5000 rpm for thirty (30) minutes and decanting excess water followed by the standard analysis protocol. For soils having  $\text{pH} > 7.4$  and calcium carbonate  $> 0.5\%$ , should be extracted with ammonium acetate  $\text{pH} 8.5$  to avoid the dissolution of calcium carbonate.
3. Check repipette dispensing volume, calibrate using an analytical balance.
4. Extracts may be stored for one week under refrigeration.
5. Filter paper should be checked periodically (monthly) for possible contamination of alkali metals. If contamination is  $> 5 \text{ mg kg}^{-1}$  on a soil basis select an alternative supply of filter paper.
6. Results may be expressed as  $\text{meq } 100 \text{ g}^{-1}$  or  $\text{cmol kg}^{-1}$ . Divide concentration of K, Ca, Mg and Na by 391 mg, 200 mg, 121.5 mg, and 230 mg respectively.
7. Generally, soils having less than  $100 \text{ mg kg}^{-1}$  potassium will respond to applications of potassium fertilizers for most crops.

### Literature

- Doll, E.C. and R. E. Lucas. 1973. Testing soil for potassium, calcium and magnesium. p 133-152. *In*: L.M. Walsh and J.D. Beaton. (ed.) Soil testing and plant analysis. SSSA Madison, WI.
- Knudsen, D., G.A. Peterson and P.F. Pratt. 1982. Lithium, sodium and potassium. *In*: A.L. Page (ed) Methods of soil analysis Part 2. Agronomy Monograph 9. 2nd ed. ASA and SSSA, Madison WI.
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## Scope and Application

This method semiquantitatively determines the amount of soil plant available K, Ca, Mg, and Na residing on the soil colloid exchange sites by displacement with ammonium acetate solution buffered to pH 8.5. Cation concentrations are determined using atomic emission (AES), absorption spectrometry (AAS) or ICP-AES instrumentation. A chemical interference solution is used to minimize chemical matrix effects for the AES method. It is based on a modification of the procedure outlined by Knudsen et al. (1982) for exchangeable K but is buffered to pH 8.5 to minimize dissolution of calcium carbonate ( $\text{CaCO}_3$ ). Generally, these cations are associated with the exchange sites. The exception are soils that have high soluble salts and are saline, which requires a special preanalysis treatment. The method doesn't correct for calcium and magnesium extracted as free carbonates or gypsum. In the northern Great Plains the method has been used to determine available sulfur. The method detection limit is approximately of  $25 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible  $\pm 7\%$ .

## Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Reciprocating horizontal mechanical shaker, 180 oscillations per minute (opm).
3. Repipette dispenser, calibrated to  $25.0 \pm 0.2 \text{ mL}$ .
3. Extraction vessel and associated filtration vessel.
4. Whatman No. 42, No. 2 or equivalent highly retentive filter paper.
5. Atomic Emission/Absorption Spectrophotometer (AAE) (AAS) or ICP-AES.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Ammonium acetate (1.0 N) extraction solution neutral @ pH 8.50: Add 570 mL of glacial acetic acid  $\text{CH}_3\text{COOH}$  (99%) to 8000 mL of deionized water. Add 680 mL of concentrated ammonium hydroxide adjust pH to 8.5 with 3.0 N glacial acetic acid or 3.0 N ammonium hydroxide and dilute to 10 L final volume. Check solution for possible contamination of K, Na, Ca, and Mg (see comment #1).
3. Chemical interference solution: Dissolve  $5000 \text{ mg L}^{-1}$  lanthanum oxide ( $\text{La}_2\text{O}_3$ ) and  $2000 \text{ mg L}^{-1}$  cesium chloride solution: Dissolve 4.691 g  $\text{La}_2\text{O}_3$  and 5.071 g CsCl in 1.5 L of deionized water and add 25 mL of  $\text{HClO}_4$  and 25 mL of  $\text{HNO}_3$  and dilute to 2000 mL. Check solution for contamination of K, Na and/or other cations.
4. Standard calibration solutions of K, Ca, Mg and Na. Prepare six calibration standard solutions of 1.0 - 20  $\text{mg L}^{-1}$  of K, 1.0 - 70  $\text{mg L}^{-1}$  of Ca, and 0.50 - 20  $\text{mg L}^{-1}$  for Mg and Na prepared from 1000  $\text{mg L}^{-1}$  stock solutions. Dilute calibration solutions with a solution containing the chemical interference solution and 0.04 N ammonium acetate solution.

## Procedure

1. Weigh  $2.50 \pm 0.05 \text{ g}$  of air dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into a 120 mL extraction vessel (See Comment #2). Add 25.0 mL of  $\text{NH}_4\text{OAc}$  extraction solution and place on orbital mechanical shaker for thirty (30) minutes (See Comment #3). Include a method blank.
2. Filter, refilter if filtrate is cloudy (See Comment #4 and #5).
3. Dilute 1:25 an aliquot of the soil extract with the chemical interference solution.
4. Adjust and operate AAE, AAS or ICP-AES instrument with six standards in accordance with manufacturer's instructions. Calibrate using prepared calibration solutions. Determine K, Mg, Ca and Na concentration of extract, of method blank, and record results as  $\text{mg L}^{-1}$  of cation in solution.

## Calculations

Report as mg kg<sup>-1</sup> of K, Ca, Mg and Na to the nearest 1 mg kg<sup>-1</sup> (See Comment #7 and #8):

mg kg<sup>-1</sup> cations in soil = (mg L<sup>-1</sup> cation in solution - method blank) × 25 × 10

### Comments

1. Extraction solution pH should be monitored daily.
2. Soils high in soluble salts ( $EC_e > 1 \text{ dS m}^{-1}$ ) should be washed with deionized water before adding extraction with ammonium acetate to reduce potential errors of soluble salts. This can be accomplished by adding 50 mL of deionized water to 2.5 g of soil, placed in a 50 mL centrifuge tube and centrifuge at 5000 rpm for thirty (30) minutes and decanting excess water followed by the standard analysis protocol. For soils having pH > 7.4 and calcium carbonate > 0.5%, should be extracted with ammonium acetate pH 8.5 to avoid the dissolution of calcium carbonate.
3. Check repipette dispensing volume, calibrate using an analytical balance.
4. Extracts may be stored for one week under refrigeration.
5. Filter paper should be checked periodically (monthly) for possible contamination of alkali metals. If contamination is > 5 mg kg<sup>-1</sup> on a soil basis select an alternative supply of filter paper.
6. Results may be expressed as meq 100 g<sup>-1</sup> or cmol kg<sup>-1</sup>. Divide concentration of K, Ca, Mg and Na by 391 mg, 200 mg, 121.5 mg, and 230 mg respectively.
7. Generally, soils having less than 100 mg kg<sup>-1</sup> potassium will respond to applications of potassium fertilizers for most crops.

### Literature

Doll, E.C. and R. E. Lucas. 1973. Testing soil for potassium, calcium and magnesium. p 133-152. *In*: L.M. Walsh and J.D. Beaton. (ed.) Soil testing and plant analysis. SSSA Madison, WI.

Knudsen, D., G.A. Peterson and P.F. Pratt. 1982. Lithium, sodium and potassium. *In*: A.L. Page (ed) Methods of soil analysis Part 2. Agronomy Monograph 9. 2nd ed. ASA and SSSA, Madison WI.

McKeague, J.A. ed. 1981. Extractable cations. *In*: Manual of soil sampling and methods of analysis. Canadian Soil Survey Committee, prepared by subcommittee of methods of analysis.

Munter, R. 1988. Laboratory factors affecting the extractability of nutrients. p. 8-10. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region, North Dakota Agricultural Experiment Station Bulletin No. 499 (revised).

## Scope and Application

This method semiquantitatively determines the amount of soil K fixation potential associated with illitic clay soils. Potassium is added to soil, then dried and incubated for seven days. Potassium concentration is determined using atomic emission (AES), absorption spectrometry (AAS) and inductively Coupled Plasma (ICP) instrumentation and fixation capacity calculated based on the amount of K recovered. It is based on a modification of the procedure first proposed by Cassman et al. (1982) and that developed at the University California Davis for evaluating the fixation of potassium for soils of the San Joaquin valley of California. Data for cotton in California indicates that the test is useful for predicting K response on sub soils which have a high K fixation potential Miller et al. (1997). Generally, the method is reproducible within  $\pm 7\%$  K on a given soil, or 3 - 5% on a soil K fixation basis.

## Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.01$  g.
2. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute (opm).
3. Repipette dispenser, calibrated to  $25.0 \pm 0.2$  mL.
3. 50 mL graduated centrifuge tube.
4. Whatman No. 42, No. 2 or equivalent highly retentive filter paper.
5. Atomic Emission/Absorption Spectrophotometer (AES) or (AAS).

## Reagents

1. Deionized water, ASTM Type I grade.
2.  $\text{KNO}_3$  0.01 N solution: dissolve 1.011 g of potassium nitrate in 1.0 L of deionized water.
3.  $\text{CaCl}_2$ , 0.01 N solution: dissolve 2.94 g of  $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$  in 1.0 L of deionized water.
4.  $\text{NH}_4\text{Cl}$  3.0 N solution: dissolve 160.3 g of ammonium chloride in 1.0 L of deionized water.
5. Standard calibration solutions of K. Prepare two (2) sets of six (6) calibration standard solutions of  $1.0 - 20 \text{ mg L}^{-1}$  of K from  $1000 \text{ mg L}^{-1}$  stock solutions prepared in matrix of 0.01  $\text{CaCl}_2$  and cesium chloride ( 5.071 g CsCl in 2.0 L of deionized water); and a second set of K calibration standards in a matrix of 1.0  $\text{NH}_4\text{Cl}$  and cesium chloride (5.071 g CsCl in 2.0 L of deionized water).

## Procedure

1. Weigh  $3.00 \pm 0.05$  g of air dried soil pulverized to pass 10 mesh sieve ( $< 2.0$  mm) into a 50 mL volumetrically marked centrifuge tube.
2. Add 3.0 mL 0.01 N  $\text{KNO}_3$  into the centrifuge tube. Allow to air dry 48 hours. Preferred method is to place in high draft hood for forty-eight (48) hours at room temperature, avoid heating above  $30^\circ \text{C}$ . Include two centrifuge tubes containing only 3.0 mL of 10mM  $\text{KNO}_3$  as a control (See Comment #1).
3. When dry, add 30 mL of 0.01 M  $\text{CaCl}_2$  solution and 100  $\mu\text{L}$  of toluene. Cap and place horizontally on a reciprocating shaker for thirty (30) minutes (See Comment #2).
4. Incubate for seven (7) days at  $25^\circ \text{C}$ , shaking for 30 minutes daily.
5. After seven (7) days, shake thirty (30) minutes, then centrifuge 2000 rpm until supernatant is clear. Remove a 10.0 mL aliquot of supernatant and determine K concentration using matrix matched standards (prepared in 0.01 M  $\text{CaCl}_2$  solution) by AAS or AES.
6. Add 10 mL of 3.0 N  $\text{NH}_4\text{Cl}$ , and shake for 30 minutes. Centrifuge and determine K concentration in supernatant using matrix matched ( $1.0 \text{ N NH}_4\text{Cl}$  and  $0.01 \text{ M CaCl}_2$ ) standards by atomic absorption. Use control solution to verify 100% K recovery of  $\text{KNO}_3$  control or as substitution of added K value of 1170.

7. Adjust and operate AES or AAS instrument with six standards in accordance with manufacturer's instructions. Calibrate using prepared calibration solutions. Determine K, concentration of extract, of method blank, and record results as mg L<sup>-1</sup> of cation in solution.

### Calculations

Calculation of % K Fixation:

$$\frac{1170 - [(10 \times \text{mg L}^{-1} \text{ K in } 0.01 \text{ M CaCl}_2 \text{ supernatant}) + (30 \times \text{mg L}^{-1} \text{ K in NH}_4\text{Cl})]}{1170} \times 100$$

Note: Negative values indicate soil has released more K than was fixed.  
Positive values indicate soil mineral fixation of potassium (see Comment #3).

### Comments

1. Check repipette dispensing volume, calibrate using an analytical balance.
2. Toluene is added to prevent microbial growth.
3. Generally, cotton grown in California on subsoils (depths 5-15 inches) with fixation potential greater than 60%, are highly response to K fertilizers with respect to lint yields.

### Literature

Cassman, K.G., D.C. Bryant and B.A. Roberts 1990. Comparison of soil test methods for predicting cotton response to soil fertilizer potassium on potassium fixing soils. Commun. in Soil Sci. Plant Anal. 21(13-16), 1727-1743.

Miller, R.O., B.L. Weir, B. Roberts, R. Vargas, D. Munk, S. Wright, D. Munier, R. Travis and M. Keeley. 1997. Potassium in fertility in cotton production systems. UC Extension Publication #21562.

Munter, R. 1988. Laboratory factors affecting the extractability of nutrients. p. 8-10. In: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region, North Dakota Agricultural Experiment Station Bulletin No. 499 (revised).

## Scope and Application

The DTPA (diethylenetriaminepentaacetic acid) micronutrient extraction method, developed by Lindsay and Norvell (1978), is a nonequilibrium extraction for estimating the potential soil bioavailability of Zn, Cu, Mn, and Fe for neutral and calcareous soils. Analyte concentrations are determined by atomic absorption spectrometry (AAS) or inductively coupled plasma atomic emission spectrometry (ICP-AES). The quantity of micronutrient and trace metals extracted, are affected by solution pH, soil extraction ratio, shaking time, extraction time, and extractant concentration. The method has shown to be well correlated to crop response to zinc and copper fertilizers. It also shows potential for monitoring Cd, Ni and Pb in soils receiving sludge applications. The method detection limit is approximately  $0.1 \text{ mg kg}^{-1}$  for Zn, Cu, Mn, and Fe (on a dry soil basis) and is generally reproducible  $\pm 10\%$ , except for Fe which is approximately  $\pm 15\%$ .

## Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute (opm).
3. Repipette dispenser, calibrate to  $20.0 \pm 0.1 \text{ mL}$ .
4. Extraction vessel and associated filtration vessel.
5. Whatman No. 42 or No. 2, or equivalent medium retentive filter paper.
6. Atomic Absorption Spectrometer (AAS) or ICP-AES instrumentation.

## Reagents

1. Deionized water, ASTM Type I grade.
2. DTPA extraction solution, 0.005 M: Dissolve 0.005 M DTPA (diethylenetriaminepentaacetic acid), 0.01 M  $\text{CaCl}_2$  and 0.10 M Triethanolamine (TEA) adjusted to pH 7.3. For 10 L solution, dissolve 19.67 g DTPA and 149.2 g of TEA in 5 L deionized water. Add 14.69 g of  $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$  to 5 L of deionized water (See Comments #1 and #2) and add DTPA-TEA mixture and adjust the pH to 7.3  $\pm 0.05$  using 1.0 N HCl (approximately 41.5 mL of concentrated HCl are required).
3. Standard calibration solutions, Zn, Cu, Mn, and Fe. Prepare six calibration of 100 mL solutions of Zn, Cu, Mn, and Fe diluted in DTPA extraction solution from 1000  $\text{mg L}^{-1}$  standard solutions. Concentration ranges: 0.1 - 5.0  $\text{mg L}^{-1}$  for Zn; 0.1 - 10  $\text{mg L}^{-1}$  for Mn and Fe; and 0.05 - 4.0  $\text{mg L}^{-1}$  for Cu.

## Procedure

1. Weigh  $10.0 \pm 0.2 \text{ g}$  of air dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into 120 mL extraction vessel and add 20.0 mL of DTPA extraction reagent (See Comments #3 and #4). Include a DTPA method blank.
2. Place extraction vessel(s) on reciprocating mechanical shaker for two (2) hours at  $25 \text{ }^\circ\text{C}$ , and 180 oscillations per minute.
3. Immediately filter (See Comment #5), refilter if filtrate is cloudy.
4. Adjust and operate AAS instrument in accordance with manufacturer's instructions. Calibrate instrument using standard calibration solutions and determine individually (Zn, Cu, Mn and Fe) concentration of DTPA extracts, method blank and record as  $\text{mg L}^{-1}$  of analyte in solution. (See Comments #6 and #7).

## Calculations

Report as mg kg<sup>-1</sup> Mn and Fe in the soil to the nearest 0.1 mg kg<sup>-1</sup>, report Zn and Cu in the soil to the nearest 0.01 mg kg<sup>-1</sup>;

$$\text{mg kg}^{-1} \text{ in soil} = (\text{mg L}^{-1} \text{ in extract} - \text{method blank}) \times 2.0$$

(See Comment #8)

## Comments

1. Calcium chloride is added to the DTPA solution to prevent dissolution of calcium carbonate. This prevents dissolution of zinc present in calcium carbonate minerals.
2. TEA is added to buffer the DTPA solution and prevent excess dissolution of trace metals which is highly pH dependent. Check solution pH bimonthly and adjust as needed.
3. Check repipette dispensing volume, calibrate using an analytical balance.
4. Precautions should be taken to avoid potential contamination by: (1) rinsing all labware in a solution of 0.5 N HCl; (2) reagent solutions made in ASTM Type I water; and (3) filter paper checked for potential contamination.
5. Filtering should commence immediately after shaking, and should cease 15 within minutes after starting. Continued filtering will increase the amount of micro nutrient extracted.
6. Metal concentrations exceeding highest standard will require dilution for analysis.
7. Manganese concentrations will likely increase over time on soils stored for periods greater than six (6) months, the result of high oxidation. This can be minimized by storing soils in air-tight containers.
8. Soils having DTPA extractable zinc less than 0.80 mg kg<sup>-1</sup> are generally responsive to zinc fertilizer for sensitive crops such as corn. Soils having DTPA extractable copper less than 0.60 mg kg<sup>-1</sup> are generally responsive to copper fertilizer for sensitive crops such as wheat.

## Literature

Horneck, D. A., J. M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. Ag. Expt. Station SM 89:4.

Liang, J. and R.E. Karamanos. 1993. Soil reaction and exchangeable acidity. p. 87-90. *In*: M. R. Carter (ed.) Soil sampling and methods of analysis, Canadian Society of Soil Science, Lewis Publishers Ann Arbor, MI.

Lindsay, W. L. and W. A. Norvell. 1978. Development of a DTPA soil test for zinc, iron, manganese, and copper. Soil Sci. Soc. Amer. J. 42:421-428.

Whitney, D. 1988. Micronutrient soil test for zinc, iron, manganese, and copper. p. 20-22. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agricultural Experiment Station Bulletin 499 (revised).

## Scope and Application

The DTPA - Sorbitol (diethylenetriaminepentaacetic acid) micronutrient extraction method, is based on that developed by Lindsay and Norvell (1978), is a nonequilibrium extraction for estimating the potential soil bioavailability of Zn, Cu, Mn, Fe and B for mildly acid to calcareous soils. The addition of Sorbitol is added for the chelate extraction of boron. Analyte concentrations are determined by inductively coupled plasma atomic emission spectrometry (ICP-AES). The quantity of micronutrient and trace metals extracted, are affected by solution pH, soil extraction ratio, shaking time, extraction time, and extractant concentration. The method has shown to be well correlated to crop response to zinc and copper fertilizers. It also shows potential for monitoring Cd, Ni and Pb in soils receiving sludge applications. The method detection limit is approximately  $0.1 \text{ mg kg}^{-1}$  for Zn, Cu, Mn, and Fe and  $0.05 \text{ mg kg}^{-1}$  for B (on a dry soil basis). It is generally reproducible  $\pm 10\%$ , except for Fe which is approximately  $\pm 15\%$ .

## Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute (opm).
3. Repipette dispenser, calibrate to  $20.0 \pm 0.1 \text{ mL}$ .
4. Extraction vessel and associated filtration vessel - 50mL volume.
5. Whatman No. 42 or No. 2, or equivalent medium retentive filter paper.
6. ICP-AES, inductively coupled plasma atomic emission spectrometry (ICP-AES).

## Reagents

1. Deionized water, ASTM Type I grade.
2. 5% sodium hypochlorite solution.
3. DTPA-Sorbitol extraction solution, 0.005 M. Rinse all labware with 5% sodium hypochlorite solution. Dissolve 0.005 M DTPA (diethylenetriaminepentaacetic acid), 0.01 M  $\text{CaCl}_2$  and 0.10 M Triethanolamine (TEA) adjusted to pH 7.3 and 0.20M of Sorbitol. For 10 L solution, dissolve 19.67 g DTPA and 149.2 g of TEA in 5 L deionized water. Add 14.69 g of  $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$  to 5 L of deionized water (See Comments #1 and #2), 364.4g Sorbitol (Sigma, S-1876, D-Sorbitol, 98% minimum) and add DTPA-TEA mixture and adjust the pH to  $7.30 \pm 0.05$  using 1.0 N HCl (approximately 41.5 mL of concentrated HCl are required).
4. Standard calibration solutions, Zn, Mn, Fe, Cu and B. Prepare six calibration solutions of 100 mL volume of Zn, Cu, Mn, and Fe diluted in DTPA-Sorbitol extraction solution from 1000  $\text{mg L}^{-1}$  standard solutions. Concentration ranges: 0.1 - 10.0  $\text{mg L}^{-1}$  for Zn; 1.0 - 30  $\text{mg L}^{-1}$  for Mn and Fe; 0.05 - 4.0  $\text{mg L}^{-1}$  for Cu; and 0.02 - 5.0  $\text{mg L}^{-1}$  for B.

## Procedure

1. Weigh  $10.0 \pm 0.2 \text{ g}$  of air dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into 50 mL extraction vessel and add 20.0 mL of DTPA - Sorbitol extraction reagent (See Comments #3 and #4). Include a DTPA method blank.
2. Place extraction vessel(s) on reciprocating mechanical shaker for two (2) hours at  $25 \text{ }^\circ\text{C}$ , and 180 oscillations per minute.
3. Immediately filter, refilter if filtrate is cloudy (See Comment #5).
4. Adjust and operate ICP-AES instrument in accordance with manufacturer's instructions. Calibrate instrument using standard calibration solutions and determine individually (Zn, Mn, Cu, Fe and B)

concentration of DTPA-Sorbitol extracts, method blank and record as mg L<sup>-1</sup> of analyte in solution. (See Comments #6 and #7).

### Calculations

Report as mg kg<sup>-1</sup> Mn and Fe in the soil to the nearest 0.1 mg kg<sup>-1</sup>, report Zn, Cu and B in the soil to the nearest 0.01 mg kg<sup>-1</sup>;

$$\text{mg kg}^{-1} \text{ in soil} = (\text{mg L}^{-1} \text{ in extract} - \text{method blank}) \times 2.0$$

(See Comment #8 and #9)

### Comments

1. Calcium chloride is added to the DTPA solution to prevent dissolution of calcium carbonate. This prevents dissolution of zinc present in calcium carbonate minerals.
2. TEA is added to buffer the DTPA solution and prevent excess dissolution of trace metals which is highly pH dependent. Check solution pH bimonthly and adjust as needed.
3. Check repipette dispensing volume, calibrate using an analytical balance.
4. Precautions should be taken to avoid potential contamination by: (1) rinsing all labware in a solution of 0.5 N HCl; (2) reagent solutions made in ASTM Type I water; and (3) filter paper checked for potential contamination.
5. Soil DTPA-Sorbitol suspension filters slowly. Filtering should commence immediately after shaking, and should cease 15 within minutes after starting. Continued filtering will increase the amount of micro nutrient extracted.
6. Metal concentrations exceeding highest standard will require dilution for analysis.
7. Manganese concentrations will likely increase over time on soils stored for periods greater than six (6) months, the result of oxidation. This can be minimized by storing soils in air-tight containers.
8. Soils having DTPA extractable zinc less than 0.8 mg kg<sup>-1</sup> are generally responsive to zinc fertilizer for sensitive crops such as corn. Soils having DTPA extractable copper less than 0.6 mg kg<sup>-1</sup> are generally responsive to copper fertilizer for sensitive crops such as wheat.
9. The DTPA-Sorbitol extraction method extracts 5-8% less Zn, Mn, Fe and Cu than the standard DTPA extraction procedure (S-6.10). For specific soils with greater than 4% SOM, this method will extract 60% Mn of the extracted by the DTPA procedure (S-6.10). Generally DTPA-Sorbitol extracts 10% less B than the hot-water method (S-7.10).



## Literature

Geasting, W.D. and P.N. Soltanpour. 1981. Boron analysis in soil extracts and plant tissue by plasma emission spectroscopy. *Commun. Soil Sci. Plant Anal.* 12(8):733-742.

Horneck, D. A., J. M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. *Ag. Expt. Station SM 89:4.*

Liang, J. and R.E. Karamanos. 1993. Soil reaction and exchangeable acidity. p. 87-90. *In: M. R. Carter (ed.) Soil sampling and methods of analysis, Canadian Society of Soil Science, Lewis Publishers Ann Arbor, MI.*

Lindsay, W. L. and W. A. Norvell. 1978. Development of a DTPA soil test for zinc, iron, manganese, and copper. *Soil Sci. Soc. Amer. J.* 42:421-428.

Vaughan, B. and J. Howe. 1994. Evaluation of boron chelates in extracting soil boron. *Commun. Soil Sci. Plant Anal.* 25(7&8) 1071-1084.

Whitney, D. 1988. Micronutrient soil test for zinc, iron, manganese, and copper. p. 20-22. *In: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agricultural Experiment Station Bulletin 499 (revised).*

## Introduction

This method estimates the relative bioavailability of boron in soils using 0.02 M  $\text{CaCl}_2$  in conjunction with hot water which facilitates the dissolution of boron. The extraction of boron is based on that described by Bingham (1982) and adapted for the use of plastic bags as described by Mahler et al. (1983). Boron is determined spectrophotometrically by complexation of  $\text{HBO}_3$  with azomethine-H to which forms a colored complex. Subsequent measurement is at 420 nm (Horneck et al., 1989). This method is used to assess crop potential deficiencies of boron. It differs from the saturation paste soluble boron concentration (Method S - 1.40) which assesses possible supra-optimal concentrations of soil boron. The method detection limit is approximately  $0.2 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible  $\pm 12\%$ .

## Equipment

1. Analytical balance: 1000 g capacity, and minimum  $\pm 0.01 \text{ g}$ .
2. Plastic bags, sealable, 500 mL, 4 mil thickness.
3. Repipette dispenser, calibrated to  $30.0 \pm 0.3 \text{ mL}$ .
4. Pipette  $1.00 \pm 0.05$  and  $4.00 \pm 0.1 \text{ mL}$ .
5. Hot Plate and 4 liter beaker.
6. Vortex stirrer.
7. Whatman No. 42 or No. 2, or equivalent highly retentive filter paper.
8. Polypropylene sample tube (See Comment #1).
9. Spectrophotometer or automated Flow Injection Analysis (FIA) system, 420 nm.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Calcium chloride extraction solution, 0.02 M: Dissolve 2.84 g of  $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$  into 750 mL deionized water and dilute to 1000 mL.
3. Buffer masking agent: Dissolve completely 250 g ammonium acetate (reagent grade  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ ), 25 g disodium salt of ethylenedinitrilo-tetraacetic acid ( $\text{Na}_2$ -EDTA), and 10 g disodium salt of nitrilotriacetic acid ( $\text{Na}_2$ -NTA) in 400 mL deionized water in a 1 L beaker using a magnetic stirrer. Add 125 mL glacial acetic acid very slowly, while stirring. The temporary acidic conditions may cause a slight precipitation of the EDTA salts, stir solution until all the EDTA redissolves. Do not heat the solution. Adjust the buffer to a pH of 5.4 to 5.6 with acetic acid or  $\text{NH}_4\text{OH}$ . Prepare fresh solution every two months (See Comment #2).
4. Azomethine-H solution: Dissolve 0.9 g azomethine-H reagent (Pierce Chemical Co., Rockford, IL) and 2.0 g L-ascorbic acid ( $\text{C}_6\text{H}_8\text{O}_6$ ) in 50 mL of deionized water prewarmed to  $60^\circ\text{C}$ . A hot water bath facilitates dissolution. Dilute to 100 mL volume with deionized water and store in refrigerator. Solution is stable for forty-eight (48) hours.
5. Boron standard calibration solutions. Prepare five boron standards ranging from 0.1 to  $4.0 \text{ mg L}^{-1}$  boron from  $1000 \text{ mg L}^{-1}$  standard solution and diluted to volume with 0.02 M  $\text{CaCl}_2$ .

## Procedure

1. Weigh  $15.0 \pm 0.1 \text{ g}$  of air dry soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) and place in plastic bag. Add 30 mL of 0.02 M  $\text{CaCl}_2$  reagent using repipette dispenser (See Comment #3) and close bag. Include a method blank.
2. Place plastic bags into 4 L beaker containing boiling water and leave for ten (10) minutes.
3. Remove plastic bags, cool one ( $1 \pm 0.1$ ) minute and filter. Refilter if extract is cloudy.
4. Pipette 4.0 mL of soil extract into a 12 mL polyethylene sample tube.
5. Add 1.0 mL of buffer masking agent and stir with a vortex stirrer.
6. Add 1.0 mL of azomethine-H solution, vortex stir and allow to stand for sixty (60) minutes but no

- longer than three (3) hours (See Comments #4 and #5).
7. Prepare standard curve following steps 4 - 6, substituting 4.0 mL of standard calibration solution for soil extract. A method blank is prepared in the same manner using 4.0 mL CaCl<sub>2</sub> extracting solution instead of the soil extract.
  8. Read sample absorbance on a spectrophotometer set at 420 nm of a method blank, unknown samples and record results as mg L<sup>-1</sup> of B in the solution extract (See Comments #6 and #7).

### Calculations

Report as mg kg<sup>-1</sup> boron in the soil to the nearest 0.1 mg kg<sup>-1</sup>:

$$B \text{ mg kg}^{-1} \text{ in soil} = (\text{mg L}^{-1} \text{ B in extract} - \text{method blank}) \times 2$$

### Comments

1. Teflon or polypropylene labware is recommended to minimize the potential of boron contamination from boro silicate glass.
2. The EDTA and NTA chelates eliminate interferences from Al, Fe, and Cu. The concentration of these chelates should be effective for levels of these elements commonly found in soil extracts.
3. Check pipette and repipette dispensers delivery volume, recalibrate using an analytical balance.
4. The azomethine-H should be added quickly so color development is equal for all tubes. A constant check must be maintained on linearity and drift of standards when analyzing a large number of samples.
5. For soil samples with a yellow extract: Prepare a second sample solution and blank following steps 4 and 5. Add 1.0 mL of deionized water in place of azomethine-H solution. The blank for this determination consists of 5.0 mL CaCl<sub>2</sub> extracting solution and 1.00 mL buffer masking agent.
6. Boron may also be determined directly on the hot-water extract using ICP-AE using 249.699 or 208.14 nm wavelengths.
7. Dilution will be required on samples having boron concentrations exceeding the highest standard.

### Literature

Bingham, F.T. 1982. Boron. p.431-446. *In*: A.L. Page et al. (ed.) Methods of soil analysis, Part 2. Agronomy Mongr. 9. 2nd ed. ASA and SSSA, Madison, WI.

Horneck, D. A., J. M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. Ag. Expt. Station SM 89:4.

Mahler, R.L. D.V. Naylor and M.K. Fredickson. 1983. Hot water extraction of boron from soils using sealed plastic pouches. unpublished, Univ. of Idaho. 12.

## Scope and Application

The Total Kjeldahl Nitrogen (TKN) method is based on the wet oxidation of soil organic matter using sulfuric acid and digestion catalyst and conversion of organic nitrogen to ammonium nitrogen. Ammonium is determined using spectrophotometric, diffusion-conductivity or distillation techniques. The method is readily adapted to manual or automated techniques. The procedure does not quantitatively digest nitrogen from heterocyclic compounds (bound in a carbon ring), oxidized compounds such as nitrate, or ammonium from within mineral lattice structures. The method has a detection limit of approximately 0.020% N and is generally reproducible within  $\pm 8\%$ .

## Equipment

1. Analytical balance: 250 g capacity minimum  $\pm 0.1$  mg.
2. Acid fume hood.
2. Volumetric digestion tubes, 75 mL and digestion heating block (400 °C).
3. Repipette dispenser, calibrated to  $3.0 \pm 0.2$  mL.
4. Spectrophotometer, diffusion-conductivity instrument or distillation apparatus.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Sulfuric acid, concentrated - reagent grade.
3. Digestion Catalyst ( $K_2SO_4$ ,  $CuSO_4$ , and  $SeO$ : ratio 100:10:1), Kjel-tab.
4. Standard Calibration solutions ( $NH_4-N$ ). Prepare six working standards of ammonium, concentration range 0.1 - 40  $mg L^{-1}$ , made from 1000  $mg L^{-1}$  ammonium nitrogen standard solution and diluted to volume with 12 % sulfuric acid (v/v).

## Procedure

1. Weight  $1.000 \pm 0.005$  g of air dry soil to pass 10 mesh sieve ( $< 2.0$  mm) into a 75.0 mL volumetric digestion tube (See Comment #1). Include a method blank.
2. Add digestion catalyst, (200 mg of mixed catalyst or Kjel-tab) and 3.0 mL of concentrated sulfuric acid (See Comment #2 and #3). Note: it is essential that all dry material be completely moistened by acid and well mixed to insure complete digestion.
3. Place tubes on a digestion block at 150 °C and after thirty (30) minutes raise to 350 °C for two (2) hours or until samples are completely digested. At completion, mineral soils will be whitish-gray and organic soils blue-green in color.
4. Remove samples from block and place under fume hood for 5-10 minutes. Add 10-20 mL of deionized water using a wash bottle to each tube to prevent hardening and crystal formation. Dilute digest to volume with deionized water, cap, invert three times, and allow digest to clear.
5. Determine digest ammonium concentration using the spectrophotometric, diffusion-conductivity instruments or distillation techniques using standard calibration solutions (See Comment #4 and #5). The ammonium nitrogen content of the digest solution can be determined with a rapid flow analyzer (Technicon Method No. 334-74A/A) or an flow injection analyzer (FIA). This determination can also be made using the Kjeldahl distillation method. Adjust and operate instruments in accordance with manufacturer's instructions. Determine ammonium concentration of a method blank, unknown samples and record ammonium concentration as  $mg L^{-1}$  of  $NH_4-N$  in soil digest.

## Calculations

Report total Kjeldahl nitrogen results to the nearest 0.001% as:

$$\% \text{ N} = \frac{(\text{mg L}^{-1} \text{ NH}_4\text{-N in digest} - \text{method blank}) \times (0.075) \times (100)}{(\text{Sample size mg})} [\text{equ. S -8.10-1}]$$

## Comments

1. Use 250 mg of soil if sample is greater than approximately 10% organic matter.
2. Check repipette dispenser delivery volume, recalibrate using an analytical balance.
3. When adding reagents to vessels always wear protective clothing (i.e. eye protection, lab coat, disposable lab gloves, and shoes). Always handle reagent and digestion labware in hoods capable of high air flow, 100 cfm.
4. The Kjeldahl method outline by Bremmer and Mulvaney (1982) is modified eliminating the water from the digestion step.
5. Kjeldahl soil acid digest is classified as hazardous waste and must be disposed of in a suitable manner.

## Literature

Bremmer, J. M. and C. S. Sulvaney. 1982. Total nitrogen p. 595-624. *In*: A.L. Page et al. (eds.) Methods of soil analysis, part 2. Agron. Mongr. 9. 2nd ed. ASA and SSSA, Madison, WI.

Carlson, R.M. 1978. Automated separation and conductimetric determination of ammonia and dissolved carbon dioxide. *Anal. Chem.* 48:1528-1531.

Horneck, D. A., J. M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. *Ag. Expt. Station SM 89:4.*

### Scope and Application

This method semiquantifies the amount of oxidizable organic matter as determined by the gravimetric weight change associated with high temperature oxidation of soil organic matter in a muffle furnace. The method is based that described by Storer (1984) and is an estimate. As a result of chromium hazardous waste disposal costs associated with Walkley-Black method (S - 9.00), many laboratories have chosen to switch to loss on ignition as a means for estimating soil organic matter content. The LOI method is poorly correlated with the Walkley-Black method for soils containing less than 3% organic matter. Soil organic matter values are used to estimate potential nitrogen mineralization, pesticide management and for crop production management. The method is generally reproducible within  $\pm 20\%$ .

### Equipment

1. Analytical balance: 100 g capacity, resolution  $\pm 0.001$  g.
2. High temperature, crucibles 20 cc capacity.
3. Drying oven, 105 °C.
4. Dessicator, containing desiccating agent.
5. Muffle furnace capable of heating to 360 °C.

### Reagents

1. Calcium carbonate, reagent grade (See Comment #1).

### Procedure

1. Weigh  $10.0 \pm 1.0$  g of air dry soil pulverized to pass 10 mesh sieve ( $< 2.0$  mm) into a preweighed crucible (record to the nearest  $\pm 0.001$  g). Prepare a crucible containing calcium carbonate (See Comment #1).
2. Place in drying oven for two (2) hours at 105 °C. Place in dessicator for one (1) hour.
3. Record crucible + soil as Initial wt to the nearest  $\pm 0.001$  g.
4. Heat in muffle furnace to 360 °C for two (2) hours (after temperature reaches 360 °C).
5. Place in drying oven at 105°C for one (1) hour and place in dessicator for one (1) hour.
6. Record crucible + soil as Final wt sample weight to the nearest  $\pm 0.001$  g.

### Calculations

$$\text{LOI \%} = \frac{(\text{Initial wt at } 105\text{ }^{\circ}\text{C} - \text{Final wt. at } 105\text{ }^{\circ}\text{C}) \times 100}{(\text{Initial wt at } 105\text{ }^{\circ}\text{C} - \text{crucible wt})} \quad [\text{equ. S -9.20-1}]$$

Estimation of organic matter by LOI is done by regression analysis with organic matter. Select fifty soils covering a range in organic matter expected, determine organic matter based on Walkley-Black method (Method S - 9.10) and LOI value. Use this equation to convert LOI values to estimated percent organic matter (See Comment #2).

## Comments

1. Calcium carbonate is included as a method standard to evaluate potential loss of carbonates of alkali metals. If appreciable losses (>0.05% weight change) are noted check temperature calibration of the muffle furnace, reduce oven temperature 10 °C and repeat.
2. The regression slope for organic matter for the LOI method on the Walkley-Black method ranges from 0.68 to 1.04 for soils reported in the literature. Regression intercept values range from -0.04 to -0.36, (Schultee and Hopkins, 1996).

## Literature

Schulte, E.E. 1988. Recommended soil organic matter tests. p. 29-31. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agricultural Experiment Station Bulletin 499 (revised).

Schulte, E.E. and B.G. Hopkins. 1996. Estimation of soil organic matter by weight Loss-On-Ignition. P. 21-32. *In*: Soil Organic matter: Analysis and Interpretation. (ed.) F.R. Magdoff, M.A. Tabatabai and E.A. Hanlon, Jr. Special publication No. 46. Soil Sci. Soc. Amer. Madison, WI.

Storer, D. A. 1984. A simple high sample volume ashing procedure for determining soil organic matter. *Commun. Soil Sci. Plant Anal.* 15:759-772.

## Scope and Application

This method quantitatively determines the amount of organic carbon in soil materials by combustion of the sample in an O<sub>2</sub> environment using an automated resistance furnace and with subsequent quantification of nitrogen by thermal conductivity detector and CO<sub>2</sub> using an infrared or conductivity detector. For very low nitrogen analysis (< 0.05%) specific instruments are available with chemi-luminescence detectors. Soils with a pH >7.4 (method S - 2.10) and containing inorganic carbon (carbonates), organic carbon is determined by the difference between total carbon by combustion minus the quantity of inorganic carbon as determined by Method S - 13.10 or S - 13.20. The method for specific instruments requires that soils be pulverized to pass 60 mesh sieve to ensure homogeneity. It is based on the method originally described by Dumas whereby soil samples encased in tin (Sn) foil, are ignited in a furnace in excess of 1000 °C, in a helium and oxygen environment in a quartz combustion tube. The combustion gas is passed through a catalyst (instrument manufacturer dependent) to complete conversion of CO to CO<sub>2</sub>, scrubbed of moisture and CO<sub>2</sub> determined by an infrared detector and for nitrogen by thermal conductivity detector. Specific instruments provide for the simultaneous determination of H or S. Total nitrogen and organic carbon is used to assess nutrient mineralization, water infiltration, soil structure and absorption or deactivation of agricultural chemicals. The method has a detection limit of 0.03% N and 0.02% TOC (dry sample basis instrument dependent) and is generally reproducible to within ±7.0% for nitrogen and ±5.0% for organic carbon..

## Equipment

1. Analytical balance: 250 g capacity, resolution ± 0.1 mg.
2. Total Nitrogen Analyzer: Leco CHN-1000, CNS-2000, Elementar, Carl-Erba, Perkin-Elmer or Elementar, with a resistance furnace with infrared and/or thermal conductivity detector and operating supplies.
3. Tin foil encapsulating capsules (See Comment #1).
4. Desiccator, containing a desiccating agent.

## Reagents

1. Compressed Oxygen, 99.99% purity.
2. Helium carrier gas 99.99% purity.
3. Total Organic Carbon calibrations standards: EDTA, 9.57% ± 0.05% N; sulfanilic acid (C<sub>6</sub>H<sub>7</sub>NO<sub>3</sub>S) 41.6%C; Leco part number 502-203 soil, 2.75% ±0.09 %C; and Leco part number 502-062 soil, 0.85% ±0.05%C.

## Procedure

1. Determine the soil moisture content (See comment #2).
2. Weigh of air dry soil (quantity is instrument dependent) pulverized to pass a 30 mesh sieve (< 250 μm) (See Comments #3 and #4 #5) and place in into a tarred tin foil container, encapsulate, close and record sample weight to the nearest 0.1 mg.
3. Initialize the instrument following manufacturers suggested protocol. Conduct a system leak check on combustion system. Perform blank stabilization test, analyze consecutive blanks until the blanks stabilize at a constant value (See Comment #6).
4. Adjust and operate the instrument according to manufacturer instructions using calibration standards. Enter sample dry matter content and analyze unknown sample for total nitrogen. Report results to the nearest 0.001% carbon (See Comment #7).



## Calculations

Report total nitrogen results to the nearest 0.001%

Report total organic carbon results to the nearest 0.01%

## Comments

1. Tin (Sn) foil capsules is utilized as combustion catalyst. Capsules can be obtained from the manufacturer's, after market vendors, or fabricated from sheets of tin foil material.
2. Samples limited in material, should be dried over phosphorus pentoxide or magnesium perchlorate for forty-eight (48) hours and analyzed with no correction for moisture content or reported on as received basis.
3. Sample particulate material must be ground to pass a 30 mesh sieve ( $< 600 \mu\text{m}$ ) for macro analysis instruments which require sample sizes in excess of 250 mg (i.e. LECO, CHN-2000 and Elementar) in order to assure adequate sample homogeneity. For instruments utilizing a sample sizes 5-10 mg (Carlo-Erba and Perkin-Elmer) samples must be ground to pass 100 or 140 ( $106\text{-}150 \mu\text{m}$ ) mesh sieve.
4. Soils containing free carbonates must be analyzed for free calcium carbonate according to Methods S - 13.10 or S- 13.20 for the determination of inorganic carbon. Total organic carbon is calculated by subtracting the inorganic carbon from the total carbon value determined by the combustion instrument.
5. Sample weight may be entered into instrument software using a balance interface.
6. All soil calibration samples should be: (1) checked for homogeneity; and (2) nitrogen and organic carbon content verified using standard addition techniques using two chemical standards such as EDTA and sulfanilic acid. A quality control certified reference sample (NIST SRM 2704,  $3.348\% \pm 0.10\% \text{ C}$ ) is available from the National Institute of Standards and technology, see appendix B.
7. To convert total soil organic carbon (%C) to soil organic matter (SOM by Walkely -Black method, S - 9.10), multiply %C by 1.724 to estimate soil organic matter. The conversion factor ranges from 1.6 to 2.5 dependent on the soil and cropping system management.

## Literature

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### Scope and Application

Cation exchange capacity (CEC) is the measure of a soil to retain readily exchangeable cations which neutralize the negative charge of soils. This method involves saturation of the cation exchange sites with ammonium, equilibration, removal of the excess ammonium with ethanol, replacement and leaching of exchangeable ammonium with protons from HCl acid (Horneck, et al. 1989). This method may be poorly suited to soils containing carbonates, vermiculite, gypsum and zeolite minerals. The procedure is time consuming and labor intensive. The speed at which samples filter depends on the strength of the vacuum applied and sample makeup. If using a water aspirated generated vacuum, some samples may never filter because of clogged filter paper. The method detection limit is approximately  $1.0 \text{ cmol kg}^{-1}$  (or meq/100 gm on a dry soil basis) and is generally reproducible within  $\pm 10\%$ .

### Equipment

1. Analytical balance: 250 g capacity, resolution 0.01 g.
2. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute.
3. Repipette dispensers, calibrated to  $20 \pm 0.1$ ,  $8.0 \pm 0.1$  mL and  $5.0 \pm 0.1$  mL.
4. Whatman No. 1, No. 2 or equivalent filter paper.
5. Buchner funnels vacuum flasks and source of vacuum
6. Auto analyzer or Kjeldahl distillation equipment.

### Reagents

1. Deionized water ASTM Type I Grade.
2. Ammonium acetate (1.0 N) extraction solution neutral: Add 570 mL of glacial acetic acid  $\text{CH}_3\text{COOH}$  (99%) to 8000 mL of deionized water. Add 680 mL of concentrated ammonium hydroxide adjust pH to 7.0 with 3.0 N glacial acetic acid or 3.0 N ammonium hydroxide and dilute to 10 L final volume.
3. Ethanol, 95%
4. Hydrochloric acid, 0.1 N HCl - Dilute 8.3 mL of concentrated HCl reagent to 1000 mL with deionized water.
5. Standard calibration solutions of  $\text{NH}_4\text{-N}$ . Prepare six calibration standard solutions of 1.0 - 20  $\text{mg L}^{-1}$  of  $\text{NH}_4\text{-N}$   $\text{mg L}^{-1}$  from 1000  $\text{mg L}^{-1}$  stock solutions. Dilute calibration solutions with 0.1 N HCl solution.

### Procedure:

1. Weigh  $10.0 \pm 0.1$  g of air dry soil pulverized to pass 10 mesh sieve ( $< 2.0$  mm) soil into a 125 mL Erlenmeyer flask. Add 50 mL of ammonium acetate solution (See Comment #1) and place the flask reciprocating shaker for thirty (30) minutes. Include a method blank.
2. Connect a 1 L vacuum extraction flask to a Buchner funnel fitted with a Whatman No. 5 or equivalent filter paper. Moisten the filter paper with 2 mL deionized water (See Comment #2).
3. Transfer the soil suspension into the Buchner funnel and leach the sample with 175 mL of 1 N ammonium acetate. The soil extract may be analyzed for extractable K, Ca, Mg, and Na.

4. Rinse the excess ammonium acetate from the soil sample in the Buchner funnel by leaching with a total volume of ethanol and discard the leachate. Note: Be sure to gently fill funnel to remove all excess ammonium and allow it to drain until only damp soil remains. Continue adding ethanol in this manner until 200 mL of solution has been used.
5. Change to a clean 500-mL suction flask and leach the soil sample with 225 mL of 0.1 N HCl to replace the exchangeable ammonium. Bring leachate to a final volume of 250 mL volumetric flask using deionized water.
6. The concentration of ammonium-N in the final leachate can be determined with an ALPKEM rapid flow analyzer (RF-300), which relies on ammonium to complex with salicylate to form indophenol blue (Technicon Method No. 334-74A/A). This color is intensified with sodium nitroprusside and measured at 660 nm. This determination can also be made using the Kjeldahl distillation method (See Comment #3 and #4).

### Calculation

$$\text{CEC in meq per 100 g of soil} = (\text{mg L}^{-1} \text{ of NH}_4\text{-N in leachate}) \times \frac{0.25}{14} \times \frac{100}{\text{sample size (g)}}$$

mg L<sup>-1</sup> NH<sub>4</sub>-N in leachate is determined using a standard curve.

### Comments

1. Check repipette dispensing volume calibration using an analytical balance.
2. Check filter paper supply for possible contamination of and NH<sub>4</sub>-N. If contamination is greater than 0.2 mg L<sup>-1</sup> on a soil extract basis, rinse filter paper with 0.1 N HCl solution.
3. Samples having ammonium concentrations exceeding the highest standard will require dilution and reanalysis.
4. This procedure used is essentially the same as that of Schollenberger (1945) except that determination of NH<sub>4</sub>-N is done spectrophotometrically rather than Kjeldahl distillation and titration. To determine the NH<sub>4</sub>-N content using the Kjeldahl distillation method, follow steps 1 through 5 above, then proceed to Kjeldahl distillation. Care must be taken not to allow soil to dry and crack between ethanol leaching, as this could result in incomplete removal of excess NH<sub>4</sub>-N. A similar procedure is described by Rhoades (1982).

### Literature

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## Scope and Application

Cation exchange capacity (CEC) is the measure of a soil ability to retain readily exchangeable cations which neutralize the negative charge of soils. This method involves saturation of the cation exchange sites with barium, equilibration, removal of the excess barium with ethanol, replacement and leaching replacement with ammonium. Other cations such as sodium have been used as the exchanging cation with measurement by atomic absorption spectrometer (Rhoades, 1982). This method differs from that using ammonium and is more appropriate for acid soils ( $\text{pH} < 7$ ). For arid soils,  $\text{pH} > 7.5$  and high in carbonates it is recommended to use the method of Polemio and Rhoades (1977). Whatever the replacement cation, the procedure is time consuming. The speed at which samples filter depends on the strength of the vacuum applied and sample makeup. Some samples may never filter because of clogged filter paper. The method detection limit is approximately  $1.0 \text{ cmol kg}^{-1}$  ( $\text{meq}/100 \text{ gm}$ ) (on a dry soil basis) and is generally reproducible within  $\pm 10\%$ .

## Equipment

1. Analytical balance: 250 g capacity, resolution 0.01 g.
2. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute.
3. Repipette dispensers, calibrated to  $20 \pm 0.1$ ,  $8.0 \pm 0.1$  mL and  $5.0 \pm 0.1$  mL.
4. Whatman No. 1, No. 2 or equivalent filter paper.
5. Buchner funnels with vacuum flasks and source of vacuum.
6. Atomic absorption spectrophotometer (AAS) or ICP-AES instrument.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Barium chloride extracting solution 0.5 N: Dissolve 122 g of  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  into deionized water and dilute to 1000 mL.
3. Ethanol, 90%
4. Ammonium acetate (1.0 N) extraction solution neutral: Add 570 mL of glacial acetic acid,  $\text{CH}_3\text{COOH}$  (99%) to 8000 mL of deionized water. Add 680 mL of concentrated ammonium hydroxide adjust pH to 7.0 with 3.0 N glacial acetic acid or 3.0 N ammonium hydroxide and dilute to 10 L final volume.
5. Standard calibration solutions of Barium. Prepare six standard solutions of 5.0 - 100  $\text{mg L}^{-1}$  of Barium prepared from 1000  $\text{mg L}^{-1}$  stock solutions. Dilute solutions with 1.0 N ammonium acetate solution.

## Procedure

1. Weigh  $5.0 \pm 0.1$  g of air dry soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into a 125 mL Erlenmeyer flask. Add 50 mL of 0.5 N Barium chloride solution (See Comment #1 and #2) and place the flask reciprocating shaker for thirty (30) minutes. Include a method blank.
2. Connect a 1 L vacuum extraction flask to a Buchner funnel fitted with a Whatman No. 5 or equivalent filter paper. Moisten the filter paper with 2 mL of deionized water.
3. Transfer the soil suspension into the Buchner funnel and leach the sample with 100 mL of 1 N barium chloride.

4. Rinse the excess barium chloride from the soil sample in the Buchner funnel by leaching with a total volume of ethanol and discard the leachate. Note: Be sure to gently fill funnel to remove all excess barium and allow it to drain until only damp soil remains. Continue adding alcohol in this manner until 200 mL of ethanol has been used.
5. Change to a clean 500-mL suction flask and leach the soil sample with 225 mL of 1.0 ammonium acetate to replace the exchangeable barium. Bring leachate to volume in a 250 mL volumetric flask using deionized water.
6. The concentration of Ba in the final leachate can be determined using atomic absorption or ICP-AES. Record barium concentration to the nearest 10 mg L<sup>-1</sup> (See Comment #3 and #4).

### Calculation

$$\text{CEC in meq per 100 g of soil} = (\text{mg L}^{-1} \text{ Ba in leachate}) \times \frac{0.25}{69} \times \frac{100}{\text{sample size (g)}}$$

### Comments

1. For soils high in salts (EC > 4 dSm<sup>-1</sup>) wash the soil with 30 mL of deionized water to remove excess salts.
2. Check repipette dispensing volume calibration using an analytical balance.
3. Samples having barium concentrations exceeding the highest standard will require dilution and reanalysis.
4. This procedure used is essentially the same as that of Rhoades (1982), except that the barium concentration is increased. For strong acid soils, pH < 6, reduce the concentration of barium to 0.2 N such that the ionic strength matches that of the soil.

### Literature

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Thomas, G.W. 1982. Exchangeable cations. In A.L. Page (ed.). Methods of soil analysis, Part 2, Second Edition, Agronomy Monograph 9, American Society of Agronomy, Madison, WI.

## Scope and Application

This method semiquantitatively determines the amount of sulfate-sulfur ( $\text{SO}_4\text{-S}$ ) in soils by extraction with  $\text{Ca}(\text{H}_2\text{PO}_4)_2 \cdot \text{H}_2\text{O}$  with subsequent determination of  $\text{SO}_4\text{-S}$  by turbidimetric measurement. Calcium phosphate is utilized to suppress the dissolution of organic matter and for the removal of sulfate that maybe absorbed. Turbidimetric analysis is based on the formation of  $\text{BaSO}_4$  crystals in a suspension and subsequent measurement of optical density. The turbidimetric method will require practice to become proficient. It is sensitive to high concentrations of dissolved soil organic carbon which may lead to an under estimation of  $\text{SO}_4\text{-S}$  (Ajwa and Tabatabai, 1993). Sulfate-sulfur can also be determined using ion chromatography using an AS4A column. Soil  $\text{SO}_4\text{-S}$  is only an index of plant available since sulfur is also available from the mineralization of organic matter, irrigation water and atmospheric deposition. The method has a detection limit of  $2.0 \text{ mg kg}^{-1}$  (dry basis) and is generally reproducible to within  $\pm 15\%$ .

## Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute.
3. Erlenmeyer flasks, 125 mL.
4. Magnetic stirrer.
5. Whatman No. 42, S&S 597 or equivalent highly retentive filter paper.
6. Repipette dispenser calibrated to  $25.0 \pm 0.2 \text{ mL}$ .
7. Pipette 10.0 mL.
8. Nephelometer, Turbidimeter or spectrophotometer 420 nm filter.

## Reagents

1. Deionized water, ASTM Type I grade.
2. Calcium phosphate extraction solution, 0.08 M (containing  $500 \text{ mg L}^{-1} \text{ PO}_4\text{-P}$ ): Dissolve 2.03 g of analytical grade  $\text{Ca}(\text{H}_2\text{PO}_4)_2 \cdot \text{H}_2\text{O}$  in 1000 mL of deionized water (see Comment #1).
3. Activated Charcoal (See Comment #2). Mix 100 g of Darco G-6 activated carbon with calcium phosphate extraction solution and thoroughly wet carbon. Cap container, shake, and filter slowly through buchner funnel. Wash three times with deionized water and verify removal of  $\text{SO}_4\text{-S}$ .
4. Acid "seed" solution,  $20 \text{ mg L}^{-1} \text{ S}$  in 5.8 N HCl : Dissolve 0.1087 g analytical grade  $\text{K}_2\text{SO}_4$  in 500 mL of deionized  $\text{H}_2\text{O}$  and add 500 mL of concentrated HCl. Add Teflon coated magnetic stir bar and place on stirrer. Add 5.0 g of powdered gum acacia, or gelatin (See Comment #3) suspension agent slowly add 400 mL of  $40^\circ\text{C}$  deionized water, dissolve lumps and bring to 500 mL. Bring to 1000 mL final volume with acetic acid,  $\text{CH}_3\text{COOH}$  (99%), slowly stirring.
5. Barium chloride crystals. Parr turbidimetric grade,  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  crystals 20 - 30 mesh. Use high purity  $\text{BaCl}_2$ , as low purity may result in low recovery of  $\text{SO}_4\text{-S}$  (See Comment #4).
6. Standard  $\text{SO}_4\text{-S}$  calibration solutions. Prepare  $100 \text{ mg L}^{-1} \text{ SO}_4\text{-S}$  calibration solution, dissolve 0.5434 g of oven dry  $\text{K}_2\text{SO}_4$  in 500 mL of deionized water and dilute to one liter. Prepare six 100 mL calibration solutions of: 0.5, 1.0, 2.0, 4.0, 6.0, and  $8.0 \text{ mg L}^{-1} \text{ SO}_4\text{-S}$  and diluted to final volume with calcium phosphate extraction solution.

## Procedure

1. Weigh  $10.0 \pm 0.1 \text{ g}$  of air dry soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into 125 mL Erlenmeyer extraction flasks (See Comment #5). Add 25 mL of calcium phosphate extraction solution using repipette dispenser and place on reciprocating mechanical shaker for thirty (30)

minutes. Include a method blank.

2. Add 0.15 g of activated charcoal and shake for an additional three (3) minutes. Repeat with 25 mL aliquot of SO<sub>4</sub>-S calibration solutions.
3. Filter extract through paper, refilter if filtrate is cloudy.
4. Place 10.0 mL aliquot of soil extract in 25 mL flask, and add 1.0 mL of seed solution and swirl. Repeat with SO<sub>4</sub>-S calibration solution and method blank (See Comment #6 and #7).
5. Place flask on magnetic stirrer and add 0.3 g of BaCl<sub>2</sub> · 2H<sub>2</sub>O crystals. Stir for five (5) minutes and then read optical density (or percent transmittance) on a nephelometer, turbidimeter or spectrophotometer at 420 nm (See Comment #8). Zero optical density with deionized water. Repeat with SO<sub>4</sub>-S calibration solutions and method blank. Using standard calibration solutions determine SO<sub>4</sub>-S concentration of soil extracts and method blank by plotting log of transmittance versus standard concentration. Record as mg L<sup>-1</sup> of analyte in extract solution to the nearest 0.5 mg L<sup>-1</sup>.

## Calculations

Report soil sulfate-sulfur (SO<sub>4</sub>-S):

$$\text{mg kg}^{-1} = (\text{mg L}^{-1} \text{ SO}_4\text{-S in soil extract} - \text{method blank}) \times (2.5)$$

## Comments

1. For acidic soils (pH < 5.5) add 10 mL of concentrated HCl to extractant solution. Acidified phosphate extraction solutions are more reliable for use on acid soils which may contain absorbed sulfate.
2. For soils containing low concentrations of SO<sub>4</sub>-S labile organic matter may prevent the formation of barium sulfate crystals resulting in a low bias SO<sub>4</sub>-S concentration. Labile organic matter may be removed by the addition of activated charcoal or hydrogen peroxide. For organic soils the volume of extraction (1:5) solution must be increased to account for high potential SO<sub>4</sub>-S concentrations.
3. A number of suspension agents have been reported in the literature which include: gum acacia, gelatin, glycerol, PVP-K30 (polyvinylpyrrolidone), and Tween 80 which have proven effective in turbidimetric analysis. Each of these will require experimentation and practice using SO<sub>4</sub>-S spiking to fully refine the technique. For use of PVP-K30 (polyvinylpyrrolidone) add 10 g to 700 mL and dilute to 1000 mL final volume.
4. Use BaCl<sub>2</sub> specifically designated for turbidimetric determination of sulfate-sulfur. Sources: J.T. Baker Cat. Parr Turbidimetric BaCl<sub>2</sub>, JT0974-5; VWR JT0974-5; and GFS Chemicals Reagent Grade ACS #602.
5. Pre-rinse all extraction flasks, turbidimetric and spectrometer cuvette in hot water followed by 0.5 N HCl rinse.
6. Check repipette volume, calibrate using an analytical balance.
7. HCl is included in the turbidimetric "seed" solution to prevent the precipitation of alkali metals with carbonate and phosphate and provide nucleus for initiating precipitation.
8. Development of BaSO<sub>4</sub> suspension requires continuous mixing and development may require from 1 - 10 minutes. It is essential that time of development be equivalent for both standards and unknown samples.

## Literature

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### Scope and Application

This method involves the quantitative extraction of chloride (Cl) from soils using 0.1 M  $\text{Ca}(\text{NO}_3)_2$ . Chloride in the extract is determined spectrophotometrically by complexation with mercury(II) thiocyanate. The procedure outlined follows that outlined by Fixen, et al. (1988) for determining chloride and is readily adapted to manual or automated techniques. Chloride may also be determined using ion specific electrode using 0.5 M  $\text{K}_2\text{SO}_4$  as the extractant. Care must be taken to avoid Cl contamination from filter paper and operator handling as chloride is a typical contaminate in laboratory operations. Mercury is a hazardous material, follow manufacturers recommendations in handling this material. Soil chloride can be used to predict small grain response to Cl fertilizers. The method detection limit is approximately  $2.0 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible  $\pm 7\%$ .

### Equipment

1. Analytical balance: 1000 g capacity, resolution  $\pm 0.1 \text{ g}$ .
2. Repipette dispenser, calibrated to  $25.0 \pm 0.2 \text{ mL}$ .
3. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute (opm).
4. Extraction vessels and associated filtration vessel (See comment #1).
5. Whatman No. 42 or equivalent highly retentive filter paper.
6. Spectrophotometer, autoanalyzer, or flow injection analyzer (FIA) instrument.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Calcium nitrate extracting solution, 0.1M  $\text{Ca}(\text{NO}_3)_2$ : Dissolve 4.72 g of reagent grade  $\text{Ca}(\text{NO}_3)_2$  into 1200 mL deionized water and dilute to a 2 L in a volumetric flask.
3. Saturated mercury (II) thiocyanate [ $\text{Hg}(\text{SCN})_2$ ] solution. 0.75%: Add approximately 0.75 g  $\text{Hg}(\text{SCN})_2$  to 1 L of distilled water and stir overnight. Filter through Whatman No. 42 paper. Saturated solutions maybe stored for long periods of time.
4. Ferric (III) nitrate nonahydrate [ $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ ] solution, Dissolve 20.2 g of  $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$  in approximately 500 mL of deionized water and add concentrate  $\text{HNO}_3$  until the solution is almost colorless. Dilute to 1 L final volume.
5. Standard calibration solutions of Cl. Prepare six calibration standards ranging from 0.5 to 10.0 mg  $\text{L}^{-1}$  concentration, diluted in calcium nitrate extracting solution.

### Procedure

1. Weigh  $10.0 \pm 0.1 \text{ g}$  of air-dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into extraction vessel. Add 25.0 mL of calcium nitrate extracting solution, 0.1  $\text{Ca}(\text{NO}_3)_2$  using repipette dispenser (See Comment #2). Include a method blank.
2. Place extraction vessel(s) on reciprocating horizontal mechanical shaker for fifteen (15) minutes.
3. Filter extract (See Comment #3), refilter if filtrate is cloudy.
4. Chloride content of the extract is determined using a spectrophotometer, automated flow analyzer or FIA instrument. Calibrate using standard calibration solutions and operate instrument in accordance with manufacturer instructions. Determine chloride concentration of 0.1M  $\text{Ca}(\text{NO}_3)_2$  extract, method blank, unknown samples and record results as mg  $\text{L}^{-1}$  of chloride in extract solution (See Comment #4 and #5).

## Calculation

Cl mg kg<sup>-1</sup> in soil = (Cl mg L<sup>-1</sup> in filtrate - method blank) × 2.5

Report soil chloride concentration to the nearest 0.1 mg kg<sup>-1</sup> (See Comment #6)

## Comments

1. Rinse all extraction labware with 0.1M Ca(NO<sub>3</sub>)<sub>2</sub> to minimize Cl contamination of unknown samples and standards.
2. Check repipette dispensing volume calibration using an analytical balance.
3. Check filter paper supply for possible contamination of and Cl. If contamination is greater than 0.5 mg L<sup>-1</sup> on a solution basis, rinse filter paper with 0.1M Ca(NO<sub>3</sub>)<sub>2</sub>.
4. Samples having chloride concentrations exceeding the highest standard will require dilution and reanalysis.
5. Chloride can also be extracted with 0.5 M K<sub>2</sub>SO<sub>4</sub> and determined using potentiometric known addition methods or extracted with 0.5 M NaHCO<sub>3</sub> with subsequent analysis using ion exchange chromatography.
6. Chloride (Cl) results can be expressed on a volume basis. Assuming the sample represents a 0-6 inch (0-15 cm) depth of the soil, then: Cl mg kg<sup>-1</sup> × 2.0 ≈ Cl lbs ac<sup>-1</sup>

## Literature

Fixen, P.E., R.H. Gelderman and J.L. Denning. 1988. Chloride Tests. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bulletin No. 499 (revised).

Gelderman, R.H and P.E. Fixen. 1988. Recommended nitrate-N tests. p. 10-11. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bulletin No. 499 (revised).

Keeney, D.R. and D.W. Nelson. 1982. Nitrogen - inorganic forms. p. 643-698. *In* A.L. Page (ed.) Methods of soil analysis, part 2. Agron. Monogr. 9, 2nd ed. ASA and SSSA, Madison, WI.

### **Scope and Application**

This method involves tests for the presence of carbonates in soil materials and is also maybe known as the “fizz” test. The method is based on the reaction of HCl with soil carbonates and visual observation of gaseous loss of CO<sub>2</sub> from the sample as described by the U.S. Salinity Laboratory Staff (1954). The method is not quantitative. Soils may be categorized as slightly reactive, moderately reactive or highly reactive. The method detection limit is approximately 0.2% CaCO<sub>3</sub> equivalent (on a dry soil basis).

### **Equipment**

1. 50 mm watchglass.
2. Pipette dispenser, calibrated to 2.0 ± 0.2 mL.

### **Reagents**

1. Deionized water, ASTM Type I grade.
2. Hydrochloric acid (HCl), 3 N. Transfer 250 mL of concentrated HCl to 500 mL of deionized water and dilute to 1.0 L with deionized water.

### **Procedure**

1. Place 4 - 6 g of soil on a small watchglass. Using a pipette add sufficient water to saturate the sample (See Comment #1). Using a pipette add a few drops of 3.0 N hydrochloric acid solution. Note any effervescence that occurs. The soil may be classified as slightly, moderately or highly calcareous in accordance to the degree of effervescence.

### **Comments**

1. Water is added to the sample to displace soil air as to avoid confusion of the loss of soil air with effervescence of lime.

### **Literature**

Allison L.E. and C.D. Moodie. 1965. Carbonate. P. 1379-1400. *In*: C.A. Black et al. (Ed.) Methods of soil analysis. Part 2. 2nd ed. Agron. Monogr. 9. ASA, CSSA and SSSA, Madison, WI.

Sobeck, A.A., W. A. Schuller, J.R. Freeman, R. M. Smith, 1978 Field and Laboratory Methods to Overburden and Minesoils. U.S. Department of Commerce, National Technical Information Service.

U.S. Salinity Lab. Staff. 1954. Methods for soil characterization. p. 83-147. *In*: Diagnosis and improvement of saline and alkali soils. Agr. Handbook 60, USDA, Washington, D.C.

## Scope and Application

This method involves the quantitative determination of calcium carbonate by gravimetric analysis. The method is based on the reaction of HCl with calcium carbonate and the gravimetric loss of CO<sub>2</sub> from the sample as described by the U.S. Salinity Laboratory Staff (1954). Major sources of error are: evaporation of water and failure to adequately degas CO<sub>2</sub> from the sample. Soil carbonates are measured to determine soil buffering capacity with relation to soil fertility, chemical and pedogenic processes. The method detection limit is approximately 0.2% CaCO<sub>3</sub> equivalent (on a dry soil basis) and is generally reproducible ±10%.

## Equipment

1. Analytical balance: 100 g capacity, resolution ± 0.1 mg.
2. Repipette dispenser, calibrated to 25.0 ± 0.2 mL.
3. Orbital mechanical shaker, capable of 180 oscillations per minute (opm).
4. Extraction vials, 70 mL with snap-lids having 1 mm holes for gas exchange or 50 mL Erlenmeyer flask with cap (See comment #1).

## Reagents

1. Deionized water, ASTM Type I grade.
2. Hydrochloric acid (HCl), 3.0 M. Transfer 250 mL of concentrated HCl to 500 mL of deionized water and dilute to 1.0 L with deionized water.
3. Calcium carbonate (CaCO<sub>3</sub>), fine ground (100 or 140 mesh sieve, 106-150 μm), reagent grade.

## Procedure

1. Weigh to the nearest 0.1 mg a 70 mL extraction vial containing 10 mL of 3.0 M HCl and record tare weight. Transfer 2.000 to 5.000 g of air-dried soil pulverized to pass 10 mesh sieve (< 2.0 mm) containing 100 to 300 mg of CaCO<sub>3</sub> equivalent, in incremental units to avoid frothing. Accurately record the weight of soil transferred to the nearest 0.1 mg. Include three calcium carbonate standards ranging from 100 to 300 mg. After the effervescence has subsided replace snap-lid and place on orbital shaker for fifteen (15) minutes. Include 3 blanks to determine water vapor loss.
2. After two (2) hours weigh vial to the nearest 0.1 mg and record the mass. (See Comment #3 and #4). Verify recovery of calcium carbonate standards, approximately 100%.

## Calculation

Weight loss of CO<sub>2</sub> (g) = Initial weight (g) - final weight (g) (vial +stopper +acid + soil)

$$\text{CO}_3\text{-C, \%} = \left( \frac{\text{g CO}_2 \text{ lost}}{\text{g air-dry soil}} \right) (0.2727) \times 100 \quad [\text{equ. S -13.1-1}]$$

$$\text{CaCO}_3\text{-C, \%} = \left( \frac{\text{g CO}_2 \text{ lost}}{\text{g air-dry soil}} \right) (2.273) \times 100 \quad [\text{equ. S -13.1-2}]$$

## Comments

1. Vials snap-lids should be large enough to permit gas exchange of CO<sub>2</sub>, yet small enough to minimize loss of water vapor.
2. Soil weight should be less than 4.0 g for soils with less than 20% CaCO<sub>3</sub>, 2.0 g for soils 20 to 40% CaCO<sub>3</sub>.
3. Use blank subtraction to compensate for water vapor loss. If water vapor loss is > 0.003 g subtract water vapor loss (g) from g CO<sub>2</sub> weight loss.
4. To convert from CaCO<sub>3</sub>-C, % to total inorganic carbon (TIC) multiply value by 0.12 .

## Literature

Allison L.E. and C.D. Moodie. 1965. Carbonate. P. 1379-1400. *In*: C.A. Black et al. (Ed.) Methods of soil analysis. Part 2. 2nd ed. Agron. Monogr. 9. ASA, CSSA and SSSA, Madison, WI.

Loeppert, Richard H. And Donald L. Suarez. 1996. Carbonate and gypsum. p. 437-474. *In*: J.M. Bartels et al. (ed.) Methods of soil analysis: Part 3 Chemical methods. (3rd.ed.) ASA and SSSA, Madison, WI. Book Series no. 5

U.S. Salinity Lab. Staff. 1954. Methods for soil characterization. p. 83-147. *In*: Diagnosis and improvement of saline and alkali soils. Agr. Handbook 60, USDA, Washington, D.C.

### Scope and Application

This method involves the quantitative determination of inorganic carbon by volumetric displacement. The method is based on the reaction of HCl with carbonate and the measurement of the loss of CO<sub>2</sub> based on the equations described by Loeppert and Suarez (1996) and by Wagner et al., (1998). The method range is from 0.25 to 100% percent CaCO<sub>3</sub> for a 20 mL serum bottle used as the reaction vessel. The method range for the 100 mL serum vial used as the reaction vessel is 2.0 to 100 percent. Soil inorganic carbon is measured to determine soil buffering capacity with relation to soil fertility, chemical and pedogenic processes and to obtain organic carbon from combustion methods that produce total carbon. The method detection limit is approximately 0.25 % CaCO<sub>3</sub> and is reproducible within the ± 5% using the 20 mL serum bottles and 2.0% for the 100 mL serum bottles.

### Equipment

1. Analytical balance: 100 g capacity, resolution ± 0.1 mg.
2. Repipette dispenser, calibrated to 2.0 mL.
3. Reaction vessels, 100 mL capacity wheaton serum bottles.
4. 0.5 dram vials (2.0 mL capacity).
5. Gray butyl rubber stoppers.(See comment #1).
6. Tear-off Aluminum serum bottle seals.
7. Hand-operated crimpers.
8. Power supply (24 volt DC. - 2 amp).
9. Digital voltage meter, capable of reading 0.01 volts resolution.
10. Pressure transducer 0-105 kPa (Setra Model 280E).

### Apparatus

1. The pressure transducer is connected to the power supply with 14 gauge wire with a digital volt meter wired in line to monitor the output.
2. Attached to the base of the pressure transducer is 20 cm of tubing attached to a 18 gauge Luer-loc hypodermic needle with a particle filter in the middle to collect any reflux from reaching the pressure transducer.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Calcium carbonate (CaCO<sub>3</sub>), fine ground (100 mesh sieve, 150 um), reagent grade.
3. Hydrochloric acid (HCl) 6N with 3% by weight ferrous chloride). Transfer 500 mL of HCl to 400 mL of deionized water and add 30 g of FeCl<sub>2</sub> · 4H<sub>2</sub>O and bring to 1.0 L volume with deionized water.
4. Prepare CaCO<sub>3</sub> standards of 0.25, 0.50, 1, 2, 4, 8, and 15 percent by weight using laboratory sand that has been powder ground and oven dried reagent grade CaCO<sub>3</sub> for 20 mL reaction vessels for the 0-30 percent range. For soils higher than 30% CaCO<sub>3</sub>, use 100 mL reaction vessels with standard concentrations of 10, 20, 30, 40, 50 and 80% CaCO<sub>3</sub> .

## Procedure

1. Weigh 1.00 g of soil into a 20 mL Wheaton serum bottle if soil is expected to have less than 30% CaCO<sub>3</sub> or 2.00g of soil into 100 mL Wheaton serum bottle if sample is expected to have more than 30% CaCO<sub>3</sub>. Use soil screened to pass through a 2mm sieve. Place CaCO<sub>3</sub> appropriate standards into 20 mL or the 100 mL Wheaton serum bottle (See Comment #1).
2. Pipet 2.0 mL of 6.0N HCl reagent into 0.50 dram glass vial. Gently insert acid dram vial into reaction vessel with sample, but do not allow solution to mix with sample. Cap reaction vessel with gray butyl rubber stoppers and crimp with aluminum tear-off seals.
3. Shake reaction vessel vigorously to ensure that acid solution in the dram vial has mixed with the soil. Run three blanks (1.00 g of laboratory sand with acid vial) with each analysis run.
4. Prior to reading the samples on the pressure transducer (15 min), rotate the acid along the sides of the reaction vessel to ensure that soil on the sides is reacting with the acid.
5. After two (2) hours of reaction time with the acid the samples and standards are ready to read on the pressure transducer.
6. Record the voltage output to 2 decimal places. Subtract the average voltage of the blanks from the standards and samples to obtain the change in pressure due to CO<sub>2</sub>.

## Calculation

Using linear regression determine the slope (regression coefficient) and the intercept (b) of the curve of pressure change vs. the dependent variable of percent CaCO<sub>3</sub>. Inorganic carbon can be obtained by dividing the formula weight of CaCO<sub>3</sub> (100) by the formula weight of carbon (12) and multiply this by the % CaCO<sub>3</sub>.

$$\% \text{ CaCO}_3 = (\text{regression coefficient}) \times (\text{delta pressure in volts}) + b.$$

$$\% \text{ Inorganic Carbon} = \% \text{ CaCO}_3 / 8.33 \quad [\text{equ. S -13.3-1}]$$

$$\text{Inorganic carbon g/kg} = 10 \times (\text{Inorganic Carbon percent})$$

## Comments

1. Caution should be used when there is no information available on CaCO<sub>3</sub> content of the soil as vessels may become over pressurized. Use 1.00 g of soil and the 100 mL Wheaton serum bottles as a pre-screen of CaCO<sub>3</sub> content prior to quantitative determination using the appropriate mass and bottle. Sample size maybe increased to 2.0 g using the 20 mL Wheaton serum bottle for samples containing less than 15% CaCO<sub>3</sub> to improve method detection and precision.
2. Soils containing dolomite need longer reaction times with acid.
3. Use caution when shaking the reaction vessel (serum bottle containing sample and acid vial) as cracks in the serum bottle can occur with reuse which can weaken the glass if excessive pressure's are encountered.
4. Check 20 and 100 mL serum reaction vessels for irregularities. Discard if irregularities in glass is found.

## Literature

Wagner, S.W., J.D. Hanson, Alan Olness, and W.B. Voorhees. 1998. A volumetric inorganic carbon analysis system. Soil Sci. Soc. Am. J. 62:690-693.

Loeppert, Richard H. and Donald L. Suarez. 1996. Carbonate and gypsum. p. 437-474. In: J.M. Bartels et al. (ed.) Methods of soil analysis: Part 3 Chemical methods. (3rd.ed.) ASA and SSSA, Madison, WI. Book series no. 5.

### Scope and Application

This method quantitatively determines the physical proportions of three sizes of primary soil particles as determined by their settling rates in a aqueous solution using a hydrometer. Proportions are represented by stated class sizes: sand ranging from 2000 - 50  $\mu\text{m}$ ; silt ranging from 50 - 2.0  $\mu\text{m}$  and clay < 2.0  $\mu\text{m}$  and those stated by the USDA Soil Survey and Canadian Soil Survey Committee. Settling rates of primary particles are based on the principle of sedimentation as described by Stokes' Law and measured using a hydrometer. The use of the ASTM 152H-Type hydrometer is based on a standard temperature of 20 °C and a particle density of 2.65 g cm<sup>-3</sup> and units are expressed as grams of soil per liter. For specific samples the method may require the pretreatment removal of soluble salts, organic matter, carbonates and iron oxides with subsequent dispersion using sodium hexametaphosphate (Day 1965). Corrections for temperature and for solution viscosity is made by taking a hydrometer reading of a blank solution. For further information consult Gee and Bauder (1986). Generally this method is of lower precision than the pipette or sedimentation methods and is used to determine soil texture. The method has a detection limit of 2.0% sand, silt and clay (dry basis) and is generally reproducible to within  $\pm 8\%$ .

### Equipment

1. Analytical balance: 100 g capacity, resolution  $\pm 0.01$  g.
2. Standard hydrometer, ASTM No. 1. 152H-Type with Bouyoucos scale in g L<sup>-1</sup>.
3. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute.
4. Sedimentation cylinder with 1.0 L mark 36 $\pm$  2 cm from the bottom.
5. Shaker bottle 200 mL and cap (polypropylene or glass).

### Reagents

1. Deionized water, ASTM Type I grade.
2. Amyl alcohol.
3. Sodium Hexametaphosphate (HMP), 5% dispersing solution. Dissolve 50.0 g Na-hexametaphosphate in 1.0 L.

### Procedure

1. Weigh 40.0  $\pm$  0.05 g of air-dried soil pulverized to pass 10 mesh sieve (< 2.0 mm) into 200 mL container (See Comments #1, #2, #3 and #4). Determine oven dry soil moisture on a 2nd sample of soil.
2. Add 100 mL of HMP solution, cap and place on reciprocating horizontal shaker for sixteen (16) hours (See Comment #5).
3. Quantitatively transfer the suspension to the sedimentation cylinder and add deionized water to bring to 1.0 L final volume.
4. Allow the suspension to equilibrate to room temperature for two (2) hours.
5. Insert plunger and thoroughly mix contents, dislodging sediment from the bottom of the cylinder. Finish stirring with two or three smooth stokes. As an alternative mixing procedure stopper the cylinder and used end over end shaking for one (1) minute. Add 2 mL of amyl alcohol to the surface to suspensions covered in foam. Repeat the process and determine hydrometer reading on a blank solution and to the nearest  $\pm 0.5$  g L<sup>-1</sup> as  $R_{C1}$ .
6. Lower the hydrometer carefully into the suspension after thirty (30) seconds and take a reading after forty (40) seconds and record to the nearest  $\pm 0.5$  g L<sup>-1</sup> as  $R_{\text{sand}}$  (See Comment #6).
7. Remove the hydrometer carefully, rinse and wipe dry.
8. After six (6) hours record temperature of the suspension to the nearest  $\pm 1^\circ\text{C}$ . Using the temperature correction values in Table 14.1-A determine the settling time for the 2.0  $\mu\text{m}$  size fraction (See Comment #7). Based on time after initiation of settling, reinsert the hydrometer carefully and take a reading and record as  $R_{\text{clay}}$  to the nearest  $\pm 0.5$  g L<sup>-1</sup>. Repeat the process determining hydrometer



reading on a blank solution and record as  $R_{C2}$  to the nearest  $\pm 0.5 \text{ g L}^{-1}$  (See Comments #8).  
 Table 14.10-A The influence of suspension temperature on the hydrometer determination of soil clay ( $2 \mu\text{m}$ ) based on a particle density of  $2.65 \text{ g cm}^{-3}$  and a solution density of  $0.5 \text{ g L}^{-1}$ .

Temperature °C	Settling time for clay hours and minutes
18	8:09
19	7:57
20	7:45
21	7:35
22	7:24
23	7:13
24	7:03
25	6:53
26	6:44
27	6:35
28	6:27

Gee and Bauder (1986).

### Calculations

Report results to the nearest 0.1% content (See Comment #9 and 10):

$$\text{Sand \%} = \frac{(\text{oven dry soil mass}) - (R_{\text{sand}} - R_{C1})}{(\text{oven dry soil mass})} \times 100 \quad [\text{equ. S -14.1-1}]$$

$$\text{Clay \%} = \frac{(R_{\text{clay}} - R_{C2})}{(\text{oven dry soil mass})} \times 100 \quad [\text{equ.S -14.1-2}]$$

$$\text{Silt \%} = 100 - (\text{Sand \%} + \text{Clay \%}) \quad [\text{equ. S -14.1-3}]$$

### Comments

1. The exact sample size is soil texture dependent. For fine textured soils, silts or clays, 10 - 20 g may be adequate. For coarse textured soils 60 - 100 g will be needed in order to obtain reproducible results. For moist soils dry overnight at  $105 \text{ }^\circ\text{C}$  and correct for moisture content.
2. For soils containing carbonates ( $\text{CaCO}_3 > 2.0\%$ , see Methods 13.1 or 13.2) and/or high in soluble salts ( $\text{EC}_e > 2.0 \text{ dS m}^{-1}$ ) it is recommended soils be pre-treated. Place 40.0 g of soil in 250 mL centrifuge tube, add 100 mL deionized water and 10.0 mL of 1.0 M Na acetate (pH 5.0). Mix, and centrifuge for 10 min at 1500 rpm) until the supernatant is clear. Decant and wash two more times with 50 mL of deionized water.
3. For soils containing organic matter contents greater than 3.5%, after removal of carbonates, add 25 mL of water and add 5 mL of  $\text{H}_2\text{O}_2$  to the suspension. If excessive frothing occurs, cool and add additional

H<sub>2</sub>O<sub>2</sub> when reaction subsides. Heat to 90 °C when frothing ceases. Continue treatment until organic matter is oxidized (as judged by rate of reaction and bleached color).

4. For removal of Iron oxides add 150 mL to the H<sub>2</sub>O<sub>2</sub> treated sample of a solution 0.3 M sodium citrate and 84 g/L sodium bicarbonate. Shake for 30 minutes to disperse the soil and add 3.0 g of sodium dithionite (Na<sub>2</sub>S<sub>2</sub>O<sub>4</sub>). Place in water bath 80 °C and stir intermittently for 20 minutes. Remove and add 10 mL of a 10% NaCl solution, centrifuge and decant. If sample is brownish in color repeat with the sodium citrate - sodium bicarbonate step. If sample is gleyed (gray), repeat with 10% solution of NaCl, and two deionized water rinses. Proceed with HMP addition.
5. It is recommended to use horizontal reciprocating shaking for the dispersing the samples. The use of electric stirrer at high rpm may result in significant grinding of sample primary minerals.
6. It may be advantageous to take a reading at 35 seconds and a 2nd at 45 seconds and record the average as the hydrometer value is dynamic.
7. The hydrometer procedure has been used by number of laboratories. Readings are frequently made at 40 seconds and two (2) hours for the determination of sand and clay. Based on theoretical considerations the two (2) hour reading is a estimate of the 5.0  $\mu$ m fraction and errors often exceed 10% by weight from that of actual clay (Gee and Bauder, 1979).
8. For determining sand fractions, quantitatively transfer the sediment suspension through a 270 mesh (53.0  $\mu$ m) sieve and wash with deionized water using a wash bottle. Transfer the sand to a tared beaker, dry at 105 °C and weigh. The dried sand may be placed in nested sieves to determine individual sand fraction size analysis.
9. For soils having clay particle densities less than 2.65 g cm<sup>-3</sup>, settling time will increase and for soils greater than 2.65 g cm<sup>-3</sup> it will decrease, consult Gee and Bauder (1986).
10. An error of 1.0  $\pm$  g L<sup>-1</sup> in the hydrometer reading results in an error of  $\pm$  2.5% for clay size Fraction and on a percentage basis translates to an error of  $\pm$ 12.5% for 40 g of soil containing 20% clay. For more accurate and precise measurement of silt and clay size fractions it recommended that the pipette method be used which has a precision of  $\pm$ 2-3%.
11. Using soil size particle analysis data, soil textural classification can be determined (Appendix E). From texture an approximate bulk density can be determined (Saxton, et al. 1986).

## Literature

American Society for Testing and Materials. 1985d Standard test method for particle-size analysis of soils D 422-63 (1972). 1985 Annual Book of ASTM Standards 04.08:117-127. American Society for Testing Materials, Philadelphia.

Day, P.R. 1965. Particle fractionation and particle-size analysis p. 545-567. *In* C.A. Black et al (ed.) Methods of soil analysis, Part 1. Agronomy 9:545-567.

Gee, G. W., and J. W. Bauder. 1979. Particle size analysis by hydrometer: a simplified method for routine textural analysis and a sensitivity test of measured parameters. *Soil Sci Soc. Am. J.* 43:1004-1007.

Gee, G. W., and J. W. Bauder. 1986. Particle-size Analysis. P. 383 - 411. *In* A.L. Page (ed.). Methods of soil analysis, Part 1, Physical and mineralogical methods. Second Edition, Agronomy Monograph 9, American Society of Agronomy, Madison, WI.

Saxton, K.E., W.J. Rawls, J.S. Romberger, and R.I. Papendick. 1986. Estimating generalized soil-water characteristics from texture. *Soil Sci. Soc. Am. J.* 50(4):1031-1036.

U.S. Salinity Lab. Staff. 1954. Methods for soil characterization. p. 83-147. Diagnosis and improvement of saline and alkali soils. Agr. Handbook 60, USDA, Washington, D.C.

### Scope and Application

This method involves the semiquantitative extraction of exchangeable aluminum from soils using 2.0 N KCl. Aluminum is determined by Inductively Coupled Plasma spectrometry (ICP-AES). The method doesn't quantitatively extract aluminum from mineral structures or bound to organic compounds. Care must be taken to avoid contamination from filter paper and operator handling. Soil aluminum concentrations are generally low in mineral soils ( $< 1.0 \text{ mg kg}^{-1}$ ), with the exception of soils with a soil:water (1:1) pH less than 5.40 (Method S - 2.30). The method detection limit is approximately  $0.5 \text{ mg kg}^{-1}$  (on a dry soil basis) and is generally reproducible  $\pm 10\%$ .

### Equipment

1. Analytical balance: 1000 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. Repipette dispenser, calibrated to  $25.0 \pm 0.2 \text{ mL}$ .
3. Reciprocating horizontal mechanical shaker, 180 oscillations per minute (opm).
4. Extraction vessels and associated filtration vessel.
5. Whatman No. 42 or equivalent highly retentive filter paper.
6. Inductively Coupled Plasma Atomic Emission Spectrometry (ICP-AES).

### Reagents

1. Deionized water, ASTM Type I grade.
2. Potassium chloride extracting solution, 1.0 N KCl: Dissolve 75 g of reagent grade KCl in 500 mL deionized water and dilute to a 1000 mL (See Comment #1).
3. Standard calibration solutions of  $\text{NH}_4\text{-N}$ . Prepare five calibration standards ranging from 0.1 to 50.0  $\text{mg L}^{-1}$  concentration, diluted in 1.0 N KCl extraction solution prepared from 1000  $\text{mg L}^{-1}$  Al standard solution.

### Procedure

1. Weigh  $5.0 \pm 0.05 \text{ g}$  of air-dried soil pulverized to pass 10 mesh sieve ( $< 2.0 \text{ mm}$ ) into extraction vessel. Add 25.0 mL of 1.0 N KCl extraction reagent using repipette dispenser (See Comment #2). Include a method blank.
2. Place extraction vessel(s) on reciprocating mechanical shaker for thirty (30) minutes.
3. Filter extract (See Comment #3), refilter if filtrate is cloudy (comment #4).
4. Aluminum content of the extract is determined using a Inductively Coupled Plasma Atomic Emission Spectrometry (ICP-AES). Suggested wavelengths are: 309.271, 396.152 and 237.335 nm. Adjust and operate instrument in accordance with manufacturer's instructions. Determine aluminum concentration of a method blank reference checks and unknown samples (See Comment #5).

### Calculation

$$\text{Al mg kg}^{-1} \text{ in soil} = (\text{Al mg L}^{-1} \text{ in filtrate} - \text{method blank}) \times 5$$

Report soil Al concentration to the nearest  $0.1 \text{ mg kg}^{-1}$  (See Comment #5)

## Comments

1. Soils may be extracted with 1.0 N KCl for the simultaneous determination of nitrate (Method 3.10).
2. Check repipette dispensing volume calibration using an analytical balance.
3. Check filter paper supply for possible contamination of and Al. If contamination is greater than 0.2 mg L<sup>-1</sup> on a solution basis, rinse filter paper with 1.0 N KCl.
4. It is recommended that soils extracted for aluminum be analyzed within two (2) hours after extraction.
5. Samples having aluminum concentrations exceeding the highest standard will require dilution and reanalysis.

## Literature

Bertsch, P.M. and P.R. Bloom. 1996. Aluminum. p. 517-550. *In*: J.M. Bartels et al. (ed.) Methods of soil analysis: Part 3 Chemical methods. (3rd.ed.) ASA and SSSA, Madison, WI. Book series no. 5

Keeney, D.R. and D.W. Nelson. 1982. Nitrogen - inorganic forms. *In* A.L. Page (eds.) Methods of soil analysis, part 2. Agron. Monogr. 9, 2nd ed. ASA and SSSA, Madison, WI. p. 643-698.

# **BOTANICAL ANALYSIS METHODOLOGIES**

## DETERMINATION OF DRY MATTER CONTENT OF BOTANICAL MATERIALS B - 1.10

### Gravimetric Moisture

#### Scope and Application

This method quantitatively determines the dry matter percentage in botanical materials based on the gravimetric loss of free water associated with heating to 105 °C for a period of two hours. The method is destructive with respect to the sample. Dry matter fraction is used to correct the sample element concentration to an absolute dry matter basis. The method does not remove molecular bound water and is generally reproducible within ± 37%.

#### Equipment

1. Analytical balance: 250 g capacity, resolution ±0.001 g.
2. Aluminum weight dish with handle.
3. Drying oven, preheated to 105 °C.
4. Desiccator, containing a desiccating agent.

#### Procedure

1. Weigh approximately 2 g of air dry botanical sample material into a tared aluminum weigh pan (preweighed to nearest 0.001 g) and record moist sample weight to the nearest 0.001 g.
2. Place sample and weigh pan in drying oven for a minimum of two (2) hours.
3. Remove and place pan in desiccator for one (1) hour.
4. Weigh sample and pan on balance, weigh and record mass to nearest 0.001 g.
5. Dispose of sample (see comment #1).

#### Calculation

$$\text{Sample dry matter \%} = \left( 1 - \frac{(\text{Sample moist wt.}) - (\text{sample dry wt.} - \text{pan tare wt.})}{(\text{Sample dry weight} - \text{pan tared weight})} \right) \times 100$$

Report dry matter content to the nearest 0.1 %.

#### Comment

1. Drying samples at 105 °C may volatilize some carbon, nitrogen and sulfur compounds. Therefore, material used for moisture content should not be used for inorganic analysis.

#### Literature

Reuter, D.J., J.B. Robinson, K.I. Peverill and G.H. Price. 1986. Guidelines for collecting, handling and analyzing plant materials. p. 11-35 *In*: D.J. Reuter and J.B. Robinson (ed.) Plant analysis an interpretation manual. Inkata Press, Melbourne, Australia.

Smith, Morris and J. Benton Storey. 1976. The influence of washing procedure on surface removal and leaching of certain elements from trees. *Hort. Sci.* 14:718-719.

# TOTAL KJELDAHL NITROGEN DETERMINATION IN BOTANICAL MATERIALS B - 2.10

## Micro-Kjeldahl

### Scope and Application

The Kjeldahl method quantitatively determines the amount of nitrogen (ammonium and protein) in botanical materials based on the wet oxidation of organic matter using sulfuric acid and digestion catalyst and conversion of nitrogen to ammonium (Issac and Johnson, 1976). Ammonium may be determined by distillation into boric acid and titration (Jones, 1989); spectrophotometric measurement (automated or manual); or diffusion-conductivity (Carlson, 1978). The method does not quantitatively recover nitrogen from heterocyclic rings (such as nicotinic acid) or from oxidized forms such as nitrate and nitrite. The Kjeldahl digest can be used for the determination of plant total phosphorus. The method is used to assess plant nitrogen sufficiency levels (Chapman and Pratt, 1961). The method detection limit is approximately 0.05% nitrogen (dry sample basis) and is generally reproducible within  $\pm 8\%$ .

### Equipment

1. Analytical balance: 100 g capacity, resolution  $\pm 0.1$  mg.
2. Acid fume hood and digestion heating block (400 °C).
3. Volumetric digestion tubes, 75 mL.
4. Repipette dispenser, calibrated  $3.0 \pm 0.1$  mL.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Digest catalyst accelerator: prepared by mixing (100:10:1) 100 g potassium sulfate ( $K_2SO_4$ ), 10 g anhydrous copper sulfate ( $CuSO_4$ ), and 1.0 g selenium (Se) metal powder. This can be purchased as a prepared material under the brand name Kjel-tab, distributed by various chemical suppliers.
3. Concentrated sulfuric acid ( $H_2SO_4$ ), reagent grade.
4. 30% hydrogen peroxide ( $H_2O_2$ ); use fresh, as this material rapidly decomposes.
5. Standard calibration solutions of  $NH_4-N$ . Prepare six calibration standards ranging from 0.2 to 40.0 mg  $L^{-1}$  concentration, diluted with 4% (v/v) sulfuric acid, prepared from 1000 mg  $L^{-1}$  ammonium nitrogen standard solution.

### Procedure

1. Determine the moisture content of the botanical material on a sub sample, Method P - 1.10.
2. Weigh  $250.0 \pm 5.0$  mg of air dried botanical material (See Comment #1) and place in into a 75 mL volumetric digestion tube (50 ml or 100 mL digestion tubes may be substituted). Include a method blank.
3. Add Kjel-tab and 6.0 mL of concentrated sulfuric acid (See Comments #2 and #3).
4. Mix on a vortex stirrer fifteen (15) seconds to thoroughly wet the sample with acid. Note: it is essential that all dry sample material be completely moistened by acid and well mixed to insure complete digestion.
5. Place the digestion tube on a digestion block, preheated to 370 °C for thirty (30) seconds or long enough to achieve complete botanical material breakup.
6. Remove from the digestion block and carefully (*slowly*) add 2-5 mL of 30% hydrogen peroxide in 1 mL increments to each digestion tube until digests begin to clear. Because this reaction takes place very rapidly, slow additions avoid excessive foaming.
7. Place the digestion tube back on the digestion block maintained at 370°C for two (2) hours. If excessive foaming occurs, remove from heat, cool two (2) minutes and add an additional 1-2 mL of hydrogen peroxide. At completion, a blue-green color may persist.



8. Remove samples from block and leave under fume hood for 5-10 minutes. Then add 10-20 mL of deionized water using a wash bottle to each tube to prevent hardening and crystal formation. Dilute digestion tubes to volume with deionized water, cap, and invert three times.
9. Sample digests can be analyzed for ammonium nitrogen by three standard methods: They are conventional ammonium distillation into boric acid and titration (Jones, 1989); spectrophotometric determination of ammonium (automated or manual); or diffusion-conductivity method of Carlson (1978). Determine ammonium concentration of a method blank, unknown samples and record results as  $\text{mg L}^{-1}$  of  $\text{NH}_4\text{-N}$  in the digest (See Comment #4 and #5).

### Calculations

Report total Kjeldahl nitrogen results to the nearest 0.01%:

$$\% \text{ N} = \frac{(\text{mg L}^{-1} \text{ NH}_4\text{-N in digest} - \text{method blank}) \times (0.075) \times (100)}{(\text{Sample size, mg}) \times (\text{Dry Matter Content (\%)} / 100)}$$

### Comments

1. Use 500 mg of sample if nitrogen content is less than 0.800%.
2. Check repipette dispenser delivery volume, recalibrate using an analytical balance.
3. When adding reagent to vessels and handling digests always wear protective clothing (i.e. eye protection, lab coat, disposable gloves and shoes). Always handle reagents and opening of vessels in an acid hood capable of high air flow, 100 cfm.
4. Samples having  $\text{NH}_4\text{-N}$  concentrations exceeding the highest standard will require dilution and reanalysis.
5. Sulfuric acid digest containing selenium is classified as a hazardous waste and must be disposed of in a suitable manner.

### Literature

Carlson, R.M. 1978. Automated separation and conductimetric determination of ammonia and dissolved carbon dioxide. *Anal. Chem.* 48:1528-1531.

Carlson, R.M., R.I. Cabrera, J.L. Paul, J. Quick, and R.Y. Evans. 1990. Rapid direct measurement of ammonium and nitrate in soil and plant tissue extracts. *Comm. in Soil Sci. Plant Anal.* 21:1519-1529.

Chapman, H.D. and P.F. Pratt. 1961. *Methods of analysis for soils, plants and waters*, Priced Publication 4034. Berkeley: University of California, Division of Agriculture Sciences.

Issac, R.A. and W.C. Johnson. 1976. Determination of total nitrogen in plant tissue. *J. of Assoc. of Off. Anal. Chem.* 59:98-100.

Jones, J.B. 1989. *Plant analysis techniques*. Athens, GA: Benton Laboratories, Inc. p. 16-18.

### Scope and Application

This method quantitatively determines the amount of nitrogen in all forms (ammonium, nitrate, protein and heterocyclic nitrogen such as nicotinic acid) in botanical materials using a resistance furnace and a thermal conductivity detector. It is based on the method originally described by Dumas and later modified by Sweeny (1989) whereby botanical samples, encased in tin (Sn) foil or another material, are ignited in a resistance furnace at approximately 900 °C, in helium and oxygen environment in a quartz combustion tube. An aliquot of the combustion gas is passed through a copper catalyst to remove oxygen and convert nitrous oxides to N<sub>2</sub>, passed over absorber columns to remove moisture and carbon dioxide, and nitrogen content determined by thermal conductivity. Similar instruments have the capability of the simultaneous analysis of carbon or sulfur. The method is used to assess plant nitrogen sufficiency levels. The method has a detection limit of 0.10% nitrogen (dry sample basis) and is generally reproducible to within ± 5.0%.

### Equipment

1. Analytical balance: 250 g capacity, resolution ± 0.1 mg.
2. Total Nitrogen Analyzer: Leco, Carl-Erba, Elementar and Perkin-Elmer with resistance furnace with thermal conductivity detector and operating supplies.
3. Tin (Sn) foil encapsulating cups.
4. Desiccator, containing a desiccating agent.

### Reagent

1. Compressed Oxygen gas, 99.995 % purity.
2. Helium carrier gas, 99.995 % purity.
3. Nitrogen calibration standards: EDTA, 9.57% ± 0.05% N; Glycine p-toluenesulfonate (C<sub>9</sub>H<sub>13</sub>O<sub>6</sub>SN), 5.66 ± 0.05% N; highly purified acetanilide (C<sub>8</sub>H<sub>9</sub>NO) 10.37 ± 0.05% N; and Leco calibration standard (PN 502-055), 2.40 ± 0.03% N.

### Procedure

1. Determine the moisture content of the botanical material, Method P - 1.10.
2. Weigh 150.0 ± 5.0 mg of air dried botanical material pulverized to pass 40 mesh sieve (See Comments #1 and #2) and place in into a tared tin foil container (instrument specific), encapsulate, and record sample weight to the nearest 0.1 mg (See Comment #3).
3. Initialize the instrument following manufacturers suggested protocol. Conduct a system leak check on combustion system. Perform blank stabilization test, analyze consecutive blanks until the blanks stabilize at a constant value.
4. Adjust and operate the instrument according to manufacturer instructions using calibration standards (provided by manufacturer or obtained commercially). Enter sample dry matter content and analyze unknown sample for total nitrogen. Report results to the nearest 0.001% nitrogen. (See Comment #4 and #5).

### Calculation

Report sample nitrogen concentration to the nearest 0.001%

## Comments

1. Samples limited in material, should be dried over phosphorus pentoxide or magnesium perchlorate for forty-eight (48) hours and analyzed with no correction for moisture content or reported on as received basis.
2. Sample particulate must be ground to pass a 40 mesh screen (< 0.40 mm) in order to assure adequate sample homogeneity for instruments utilizing sample sizes in excess of 150 mg. For instruments utilizing sample sizes 5 - 10 mg (Carl-Erba, and Perkin-Elmer) it is recommended that samples be finely ground to pass a 60 mesh sieve (<250 um) prior to analysis to ensure homogeneity. For the Elementar instrument sample size maybe increased to 1000 mg.
3. Sample weight may be entered into instrument software using a balance interface.
4. Nitrogen content as determined by automated combustion method are generally slightly greater than values determined by the Kjeldahl (TKN) method. The TKN method may have incomplete recovery from oxidized forms of nitrogen and that in heterocyclic rings.
5. Instruments may have the capability for the simultaneous analysis of sulfur, carbon and hydrogen.

## Literature

McGeehan, S.L. and D.V. Naylor. 1988. Automated instrumental analysis of carbon and nitrogen in plant and soil samples. *Comm. in Soil Sci. Plant Anal.* 19:493-50-5.

Sheldrick, B.H. 1986. Test of the LECO CHN-600 Determinator for soil carbon and nitrogen analysis. *Can. J. Soil Sci.* 66:543-545.

Shepers, J.S. D.D. Francis, and M.T. Thompson. 1989. Automated total nitrogen of soil and plant samples. *Comm. in Soil Sci. Plant Anal.* 20:949-959.

Sweeney, Rose A. 1989. Generic combustion method for determination of crude protein in feeds: Collaborative study. *J. Assoc. Off. Anal. Chem.* 72:770-774.

Yeomans, J.C. and J.M. Bremmer. 1991. Carbon and nitrogen analysis of soils by automated combustion techniques. *Comm. in Soil Sci. Plant Anal.* 22:843-850.

### Scope and Application

This method quantitatively determines the amount of sulfur in all forms (sulfate, sulfite, and protein) in botanical materials using a resistance furnace and a thermal conductivity detector. It is based on the method originally described by Dumas and later modified by Sweeny (1989) whereby botanical samples, encased in tin (Sn) foil, are ignited in a resistance furnace at approximately 900 °C, in helium and oxygen environment in a quartz combustion tube. An aliquot of the combustion gas is passed through a copper catalyst to remove oxygen and other combustion products, passed over absorber columns to remove moisture and carbon dioxide, and nitrogen content determined by thermal conductivity. Similar instruments have the capability of the simultaneous analysis of carbon or nitrogen. The method is used to assess plant sulfur sufficiency levels. The method has a detection limit of 0.10% sulfur (dry sample basis) and is generally reproducible to within  $\pm 7.0\%$ .

### Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.1$  mg.
2. Total Sulfur Analyzer: Leco, CNS-2000 or Carl-Erba with combustion furnace with detector and operating supplies.
3. Desiccator, containing a desiccating agent.

### Reagent

1. Compressed Oxygen gas, 99.995 % purity.
2. Helium carrier gas, 99.995 % purity.
3. Sulfur calibrations standard: Glycine p-toluenesulfonate ( $C_9H_{13}O_6SN$ ),  $12.2 \pm 0.05\%$  S.

### Procedure

1. Determine the moisture content of the botanical material, Method P - 1.10.
2. Weigh  $1000.0 \pm 5.0$  mg of air dried botanical material pulverized to pass 40 mesh sieve (See Comments #1 and #2) and place in into a tared tin foil container, encapsulate, and record sample weight to the nearest 1 mg (See Comment #3).
3. Initialize the instrument following manufacturers suggested protocol. Conduct a system leak check on combustion system. Perform blank stabilization test, analyze consecutive blanks until the blanks stabilize at a constant value.
4. Adjust and operate the instrument according to manufacturer instructions using calibration standards (provided by manufacturer or obtained commercially). Enter sample dry matter content and analyze unknown sample for total sulfur. Report results to the nearest 0.001% sulfur. (See Comment #4).

### Calculation

Report sample sulfur concentration to the nearest 0.01%

## Comments

1. Samples limited in material, should be dried over phosphorus pentoxide or magnesium perchlorate for forty-eight (48) hours and analyzed with no correction for moisture content or reported on as received basis.
2. Sample particulate must be ground to pass a 40 mesh screen (< 0.40 mm) in order to assure adequate sample homogeneity for instruments utilizing sample sizes in excess of 150 mg. For instruments utilizing sample sizes 5 - 10 mg (Carl-Erba) it is recommended that samples be finely ground to pass a 60 mesh sieve (<250 um) prior to analysis to ensure homogeneity.
3. Sample weight may be entered into instrument software using a balance interface.
4. Sulfur content as determined by automated combustion method are generally of greater precision than other methods using acid digestion of analysis.

## Literature

Beaton, J.D., G.K. Burns and J. Platou. 1968. Determination of sulfur in soils and plant material. Technical Bulletin No. 14, The Sulfur Institute, Washington D.C. 1968.

Sheldrick, B.H. 1986. Test of the LECO CHN-600 Determinator for soil carbon and nitrogen analysis. Can. J. Soil Sci. 66:543-545.

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Sweeney, Rose A. 1989. Generic combustion method for determination of crude protein in feeds: Collaborative study. J. Assoc. Off. Anal. Chem. 72:770-774.

Yeomans, J.C. and J.M. Bremmer. 1991. Carbon and nitrogen analysis of soils by automated combustion techniques. Comm. in Soil Sci. Plant Anal. 22:843-850.

**EXTRACTABLE POTASSIUM, NITRATE, AMMONIUM  
ORTHO-PHOSPHATE, AND CHLORIDE OF BOTANICAL MATERIALS**  
2% Acetic Acid Extraction

**B - 3.10**

### Scope and Application

The method semiquantifies the concentration of potassium (K), ammonium (NH<sub>4</sub>-N), nitrate (NO<sub>3</sub>-N), ortho-phosphate (PO<sub>4</sub>-P) and chloride (Cl) in botanical materials by extraction with a 2% acetic acid solution. Dilute acetic acid does not quantitatively extract these ions from botanical tissue. Potassium is determined by AES or AAS; ammonium spectrophotometrically; nitrate is determined spectrophotometrically at 520 nm by the Griess-Illasvay method (cadmium reduction); ortho-phosphate in the extract is determined spectrophotometrically at 880 nm by reacting with paramolybdate; and chloride is determined by coulometric titration or ion selective electrode. The method has been used primarily to determine K, NH<sub>4</sub>-N, NO<sub>3</sub>-N, PO<sub>4</sub>-P, and Cl for assessing plant fertility and chloride status (Johnson and Ulrich, 1959; Chapman and Pratt, 1961). The method can also be used to determine sulfate sulfur. Generally the method detection limit is approximately 10 mg kg<sup>-1</sup> (sample dry basis) and is generally reproducible to within ± 10.0%.

### Equipment

1. Analytical balance: 250 g capacity, resolution ± 0.1 mg.
2. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute.
3. 250 mL extraction vessel with cap and filtration container.
4. Repipette dispenser calibrated to 50.0 ± 0.2 mL.
5. Whatman No. 2V 11 cm filter paper or equivalent highly retentive paper.
6. Spectrophotometer instrument, 520 and 660 nm.
7. Atomic Emission Spectrometer (AES) or Atomic Absorption Spectrometer (AAS) instrumentation.
8. Coulometric titrator or chloride ion selective electrode.

### Reagents

1. Acetic acid extraction solution: Dilute 20 mL acetic acid, CH<sub>3</sub>COOH (99%), in 50 mL deionized water and dilute to 1.0 L. Care must be taken to use high purity acetic acid to avoid nitrate and chloride contamination.
2. Standard calibration solutions of K, NH<sub>4</sub>-N, NO<sub>3</sub>-N, PO<sub>4</sub>-P and Cl. Prepare multiple calibration standards according to specific method and manufacturer's specifications prepared from 1000 mg L<sup>-1</sup> standard solution and diluted to final volume with 2% acetic acid.

### Procedure

1. Determine the moisture content of the botanical material, Method P - 1.10.
2. Weigh out 200.0 ± 1.0 mg of air dried botanical material (See Comments #1 and #2) and place in 250 mL extraction vessel. Include a method blank.
3. Add 50.0 ± 0.2 mL of 2% acetic acid extraction solution and place on reciprocating mechanical shaker for thirty (30) minutes (See Comment #3). Include a method blank.
4. Filter, refilter if filtrate is cloudy (See Comments #4, #5, #6 and #7) and retain for analysis.

## 5. Analysis:

- i. For the determination of K, analyze for K using Atomic Emission Spectrometer (AES) or Atomic Absorption Spectrometer (AAS).
- ii. For the determination of NH<sub>4</sub>-N, analyze for NH<sub>4</sub>-N using by spectrophotometric, diffusion-conductivity instruments or distillation techniques.
- iii. For determination of NO<sub>3</sub>-N, analyze using spectrophotometric Method S - 3.10 or method 418-C or 418-F for NO<sub>3</sub>-N listed in "Standard Method for the Analysis of Waste Water", 1985. Record concentration in mg L<sup>-1</sup> NO<sub>3</sub>-N in extract.
- iv. For determination of PO<sub>4</sub>-P, analyze using spectrophotometric Method S - 4.10 or method 424-F in "Standard Method for the Analysis of Waste Water", 1985. Record concentration in mg L<sup>-1</sup> PO<sub>4</sub>-P in extract.
- v. For determination of Cl, analyze for Cl according to Method S - 1.40 or method 407-B listed in "Standard Method for the Analysis of Waste Water", 1985. Record concentration in mg L<sup>-1</sup> Cl in extract.

## Calculations

Report K and Cl in sample as to the nearest 0.01 %:

$$\% = \frac{(\text{mg L}^{-1} \text{ K or NH}_4\text{-N in extract} - \text{method blank}) \times (0.025)}{(\text{Sample size, mg}) \times (\text{Dry Matter Content } (\%) / 100)}$$

Report NH<sub>4</sub>-N, NO<sub>3</sub>-N, and PO<sub>4</sub>-P in sample as to the nearest 10 mg kg<sup>-1</sup>:

$$\text{mg kg}^{-1} = \frac{(\text{Extract analyte Conc. mg L}^{-1} - \text{method blank}) \times (250)}{\text{Dry matter content } (\%) / 100}$$

## Comments

1. Botanical materials must be ground to pass 40 mesh screen (< 0.425 mm) in order to insure adequate sample homogeneity.
2. Sample mass may be adjusted in accordance with expected analyte concentrations. For materials containing < 500 mg kg<sup>-1</sup> NH<sub>4</sub>-N or NO<sub>3</sub>-N increase sample size to 500 mg.
3. Check repipette dispensing volume, calibrate using an analytical balance.
4. Check filter paper supply for possible contamination of analytes. If significant contamination is found (> 10 mg kg<sup>-1</sup> on a sample basis), rinse filter paper with acetic acid extraction solution or filter extract with serum separator tubes.
5. Acetic acid extracts may be stored for up to 14 days, if stored at 4°C and/or with 100 uL of toluene or thymol.
6. Extracts may be retained for analysis of total potassium, ammonium nitrogen and sulfate-sulfur.
7. Samples having K, NH<sub>4</sub>-N, NO<sub>3</sub>-N, PO<sub>4</sub>-P and Cl concentrations exceeding the highest standard will require dilution and reanalysis.

## Literature

Carlson, R.M., R.I. Cabrera, J.L. Paul, J. Quick, and R.Y. Evans. 1990. Rapid direct determination of ammonium and nitrate in soil and plant tissue extracts. *Comm. in Soil Sci. Plant Anal.* 21:1519-1529.

Chapman, H.D. and P.F. Pratt. 1961. *Methods of analysis for soils, plants and waters*, Priced Publication 4034. Berkeley: University of California, Division of Agriculture Sciences.

Johnson, C.M. and A. Ulrich. 1959. Analytical methods for use in plant analysis. *Bulletin 766*. Berkeley: University of California, Agricultural Experiment Station. p. 26-78.

Standard Method of the Examination of Waste Water. 1985. p. 265-297. A.H. Franson. (ed.) 16th ed. American Public Health Association, American Water Works Association, and Water Pollution Control Federation.



### Scope and Application

The method semiquantifies the concentration of nitrate ( $\text{NO}_3\text{-N}$ ), in botanical materials by extraction with an aluminum sulfate solution and subsequent determination by ion-selective electrode (ISE). The ISE determines  $\text{NO}_3\text{-N}$  by measuring an electrical potential developed across a thin layer of water-immiscible liquid or gel ion exchanger that is selective for  $\text{NO}_3$  ions. This layer of ion exchanger is held in place by a porous membrane. The ISE is susceptible to interferences of  $\text{Cl}^-$ ,  $\text{HCO}_3^-$ ,  $\text{SO}_4^{2-}$  and is sensitive to changes in solution ionic strength (i.e. high salt). Problems with precision have been noted by Mack and Sanderson (1971) and Miller, Amacher and Dellavalle (1996). The method has been used primarily to determine  $\text{NO}_3\text{-N}$  for assessing plant nitrogen fertility (Johnson and Ulrich, 1959; Chapman and Pratt, 1961). Generally the method detection limit is approximately  $200 \text{ mg kg}^{-1}$  (sample dry basis) and is generally reproducible to within  $\pm 18.0\%$ .

### Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.001 \text{ g}$ .
2. Repipette dispenser, calibrated to  $25.0 \pm 0.2 \text{ mL}$ .
3. Reciprocating horizontal mechanical shaker, capable of 180 oscillations per minute.
4. Whatman No. 2V 11 cm filter paper or equivalent highly retentive paper.
5. Nitrate ion sensitive electrode.
6. pH/ion meter or pH-millivolt meter.

### Reagents

1. Deionized water, ASTM type I Grade.
2. Extracting Solution: ionic strength adjusting solution -  $0.01\text{M Al}_2(\text{SO}_4)_3$ ,  $0.02\text{M H}_3\text{BO}_3$ ,  $0.01\text{M Ag}_2\text{SO}_4$ , and  $0.02 \text{ M NH}_2\text{HSO}_3$  (sulfamic acid): Dissolve 67 g of  $\text{Al}_2(\text{SO}_4)_3 \cdot 18\text{H}_2\text{O}$ , 12 g of  $\text{H}_3\text{BO}_3$ , 20 g of  $\text{Ag}_2\text{SO}_4$  and 19 g of  $\text{NH}_2\text{HSO}_3$  in water and dilute to 10 liters.
3. Standard nitrate solutions. To a 1000 mL volumetric flask, add 0.7221 g of oven dry  $\text{KNO}_3$ ; make up to volume with extracting solution. This gives a solution containing  $100 \text{ mg L}^{-1}$  of  $\text{NO}_3\text{-N}$ . Prepare nitrate calibration standards from extraction solution of 5.0, 10.0, 15.0, 20.0, 30.0, and  $50.0 \text{ mg L}^{-1}$ .

### Procedure

1. Determine the moisture content of the botanical material, Method P - 1.10.
2. Weigh out  $500.0 \pm 1.0 \text{ mg}$  of air dried botanical material (See Comments #1, #2, #3) and place in 50 mL extraction vessel.
3. Add  $25.0 \pm 0.2 \text{ mL}$  of extracting solution and place on reciprocating mechanical shaker for thirty (30) minutes. Include a method blank.
4. Filter extract, refilter if filtrate is cloudy and retain for analysis.
5. Calibrate ion selective electrode/millivolt meter using standard calibration solutions and operate instrument in accordance with manufacturer instructions. Develop calibration curve for the ion selective electrode using standards. Determine nitrate concentration of plant sample and record results as  $\text{mg L}^{-1}$  of nitrate in extract solution (See Comment #4 and #5).

### Calculations

Report mg of  $\text{NO}_3\text{-N}$  in sample as to the nearest  $10 \text{ mg kg}^{-1}$ :

$$\text{mg kg}^{-1} = \frac{(\text{NO}_3\text{-N in extract mg L}^{-1} - \text{method blank}) \times (50)}{\text{Dry matter content (\%)} / 100}$$

### Comments

1. Botanical materials must be ground to pass 40 mesh screen (< 0.425 mm) in order to insure adequate homogeneity.
2. Sample mass may be adjusted in accordance with expected analyte concentrations. For materials containing < 500 mg kg<sup>-1</sup> NO<sub>3</sub>-N increase sample size to 1000 mg.
3. Check repipette dispensing volume, calibrate using an analytical balance.
4. Routinely check ISE calibration every third sample using a mid range standard. In specific instances the ISE maybe susceptible to radio frequency energy from surrounding electronic equipment (Carlson, 1992).
5. Samples having nitrate concentrations exceeding the highest standard will require dilution and reanalysis.

### Literature

Baker, A.S. and R. Smith. 1969. Extracting solution for potentiometric determination of nitrate in plant tissue. *J. of Agric. and Food Chem.* 17:1284-1287.

Carlson, R.M. 1992. Personal Communication.

Chapman, H.D. and P.F. Pratt. 1961. *Methods of analysis for soils, plants and waters*, Priced Publication 4034. Berkeley: University of California, Division of Agriculture Sciences.

Dahnke, W.C. 1971. Use of the nitrate specific ion electrode in soil testing. *Comm. in Soil Sci. and Plant Anal.* 2(2):73-84.

Mack, A.R. and R.B. Sanderson. 1971. Sensitivity of the nitrate-ion membrane electrode in various oil extracts. *Can. J. Soil Sci.* 51:95104.

Milham, P.J., A.S. Awad, R.E. Paul and J.H. Bull. 1970. Analysis of plants, soils, and waters for nitrate by using an ion-selective electrode. *Analyst* 95:751-759.

Miller, R.O., J. Amacher and N. Dellavalle. 1996. A Proficiency Testing Program for the agricultural laboratory inudstry, results of the 1994 program. *Comm. in Soil Sci. and Plant Anal.* 27(3&4):451-461.

### Scope and Application

This method quantitatively determines the amount of sulfate-sulfur ( $\text{SO}_4\text{-S}$ ) in botanical materials by extraction with a solution of 2% acetic acid with subsequent determination by turbidimetric analysis. The method may not quantitatively extract  $\text{SO}_4\text{-S}$  on some botanical materials which have a high anion exchange capacity. Turbidimetric analysis is based on the formation of  $\text{BaSO}_4$  particulates in a suspension and subsequent measurement of optical density. The turbidimetric will require practice to become proficient with the method. Sulfate-sulfur analysis can be used for the diagnosis of sulfur deficiency in specific crops (ie. alfalfa, sugar beet, clover). The method has a detection limit of  $20 \text{ mg kg}^{-1}$  (dry basis) and is generally reproducible to within  $\pm 12\%$ .

### Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.1 \text{ mg}$ .
2. Reciprocating horizontal mechanical shaker, Eberbach, capable of 180 oscillations per minute.
3. Boro-silicate test tube 16 x 150 mm with cap.
4. Repipette dispenser(s) calibrated to  $10.0 \pm 0.1 \text{ mL}$  and  $3.0 \pm 0.05 \text{ mL}$ .
5. Serum separator tubes, (PN-02-657-3, 16 x 150 mm, Fisher Sci. Co.).
6. Pipette 10.0 mL.
7. Nephelometer (preferred), turbidimeter or spectrophotometer 420 nm.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Acetic acid extraction solution, 2% (v/v): Dilute 20 mL acetic acid,  $\text{CH}_3\text{COOH}$  (99%), in 50 mL deionized water and dilute to 1000 mL. Verify purity of acetic acid, sulfate-sulfur content should be less than  $1.0 \text{ mg L}^{-1}$ .
3. Barium chloride / turbidimetric solution. Dissolve 60 g  $\text{BaCl}_2$  in 600 mL deionized water and add 150 mL of Tween 80 suspension agent. Shake vigorously let stand 24 hours. Dissolution of Tween 80 maybe facilitated by heating. (See Comment #1 and #2).
4. Acidification solution, 6 N  $\text{HCl}$ . Mix 150 mL of concentrated  $\text{HCl}$  with 150 mL deionized water. Add 1.5 mL of 2000  $\text{mg L}^{-1}$   $\text{SO}_4\text{-S}$  stock solution.
5. Standard sulfate-sulfur solutions, 2000  $\text{mg L}^{-1}$ . Dissolve 2.7176 g of oven dried potassium sulfate ( $\text{K}_2\text{SO}_4$ ) with deionized water and dilute to 250 mL final volume. Prepare sulfate-sulfur calibration solutions of concentration: 1.0, 2.0, 5.0, 10, 20, and 40  $\text{mg L}^{-1}$  and dilute to volume with 2 % acetic acid extraction solution.

### Procedure

1. Determine the moisture content of the botanical material, Method 1.10.
2. Weigh out  $100.0 \pm 1.0 \text{ mg}$  of air dried botanical material (See Comment #3 and #4) and place in 16 x 150 mm test tube.
3. Add  $10.0 \pm 0.1 \text{ mL}$  of 2 % acetic acid extraction solution and place on reciprocating mechanical shaker for thirty (30) minutes (See Comment #5). Include an extract blank and quality control samples.
4. Filter, using serum separator, refilter if filtrate is cloudy and retain for analysis.
5. Pipette 5.0 mL of standard sample into 10 mL test tube. Repeat using standard sulfate-sulfur solutions and quality control samples.

6. Add 1.5 mL of acidification solution.
7. Add 3 mL of barium chloride/ turbidimetric solution, swirl gently to mix. Let stand for five (5) minutes and repeat samples.
8. Calibrate nephelometer using standard calibration solutions and operate instrument in accordance with manufacturer instructions. Develop calibration curve using standards. Determine sulfate-sulfur concentration in extract solution and record to the nearest 0.5 mg L<sup>-1</sup> (See Comment #6).

### Calculations

1. Report plant sulfate-sulfur (SO<sub>4</sub>-S):

$$\text{mg kg}^{-1} = \frac{(\text{mg L}^{-1} \text{SO}_4\text{-S in extract} - \text{method blank}) \times (\text{DF})}{(\text{Dry Matter Content (\%)} / 100)}$$

$$\text{Dilution Factor (DF)} = 100$$

### Comments

1. Use BaCl<sub>2</sub> specifically designated for turbidimetric determination of sulfate-sulfur. Sources: J.T. Baker Cat. Parr Turbidimetric BaCl<sub>2</sub>, JT0974-5; VWR JT0974-5; and GFS Chemicals Reagent Grade ACS #602.
2. A number of suspension agents have been reported in the literature which include: gum acacia, gelatin, glycerol, PVP-K30 (polyvinylpyrrolidinone), and Tween 80 which have proven effective in turbidimetric analysis. Each of these will require experimentation and practice using SO<sub>4</sub>-S spiking to fully refine the technique. For use of PVP-K30 (polyvinylpyrrolidinone) add 10 g to 700 mL and dilute to 1 L final volume.
3. Botanical materials must be ground to pass a 40 mesh screen (< 400 μm) in order to insure adequate homogeneity.
4. Sample size maybe adjusted in accordance with expected SO<sub>4</sub>-S concentrations. For materials containing < 250 mg kg<sup>-1</sup> SO<sub>4</sub>-S use 200 mg of sample material and 10.0 mL of acetic acid extraction solution using serum separator tubes.
5. Check Repipette volume, calibrate using an analytical balance.
6. Samples having SO<sub>4</sub>-S concentrations exceeding the highest standard will require dilution and reanalysis.

### Literature

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**PHOSPHORUS, POTASSIUM, CALCIUM, MAGNESIUM, SODIUM,  
BORON ZINC, MANGANESE, IRON, COPPER AND MOLYBDENUM  
OF BOTANICAL MATERIALS**

**B - 4.10**

Dry Ash

### Scope and Application

The method quantitatively determines the concentration of P, K, Ca, Mg, Na, B, Zn, Mn, Fe, Cu and Mo in botanical materials utilizing a high temperature dry oxidation of the organic matter and dissolution of the ash with hydrochloric acid. Digest analyte concentrations are determined by atomic absorption spectrometry (AAS) and/or inductively coupled plasma atomic emission spectrometry (ICP-AES). Analysis of P and B may be conducted using spectrophotometric methods. The procedure is not quantitative for sulfur and other elements which are easily volatilized (i.e. Se, As, Hg). Ashing temperatures exceeding 500 °C will result in poor recoveries of Al, B, Cu Fe, K and Mn (Issac and Jones, 1972). Results for boron may be inconsistent due to volatilization and desorption in the muffle furnace. The method detection limit is approximately 0.04 % for P, K, Ca, and Mg and 4.0 mg kg<sup>-1</sup> for B, Zn, Mn, Fe, and Cu. The method is generally reproducible within ± 10%.

### Equipment

1. Analytical balance: 250 g capacity, resolution ± 1 mg.
2. Porcelain crucibles, 30 cc capacity.
3. Muffle furnace capable of 500 °C.
4. Repipette, 10.0 ± 0.2 mL.
5. Volumetric labware, 50 mL, plastic.
6. Atomic Absorption Spectrophotometer (AAS) and/or Inductively Coupled Plasma Atomic Emission Spectrometer (ICP-AES).

### Reagents

1. Deionized water, ASTM Type I grade.
2. 1.0 N HCl solution, prepared by mixing 83.5 mL concentrated HCl and diluting to 1.0 L.
3. Standard calibration solutions of P, K, Ca, Mg, Na, B, Zn, Mn, Cu, and Fe. From 1000 mg L<sup>-1</sup> reference solutions: Prepare five multi-element standards: of K, Ca, Mg ranging from 5.0 - 500 mg L<sup>-1</sup>; P and Na ranging from 1.0 - 100 mg L<sup>-1</sup>; and B, Zn, Mn, Fe, and Cu, ranging from 0.10 - 10.0 mg L<sup>-1</sup>. Dilute standard calibration solutions with 0.1 N HCl.

### Procedure

1. Determine the moisture content of the botanical material, Method P - 1.10.
2. Weigh 1000 ± 5.0 mg of plant material into a porcelain crucible. Include a method blank. (See Comments #1 and #2).
3. Place crucible in a muffle furnace and ramp temperature to 500 °C over two (2) hours. Ash samples for four (4) hours at 500 °C (See Comment #3).
4. Allow to cool to room temperature in muffle furnace, slowly open door and remove ashed samples. Take caution not to disturb sample ash while transferring from furnace.
5. Dissolve ash with 10.0 mL of 1.0 N HCl solution (See Comments #4 and #5). Dissolution of ash may be facilitated by heating.
6. Quantitatively transfer the contents of the crucible into a 50.0 mL volumetric flask, dilute to volume with deionized water, cap and invert three times.

7. Elemental analysis of plant digests can be made using atomic emission spectrometry (AES) atomic absorption spectrometry (AAS), inductively coupled plasma atomic emission spectrometry (ICP-AES), and/or other methodologies (See Comment #6 and #7). The method chosen will determine specific matrix modifications, calibration standard range and the need for instrument specific sample preparations and dilutions. Determination of trace elements by ICP-AES (Co, Cd, Ni, Mo, Pb) maybe facilitated by the use of an ultrasonic nebulizer (Soltanpour, 1996). Adjust and operate instruments in accordance with manufacturer's instructions. Calibrate instrument using standard calibration solutions. Determine the analyte concentrations of a method blank, unknown samples and record analyte concentrations in mg L<sup>-1</sup>.

### Calculations

For P, K, Ca, Mg and Na report results to the nearest 0.001%:

$$\% \text{ analyte} = \frac{(\text{mg L}^{-1} - \text{method blank}) \times (50) \times (0.0001)}{\text{Dry matter (\%)/100}}$$

For Mn and Fe report results to the nearest 1 mg kg<sup>-1</sup>; B, Zn and Cu the nearest 0.1 mg kg<sup>-1</sup>:

$$\text{mg kg}^{-1} \text{ analyte} = \frac{(\text{mg L}^{-1} - \text{method blank}) \times (50)}{\text{Dry matter (\%)} / 100}$$

### Comments

1. Labware should be cleaned (1) soak crucibles in 1% solution of labware detergent for one hour; (2) rinse vessels in tap water; (3) rinse in solution of 0.5 N HCl; (4) three deionized water rinses (ASTM Type I grade); and (5) dry for one hour at 80 °C.
2. Sample material must be ground to pass a 40 mesh screen (< 0.40 mm opening), to ensure homogeneity.
3. Ashing temperatures are not to exceed 500 °C to avoid potential volatilization of Al, B, Cu, K, and Mn.
4. Check pipette dispensing volume, calibrate using an analytical balance.
5. When adding reagent to vessels and handling digests always wear protective clothing (i.e. eye protection, lab coat, disposable gloves and shoes). Always handle reagents and opening of vessels in an acid hood capable of high air flow, 100 cfm.
6. Centrifuging may be necessary to clear the digest.
7. Samples having analyte concentrations exceeding the highest standard will require dilution and reanalysis.

### Literature

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**PHOSPHORUS, POTASSIUM, SULFUR, CALCIUM, MAGNESIUM  
SODIUM, ZINC, MANGANESE, COPPER, IRON AND MOLYBDENUM  
OF BOTANICAL MATERIALS**

**B - 4.20**

Nitric/Perchloric Acid Digest, Wet Ashing Open vessel

### Scope and Application

The method quantitatively determines the concentration of P, S, K, Ca, Mg, Na, Al, Zn, Mn, Fe, Cu and Mo in botanical materials utilizing a nitric-perchloric acid digestion of organic matter in conjunction with external heating. Digest analyte concentrations are determined by atomic absorption spectrometry (AAS) and/or inductively coupled plasma atomic emission spectrometry (ICP-AES). Analysis of P and S may be conducted using spectrophotometric and turbidimetric methods, respectively. The method requires predigestion with HNO<sub>3</sub>, followed by addition of HClO<sub>4</sub> and digestion at high temperatures. Extreme caution is to be followed when using perchloric acid (Schilt, 1979) which may react violently with untreated organic materials and result in an explosion. A special hood is required to handle perchloric acid fumes. Reflux funnels are placed over the digestion tubes to reduce volatilization and minimize oxygen. Alternatives to the use of perchloric acids using hydrogen peroxide have been reported (Haug and Schulte, 1985; Havlin and Soltanpour, 1980). The method can also be used for the determination of trace-elements (Pb, Ni, Cd, etc). Generally the method detection limit is approximately 0.02% for P, S, K, Ca, Mg and Na; and 0.5 mg kg<sup>-1</sup> (sample dry basis) for Al, Zn, Mn, Fe and Cu. Generally reproducible is within ± 7.0%.

### Equipment

1. Analytical balance: 250 g capacity, resolution ± 0.1 mg.
2. Block Digester (400 °C) and perchloric hood.
3. 50 mL volumetric digestion tubes and 25 mm reflux funnels.
4. Repipette dispensers, calibrated to 6.0 ± 0.05 mL and 2.0 ± 0.01 mL
5. Volumetric labware, 25 mL.
6. Atomic Absorption Spectrophotometer (AAS) and/or Inductively Coupled Plasma Atomic Emission Spectrometer (ICP-AES), vacuum system.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Concentrated nitric acid, reagent grade.
3. 70% perchloric acid, reagent grade.
4. Standard calibration solutions of P, K, S, Ca, Mg, Na, Al, Zn, Mn, Cu, Fe and Mo. Prepare five multi-element standards: of K, Ca, Mg ranging from 5 - 500 mg L<sup>-1</sup>; P, S and Na ranging from 1.0 - 100 mg L<sup>-1</sup>; and Al, Zn, Mn, Fe; and Cu and Mo , ranging from 0.02 - 10.0 mg L<sup>-1</sup>. Dilute standard calibration solutions with 5% HNO<sub>3</sub> and 1% HClO<sub>4</sub> by volume.

### Procedure

1. Determine the moisture content of the botanical material, Method P - 1.10.
2. Weigh 500.0 ± 0.5 mg of sample into a 50 mL volumetric digestion tube (See Comment #1 and #2). Include a method blank.
3. Using a repipette add 6.0 mL nitric acid, a boiling chip (teflon or glass) and swirl to thoroughly wet the sample (See Comment #3, #4 and #5).



4. Place 25 mm reflux funnels over the samples and allow to predigest at room temperature for sixty (60) minutes (16 h preferred).
5. Place the digestion tubes on a digestion block for sixty (60) minutes at 150 °C.
6. Remove, cool to room temperature and using a repipette slowly add 2.0 mL of HClO<sub>4</sub> through the funnels.
7. Place the tubes in the block at 215 °C for two (2) hours, after the HNO<sub>3</sub> fumes have evolved.
8. Remove the funnels ten (10) minutes before the end of the digestion.
9. Remove the tubes from the digestion block, cool twenty (20) minutes in a hood, and add 10 mL of deionized water on a hot plate (90 °C).
10. Mix, using a vortex stirrer, cool and dilute to final volume. Filtering or centrifuging may be necessary to remove all particulate matter in the digest prior to analysis. Quantitatively transfer contents of digestion tube into a 25 mL volumetric flask.
11. Elemental analysis of plant digests can be made using atomic emission spectrometry (AES) atomic absorption spectrometry (AAS), inductively coupled plasma atomic emission spectrometry (ICP-AES), and/or other methodologies. The method chosen will determine specific matrix modifications and the need for instrument specific sample preparations and dilutions. Determination of trace elements by ICP-AES (Cu, Mo) maybe facilitated by the use of an ultrasonic nebulizer (Soltanpour, 1996). Adjust and operate instruments in accordance with manufacturer's instructions. Calibrate instrument using calibration solutions and record concentration of analytes as mg L<sup>-1</sup> (See Comments #6, #7 and #8). Determine the analyte concentrations of a method blank, unknown samples and record concentrations in mg L<sup>-1</sup>.

### Calculations

Report P, K, S, Ca, Mg and Na results to the nearest 0.001%:

$$\% \text{ analyte} = \frac{(\text{mg L}^{-1} - \text{method blank}) \times (50) \times (0.0001)}{\text{Dry matter (\%)} / 100}$$

Report Mn Al and Fe results to the nearest 1 mg kg<sup>-1</sup>; Zn and Cu the nearest 0.1 mg kg<sup>-1</sup>:

$$\text{mg kg}^{-1} \text{ analyte} = \frac{(\text{mg L}^{-1} - \text{method blank}) \times (50)}{\text{Dry matter (\%)} / 100}$$

### Comments

1. Labware cleaning: (1) soak digestion tubes in 1% solution of laboratory detergent for one hour; (2) rinse vessels in tap water; (3) rinse in solution of 0.5 N HCl; (4) three deionized water rinses (ASTM Type I grade); and (5) dry for one hour at 80 °C.
2. Sample material must be ground to pass a 40 mesh screen (< 0.40 mm opening), to ensure homogeneity.
3. When adding reagent to vessels and handling digests always wear protective clothing (i.e. eye protection, lab coat, disposable gloves and shoes). Always handle reagents and opening of vessels in an acid hood capable of high air flow, 100 cfm.
4. It is essential that the entire sample be pretreated with nitric acid to ensure at least partial oxidation of the organic matter before the addition of perchloric acid. Caution: the use of perchloric acid in the presence of untreated organic matter can lead to rapid oxidation of the sample and a possible explosion (Blanchar, 1986).
5. Check repipette dispensing volume, calibrate using an analytical balance.

6. The method may not be quantitative for potassium since this alkali metal may form a precipitate with perchlorate.
7. Samples having analyte concentrations exceeding the highest standard will require dilution and reanalysis.
8. Increase sample mass to 2.00 g for the determination of trace metals, such as Ba, Cd, Cr, Mo, and Sr to 1000 mg and adjust dilution factor to 12.5.

## Literature

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**PHOSPHORUS, POTASSIUM, SULFUR, CALCIUM, MAGNESIUM  
SODIUM, ZINC, MANGANESE, COPPER, IRON AND MOLYBDENUM  
OF BOTANICAL MATERIALS**

**B - 4.25**

Nitric/Hydrogen Peroxide, Wet Ashing Open vessel

### Scope and Application

The method quantitatively determines the concentration of P, S, K, Ca, Mg, Na, Al, Zn, Mn, Fe, Cu and Mo in botanical materials utilizing a nitric-hydrogen peroxide digestion of organic matter in conjunction with external heating and closely follows the methods described by Haung and Schulte, 1985 and Havlin and Soltanpour, 1980. Caution is to be followed when adding using H<sub>2</sub>O<sub>2</sub> will react violently with hot untreated organic materials. Digest analyte concentrations are determined by atomic absorption spectrometry (AAS) and/or inductively coupled plasma atomic emission spectrometry (ICP-AES). Analysis of P may be conducted using spectrophotometric methods. The method can also be used for the determination of trace-elements (Pb, Ni, Cd, etc). Generally the method detection limit is approximately 0.02% for P, S, K, Ca, Mg and Na; and 0.5 mg kg<sup>-1</sup> (sample dry basis) for Al, Zn, Mn, Fe and Cu. Generally reproducible is within ± 6.0%.

### Equipment

1. Analytical balance: 250 g capacity, resolution ± 0.1 mg.
2. Block Digester (150 °C) and fume hood.
3. 20 mL volumetric digestion tubes and 25 mm reflux funnels.
4. Repipette dispensers, calibrated to 6.0 ± 0.05 mL and 1.0 ± 0.01 mL
5. Volumetric labware, 25 mL.
6. Atomic Absorption Spectrophotometer (AAS) and/or Inductively Coupled Plasma Atomic Emission Spectrometer (ICP-AES), vacuum system.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Concentrated nitric acid, reagent grade.
3. 30% hydrogen peroxide, reagent grade.
4. Standard calibration solutions of P, K, S, Ca, Mg, Na, Al, Zn, Mn, Cu, Fe and Mo. Prepare five multi-element standards: of K, Ca, Mg ranging from 5 - 500 mg L<sup>-1</sup>; P, S and Na ranging from 1.0 - 100 mg L<sup>-1</sup>; and Al, Zn, Mn, Fe; and Cu and Mo , ranging from 0.02 - 10.0 mg L<sup>-1</sup>. Dilute standard calibration solutions with 5% HNO<sub>3</sub> and 1% HClO<sub>4</sub> by volume.

### Procedure

1. Determine the moisture content of the botanical material, Method P - 1.10.
2. Weigh 500.0 ± 0.5 mg of sample into a 20 mL volumetric digestion tube (See Comment #1 and #2). Include a method blank.
3. Using a repipette add 6.0 mL nitric acid and swirl to thoroughly wet the sample (See Comment #3, #4 and #5).
4. Allow to predigest at room temperature for ten (10) minutes.
5. Place the digestion tubes on a digestion block for ten (10) minutes at 80 °C.
6. Remove, cool for two (2) minutes and then add 2.0 mL of 30% H<sub>2</sub>O<sub>2</sub> solution via a pipette in two separate aliquots of 1.0 mL each (See Comment #6)
7. Place the tubes in the block at 130 °C for one (1) hour, or until total digest volume is reduced to approximately 2.0 - 3.0 mL .
8. Mix, using a vortex stirrer, cool and dilute to final volume. Filtering or centrifuging may be necessary to remove all particulate matter in the digest prior to analysis. Quantitatively transfer contents of digestion tube into a 25 mL volumetric flask.

9. Elemental analysis of plant digests can be made using atomic emission spectrometry (AES) atomic absorption spectrometry (AAS), inductively coupled plasma atomic emission spectrometry (ICP-AES), and/or other methodologies. The method chosen will determine specific matrix modifications and the need for instrument specific sample preparations and dilutions. Determination of trace elements by ICP-AES (Cu, Mo) maybe facilitated by the use of an ultrasonic nebulizer (Soltanpour, 1996). Adjust and operate instruments in accordance with manufacturer's instructions. Calibrate instrument using calibration solutions and record concentration of analytes as mg L<sup>-1</sup> (See Comments #6, #7 and #8). Determine the analyte concentrations of a method blank, unknown samples and record concentrations in mg L<sup>-1</sup>.

### Calculations

Report P, K, S, Ca, Mg and Na results to the nearest 0.001%:

$$\% \text{ analyte} = \frac{(\text{mg L}^{-1} - \text{method blank}) \times (50) \times (0.0001)}{\text{Dry matter (\%)} / 100}$$

Report Mn Al and Fe results to the nearest 1 mg kg<sup>-1</sup>; Zn and Cu the nearest 0.1 mg kg<sup>-1</sup>:

$$\text{mg kg}^{-1} \text{ analyte} = \frac{(\text{mg L}^{-1} - \text{method blank}) \times (50)}{\text{Dry matter (\%)} / 100}$$

### Comments

1. Labware cleaning: (1) soak digestion tubes in 1% solution of laboratory detergent for one hour; (2) rinse vessels in tap water; (3) rinse in solution of 0.5 N HCl; (4) three deionized water rinses (ASTM Type I grade); and (5) dry for one hour at 80 °C.
2. Sample material must be ground to pass a 40 mesh screen (< 0.40 mm opening), to ensure homogeneity.
3. When adding reagent to vessels and handling digests always wear protective clothing (i.e. eye protection, lab coat, disposable gloves and shoes). Always handle reagents and opening of vessels in an acid hood capable of minimum air flow 100 cfm.
4. It is essential that the entire sample be pretreated with nitric acid to ensure at least partial oxidation of the organic matter
5. Check repipette dispensing volume, calibrate using an analytical balance.
6. The addition of H<sub>2</sub>O<sub>2</sub> solution will cause a vigorous reaction which may froth and rise in the digestion tube. If the froth volume expands to more than 2X of the initial digest volume, extend the cooling time to three or four minutes to minimize frothing.
7. Samples having analyte concentrations exceeding the highest standard will require dilution and reanalysis.
8. Increase sample mass to 2.00 g for the determination of trace metals, such as Ba, Cd, Cr, Mo, and Sr to 1000 mg and adjust dilution factor to 12.5.

## Literature

Chapman, H.D. and P.F. Pratt. 1961. Methods of analysis for soils, plants, and waters. University of California Berkeley, Division of Agriculture Sciences. Priced Publication 4034.

Haug, C.L. and E.E. Schulte. 1985. Digestion of plant tissue for analysis by ICP emission spectroscopy. *Comm. Soil Sci. Plant Anal.* 16: 945-958.

Johnson, C.M. and A. Ulrich. 1959. Analytical methods for use in plant analysis. Bulletin 766. Berkeley: University of California, Agricultural Experiment Station. p. 26-78.

Schilt, Alfred A. 1979 Perchloric acid and perchlorates. Fedrich Smith Chemical Company. Columbus, OH.

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Tolg, G. 1974. The basics of trace analysis. P. 698-710. *In:* E. Korte (ed.) *Methodium Chemicum.* Vol 1 Analytical Methods. Part B. Micromethods, biological methods, quality control, automation. Academic Press, New York.

Zasoski, R.J. and R.G. Bureau. 1977. A rapid nitric perchloric acid digestion method of multielement tissue analysis. *Comm. Soil Sci. Plant Anal.* 3:425-436.

**PHOSPHORUS, POTASSIUM, SULFUR, CALCIUM, MAGNESIUM,  
SODIUM, ALUMINUM, BORON, ZINC, MANGANESE, IRON, COPPER,  
AND MOLYBDENUM OF BOTANICAL MATERIALS**

**B - 4.30**

Microwave Digestion/Dissolution Closed Vessel

### Scope and Application

The method quantitatively determines the concentration of P, S, K, Ca, Mg, Na, Al, B, Zn, Mn, Fe, Cu and Mo in botanical materials utilizing a nitric acid/hydrogen peroxide digestion in conjunction with microwave heating in closed teflon vessels. Digest analyte concentrations are determined by atomic absorption spectrometry (AAS) and inductively coupled plasma atomic emission spectrometry (ICP-AES). The digestion is based on the method described by Kingston et al. (1986) using nitric acid and modified by Sah and Miller (1992) using nitric acid and hydrogen peroxide. The digest method is incomplete relative to the total oxidation of organic carbon. Phosphorus, S, and B, analyses require an ICP-AES with a vacuum spectrometer. Potassium, Ca, Mg, Na, Al, Zn, Cu, Mn, and Fe can be analyzed by AAS or ICP-AES. Microwave nitric acid/hydrogen peroxide digests may not provide 100% recovery of Al, Si, and Se. The method has a detection limit of approximately 0.01% for P, K, Ca, and Mg and 0.2 mg kg<sup>-1</sup> for B, Zn, Cu, Fe, Mn and Mo (sample dry basis). The method can also be used for the determination of trace-elements (Co, Cd, Ni, Pb, etc.) and is generally reproducible within ± 7.0% for all analytes.

### Equipment

1. Analytical balance: 250 g capacity, resolution ± 0.1 mg.
2. Microwave digestion system and teflon double wall digestion vessels (equipped with 200 psi relief seals)(See Microwave Calibration, Comment #1).
3. Repipette dispensers, calibrated to 0.5 ± 0.05 mL and 2.0 ± 0.08 mL
4. Polypropylene centrifuge tube with cap, 15 mL graduated.
5. Atomic Absorption Spectrophotometer (AAS) and/or Inductively Coupled Plasma Atomic Emission Spectrometer (ICP-AES), vacuum system.

### Reagent

1. Deionized water, ASTM Type I grade.
2. Micro® clean detergent.
3. Nitric Acid, trace metal grade, 12 N.
4. Hydrogen peroxide 30% solution.
5. Standard Calibration solutions of P, S, K, Ca, Mg, Na, Al, B, Zn, Mn, Fe, Cu and Mo. Prepare five multielement standards: of K, Ca, Mg ranging from 5 - 500 mg L<sup>-1</sup>; P, S, and Na ranging from 1.0 - 100 mg L<sup>-1</sup>; and B, Zn, Al, Mn, Fe, Mo and Cu, ranging from 0.025 - 10.0 mg L<sup>-1</sup>. Dilute standard calibration solutions with 5 % nitric acid.

### Procedure

1. Determine the moisture content of the botanical material on a sub sample, Method 1.10.
2. Weigh 250 ± 5.0 mg of dry botanical material (See Comment #2, #3 and #4) and place in 120 mL teflon digestion vessel. Include a method blank. For samples requiring Mo and Cu analyses, sample size should be increased to 500 ± 5.0 mg (dilution factor 30:1).
3. Using repipettes add 0.50 mL of trace metal grade concentrated nitric acid and 2.00 mL of 30% hydrogen peroxide to each vessel (See Comments #5 and #6). Ensure that the sample is completely wetted by the reagents.
4. Place digestion vessel in outer body shell, cap and allow the sample and reagents to predigest for thirty (30) minutes.

5. Close vessel (See Comments #7 and #8) relief valves and place samples (twelve vessels) in the microwave. Set microwave program four (4) minutes of 296 watts power and eight (8) minutes of 565 watts power based on microwave absorbed power calibration (See Comment #1).
6. At completion remove samples and place in hood to cool (optional - place in freezer to cool for thirty (30) minutes). In a hood vent vessels by rotating release valve 1/2 revolution. Vent until vessel is completely depressurized. Remove cap, rinse cap into vessel with deionized water.
7. Quantitatively transfer the contents of the digestion vessel into the centrifuge tube, dilute to 15 mL volume, cap centrifuge tube, invert three times and store (See Comment #9, #10 and #11).
8. Elemental analysis of plant digests can be made using atomic emission spectrometry (AES) atomic absorption spectrometry (AAS), inductively coupled plasma atomic emission spectrometry (ICP-AES), or other methodologies. The method chosen will determine specific matrix modifications, calibration standards used, and the need for instrument specific sample preparations and dilutions. Determination of trace elements by ICP-AES (Co, Cd, Ni, Mo, Pb) maybe facilitated by the use of an ultrasonic nebulizer (Soltanpour, 1996). Adjust and operate instruments in accordance with manufacturer's instructions. Calibrate instrument using calibration solutions. Determine the analyte concentrations of a method blank, unknown samples and record concentrations in mg L<sup>-1</sup>.

### Calculations

Report P, K, S, Ca and Mg results to the nearest 0.001%:

$$\% \text{ analyte} = \frac{(\text{mg L}^{-1} - \text{method blank}) \times (\text{DF}) \times (0.0001)}{\text{Dry matter (\%)} / 100}$$

Report Al, Mn and Fe results to the nearest 1 mg kg<sup>-1</sup>; B, Al, Zn, Cu and Mo the nearest 0.1 mg kg<sup>-1</sup>:

$$\text{mg kg}^{-1} \text{ analyte} = \frac{(\text{mg L}^{-1} - \text{method blank}) \times (\text{DF})}{\text{Dry matter (\%)} / 100}$$

Dilution Factor (DF) : for sample weights of 250 mg is 60, for 500 mg the value is 30

### Comments

1. Microwave Calibration: Place 1.0 ± 0.1 kg of deionized water in teflon beaker and determined water temperature to the nearest ± 0.5 °C. Microwave at 40% power for two (2.0 ± 0.01) minutes, vigorously stir solution for fifteen (30) seconds and record temperature to ± 0.5 °C. Repeat using successive microwave power settings of 50%, 60%, 70% 80%, 90%, 95%, 100% with a fresh aliquot of deionized water each time. Absorbed power in watts (P) can be calculated using the change in temperature according to Equation P-4.30-1:

$$P = (\Delta T) \times (34.86) \quad [\text{equ. B-4.30-1}]$$

Plot microwave calibration function of applied energy as a percent versus absorbed power in watts. Verify and check microwave calibration every three months.

2. Teflon PFA 120 mL digestion vessel liners should be cleaned according to the following procedure: (1) soak liners in 1% solution of labware detergent for one hour; (2) rinse vessels in tap water; (3) rinse in solution of 0.5 N HCl; (4) three deionized water rinses (ASTM Type I grade); and (5) dry for one hour at 80 °C. Do not brush containers to clean.
3. Sample material must be ground to pass a 40 mesh screen (< 0.40 mm opening), to ensure homogeneity.

4. Botanical materials which may be high in starch (i.e. cereal flours) may react violently and result in rupture seal failure. When digesting these materials reduce sample mass to 200 mg of sample material. Examine digest for undecomposed sample material. Redigest sample if: (1) significant residual particulate are noted in the digest or (2) the sample shows significant discolorization (i.e. gray or black, etc.).
5. Check repipette dispensing volume, calibrate using an analytical balance.
6. When adding reagent to vessels always wear protective clothing (i.e. eye protection, lab coat, disposable gloves and shoes). Always handle reagents and opening of vessels in an acid hood capable of high air flow, 100 cfm.
7. Inspect vessel rupture seal in the cap for replacement. Samples with ruptured seals will require redigestion.
8. Follow microwave manufacturer's instructions for microwave power calibration. Applying excessive microwave power may result in rupture seal or vessel failure.
9. Centrifuging may be necessary to clear the digest.
10. Samples having analyte concentrations exceeding the highest standard will require dilution and reanalysis.
11. Place 3.0 mL of concentrate Micro® clean detergent (Baxter Scientific) in digestion vessel and allow to stand 30 minutes, rinse out any particulate, and finish cleaning according to set vessel cleaning procedure.

## Literature

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# **WATER ANALYSIS METHODOLOGIES**

## Scope and Application

This method quantifies water pH and is a measure of the relative acidity or alkalinity of the solution that is in equilibrium. It is a measure of the intensity of acidity or alkalinity, but does not indicate the relative buffering capacity of water. pH is measured to assess irrigation water chemical properties, crop and soil suitability. The method is generally reproducible within  $\pm 0.05$  pH units.

## Equipment

1. pH meter, equipped with pH electrodes (indicating and reference).
2. Primary standard buffers, pH 4.0, 7.0, and 10.0.

## Procedure

1. Standardize / Calibrate the pH meter: (1) rinse electrode with deionized water and place in pH 7.0 primary standard buffer and adjust as necessary; (2) rinse electrode and place in pH 4.0 primary standard buffer; (3) adjust the slope until response is  $\pm 0.05$  units of expected response; and (4) check pH 7.0 primary standard buffer and adjust as necessary (See Comment #1). For high pH soils ( $>7.0$ ) use pH buffers 7.0 and 10.0.
2. Insert electrode into the water. When the meter has stabilized record soil pH to the nearest 0.01 pH unit.
3. Remove electrode(s), rinse with deionized water and blot excess water with filter paper (See Comment #2).

## Comments

1. Follow manufacturer's guidelines if meter does not read within 0.05 units of primary standards. Maintenance of combination electrodes differs from that of separate reference and glass electrodes; refer to manufacturer's instructions.
2. Store pH electrodes according to manufacturer's instructions (recommended practice is to store the electrodes in a primary standard buffer).

## Literature

Rhoades, J.D. and S. Miyamoto. 1990. Testing soils for salinity. p. 299-336. *In*: R.L. Westerman (ed.) Soil testing and plant analysis. 3rd ed. SSSA, Madison, WI.

Robbins, C.W. and C.L. Wiegand. 1990. Field and laboratory measurements. p. 201-219. *In*: K.K. Tanji (ed.) ASCE manuals and Reports No. 71, Agricultural salinity, assessment, and management. American Society of Civil Engineers, 245 E. 47th St., New York.

U.S. Salinity Laboratory Staff. 1954. Saturated soil paste. Diagnosis and improvement of saline and alkali soils. Agr. Handbook 60, USDA, Washington, D.C.

### Scope and Application

This method quantifies the amount of dissolved salts (mg L<sup>-1</sup>) by measurement of the electrical conductivity (EC<sub>e</sub>) of the water solution. The relationship between EC<sub>e</sub> and soluble salts is approximate due to differences in equivalent weights, ion equivalent conductivities, and relative proportions of major solutes in the solution (Robbins, 1990). The EC<sub>e</sub> measurement is sensitive to temperature and increases approximately 1.9% per °C (range 15 - 35 °C) (Rhoades, 1996). All EC<sub>e</sub> data is normalized to 25 °C. Salt tolerance crop data is generally expressed in terms of the (EC<sub>e</sub>) and used to assess the potential of soluble salts which may limit crop productivity. The method detection limit is approximately 0.01 dS m<sup>-1</sup> (mmhos cm<sup>-1</sup>) and is generally reproducible within ± 7%.

### Equipment

1. Conductance meter with dynamic range from 0.01 to 100 dS m<sup>-1</sup> conductance, temperature compensating, 25 °C.
2. Conductance cell having a cell constant (K) appropriate to the EC of the sample being measured (see Table S -1.20 -1). Pipet-type or dip-type cell is recommended that it be capable of measuring temperature.

### Reagents

1. Deionized water CO<sub>2</sub>-free, ASTM Type I grade. EC <10<sup>-4</sup> dS m<sup>-1</sup>.
2. Standard Reference Calibration Solution. Dissolve 0.7456 g KCl (previously dried at 110 °C for 2 h) in CO<sub>2</sub>-free deionized water and dilute to 1.0 L. At 25 ±0.1 °C a 0.010 N KCl solution will have an EC<sub>e</sub> of 1.412 dS m<sup>-1</sup> (mmhos cm<sup>-1</sup>). For a 0.100 N KCl solution (7.456 g KCl diluted to 1.0 L) will have an EC<sub>e</sub> of 12.900 dS m<sup>-1</sup>. Standard EC calibration solutions are listed in Table S-1.20-1 and can be purchased from a scientific supply vendor.

### Procedure

1. Calibrate conductance cell. Operate and adjust instrument in accordance with manufacturer's instructions (See Comments #2 and #3). Rinse conductance cell with three aliquots of 0.01 N KCl, adjust a fourth portion to 25 ±0.1 °C, measure R (where R is the measured resistance ohms) and temperature *t*. Repeat measurement of R until value is constant. Calculate cell constant K.

$$K = (0.001413) R_{KCl} / [1 + 0.019(25 - t)]$$

2. Rinse conductance cell with deionized water. Draw approximately 2 mL of soil saturation paste extract solution into conductance cell, rinse and replace with a second aliquot. When the meter has stabilized, record instrument reading.

### Calculations

$$EC_{25} = C_x(1000)K[1 + 0.019(25 - t)]$$

Where: C<sub>x</sub> is the instrument measured value of the sample and *t* is temperature

Report EC<sub>e</sub> to the nearest 0.01 dS m<sup>-1</sup> as EC<sub>e</sub> 25 °C.

(See Comments #4, and #5)

Table S -1.20-1 Conductivity of KCl solutions at 25 °C (Rhoades, 1996).

Concentration N	Conductivity dS m <sup>-1</sup>
0.001	0.147
0.010	1.412
0.020	2.767
0.050	6.668
0.10	12.90
0.20	24.82

### Comments

1. Exposure of the sample to the atmosphere may cause changes in conductivity due to loss or gain of dissolved gasses: CO<sub>2</sub> and NH<sub>3</sub>-N. Freshly distilled water has a conductivity of 0.005 - 0.002 dS m<sup>-1</sup> increasing after a few weeks to 0.002 -0.004 dS m<sup>-1</sup>. This of special concern on samples with very low EC<sub>e</sub>.
2. Clean platinum electrodes that are new or that are providing erratic EC readings with acid-dichromate cleaning solution. Cleaning solution: 32 mL of saturated sodium dichromate (Na<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>) and 1L 16 M sulfuric acid. Soak electrodes 16 hours followed by three rinses of deionized water. If platinum is flaked, recoat according to procedure of APHA (1985).
3. For highly saline soils (EC<sub>e</sub> >8.0 dS m<sup>-1</sup>) calibrate using 0.100 N KCl solution, EC<sub>e</sub> 12.90 dS m<sup>-1</sup>.
4. The relationship between conductivity and soluble salts is approximate due to differences in solutes, solute conductivities, and equivalent weights. The general relationship (for solutions with an EC<sub>e</sub> range of 0.10 - 2.0 dS m<sup>-1</sup>) is:

$$\begin{aligned} \text{Dissolved salt concentration (mg L}^{-1}\text{)} &\cong 640 \times \text{EC}_e, \text{ in dS m}^{-1} \\ \text{Total cations (or anions) (mmolc L}^{-1}\text{ or meq L}^{-1}\text{)} &\cong 10 \times \text{EC}_e, \text{ in dS m}^{-1} \\ \text{Osmotic potential at 25 }^\circ\text{C (KPa)} &\cong 0.39 \times \text{EC}_e, \text{ in dS m}^{-1} \end{aligned}$$

The factor for converting EC<sub>e</sub> to total dissolved salts (mg L<sup>-1</sup>) ranges from 550 to 900 dependent on the specific anions present and their concentration. For estimating approximate total cations or anions, USDA Handbook #60, Figure 4, graphically shows this relationship for typical salt concentrations.

### Literature

APHA (1985). Part 205. In Standard Methods for the Examination of Water and Wastewater'. 16<sup>th</sup> edn. American Public Health Association, Washington, DC.

Hanson, Blaine, Stephen R. Grattan, and Allan Fulton. 1993. Agricultural salinity and drainage. University of California irrigation program, Univ. California Davis.

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Robbins, C.W. and C.L. Wiegand. 1990. Field and laboratory measurements. p. 201-219. *In: K.K. Tanji (ed.) ASCE manuals and Reports No. 71, Agricultural salinity, assessment, and management. American Society of Civil Engineers, 245 E. 47th St., New York.*

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U.S. Salinity Laboratory Staff. 1954. Saturated soil paste. Diagnosis and improvement of saline and alkali soils. *Agr. Handbook 60, USDA, Washington, D.C.*

### Scope and Application

This method quantifies bicarbonate ( $\text{HCO}_3^{1-}$ ) and carbonate ( $\text{CO}_3^{2-}$ ) concentration in  $\text{mmolc L}^{-1}$  ( $\text{meq L}^{-1}$ ) in the solution. It is based on titration with 0.10 N hydrochloric acid. The determination of  $\text{HCO}_3^{1-}$  and  $\text{CO}_3^{2-}$  should be made immediately due to the potential of the extract being super saturated relative to calcium carbonate ( $\text{CaCO}_3$ ). The concentration of  $\text{HCO}_3^{1-}$  affects the solubility of calcium, the ionic strength of the extract solution and is used to calculate the adjusted SAR (Robbins, 1990 and Hanson et al. 1993). The method detection limit is approximately  $0.05 \text{ mmolc L}^{-1}$  ( $\text{meq L}^{-1}$ ) and is generally reproducible within  $\pm 10\%$ .

### Equipment

1. Titration burette  $50.0 \pm 0.2 \text{ mL}$ , or automatic titrator.
2. pH meter and combination pH electrode.
3. Pipette,  $2.0 \pm 0.05 \text{ mL}$  and  $5.0 \pm 0.05 \text{ mL}$ .
4. 50 mL beaker.
5. Magnetic stir plate and micro size (0.25 mm) Teflon coated magnetic stir bar.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Primary standard buffer solutions: pH 4.00, 7.00 and 10.0.
3. Standardized hydrochloric acid (HCl) solution, 0.020 N with respect to  $\text{H}^+$  (See Comment #1).

### Procedure

1. Standardize / calibrate the pH meter: (1) rinse electrode with deionized water and place in pH 7.0 primary standard buffer and adjust as necessary; (2) rinse electrode and place in pH 4.0 primary standard buffer; (3) adjust the slope until response is  $\pm 0.05$  units of expected response; and (4) and recheck standard buffers (See Comments #2 and #3).
2. Place 1.0 to 50 mL aliquot of water sample in beaker, and bring to 50 mL volume with deionized water and add magnetic stirrer. Place on stir plate and insert pH electrode (See Comment #4). Record amount of titrant needed to reach a pH of 8.3 for  $\text{CO}_3^{2-}$  and 4.5 for  $\text{HCO}_3^{1-}$  to the nearest 0.2 mL.
3. Determine the amount of  $\text{HCO}_3^{1-}$  in deionized water blank solution.

### Calculations

$$\text{CO}_3^{2-} \text{ mmolc L}^{-1} = \frac{(2 \times P \times N) \times 1000}{\text{aliquot (mL)}} \quad \text{HCO}_3^{1-} \text{ mmolc L}^{-1} = \frac{(T - (2 \times P)) \times N \times 1000}{\text{aliquot (mL)}}$$

P = number of mL of HCl of normality N to reach  $\text{CO}_3^{2-}$  inflection point, pH 8.3;

T = number of mL of HCl of normality N to reach  $\text{HCO}_3^{1-}$  inflection point, pH 4.5;

aliquot = volume of sample, mL.

### Comments

1. Standardized 0.020 N HCl solution can be prepared from dilution of 1.00 N HCl standard reference solution or standardized by titration of known bases (Horneck, 1989).

2. Follow manufacturer's guidelines if meter does not read within 0.05 units of primary standards. Maintenance of combination electrodes differs from that of separate reference and glass electrodes; refer to manufacturer's instructions.
3. Store pH electrodes according to manufacturer's instructions (usual recommended practice is to store the electrodes in a primary standard buffer).

## Literature

Hanson, Blaine, Stephen R. Grattan, and Allan Fulton. 1993. Agricultural salinity and drainage. University of California Irrigation Program, Univ. California Davis.

Horneck, D.A., J.M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. Ag. Expt. Station SM 89:4. p. 13.

Rhoades, J.D. 1982. Soluble salts. p. 167-178. *In*: A. L. Page et al. (ed.) Methods of soil analysis: Part 2. Agronomy Monogr. 9. 2nd ed. ASA and SSSA, Madison, WI.

Rhoades, J.D. and S. Miyamoto. 1990. Testing soils for salinity. p. 299-336. *In*: R.L. Westerman (ed.) Soil testing and plant analysis. 3rd ed. SSSA, Madison, WI.

Robbins, C.W. and C.L. Wiegand. 1990. Field and laboratory measurements. p. 201-219. *In*: K.K. Tanji (ed.) ASCE manuals and reports No. 71, Agricultural salinity, assessment, and management. American Society of Civil Engineers, 245 E. 47th St., New York.

U.S. Salinity Laboratory Staff. 1954. Saturated soil paste. Diagnosis and improvement of saline and alkali soils. Agr. Handbook 60, USDA, Washington, D.C.

### Scope and Application

This method quantifies the concentration of chloride ( $\text{mmolc L}^{-1}$  or  $\text{meq L}^{-1}$ ) in the solution. Chloride may be determined using an ion selective electrode (potentiometric), chloridometer or ion chromatography instrument methods. Plant tolerance to chloride can be related to its concentration in irrigation water. The method detection limit is approximately  $0.1 \text{ mmolc L}^{-1}$  dependent on the method of analysis and is generally reproducible within  $\pm 10\%$ . The unit  $\text{mmolc L}^{-1}$  is the accepted scientific unit for reporting the concentration of anions and cations and is equivalent to  $\text{meq L}^{-1}$ .

### Equipment

1. Solid-state chloride electrode and double junction reference electrode, chloridometer or Cl titrator.
2. pH/ion meter or millivolt meter.

### Reagents

1. Deionized water, ASTM Type I grade.
3. Chloride standard,  $1.0 \text{ mmolc L}^{-1}$ : Dissolve 74.1 mg of KCl in 500 mL of deionized water and dilute to 1.0 L.

### Procedure

1. Determine the chloride concentration by ion selective electrode, chloridometer or ion chromatography. The instrument chosen will determine specific matrix modifications and sample dilutions. Adjust and operate instruments in accordance with manufacturer's instructions. Calibrate instrument using calibration solutions and determine chloride concentration of a method blank and unknown samples (See Comments #1, #2 and #3). Report chloride concentration in water sample to the nearest  $0.1 \text{ mmolc L}^{-1}$ .

### Comments

1. Care must be taken to clean all labware prior to analysis. Wash all labware with  $0.2 \text{ N HNO}_3$  and deionized water.
2. To accurately determine saturation paste chloride concentrations less than  $2.0 \text{ mmolc L}^{-1}$ , it is advisable to use standard additions techniques and potentiometric analysis (Fixen et al., 1988)
3. Samples containing chloride concentrations greater than the highest standard will require dilution.



## Literature

Fixen, P.W., R.H. Gelderman and J.L. Denning. 1988. Chloride tests. *In*: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region. North Dakota Agr. Expt. Sta. Bull. No. 499 (revised).

Horneck, D. A., J. M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. Ag. Expt. Station SM 89:4.

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## SOLUBLE BORON

W - 1.50

Boron, Azomethine-H Spectrophotometric, ICP-AES

### Scope and Application

This procedure quantitatively determines the boron concentration in water. It is based on the complexation of azomethine-H with  $\text{HBO}_3$  to form colored complex in an aqueous matrix with subsequent spectrophotometric measurement at 420 nm (Wolf, 1974). EDTA chelate is added to minimize chemical interferences. Boron can also be determined by Inductively coupled plasma emission spectrometry (ICP-AES) using one of three wavelengths. The method is readily adapted to manual or automated techniques. The method quantifies water soluble boron concentrations which can limit crop yield or be toxic to plant growth. The method detection limit is approximately  $0.10 \text{ mg L}^{-1}$  and is generally reproducible to within  $\pm 8\%$ .

### Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.01 \text{ g}$ .
2. 15 mL test tube or vial, polypropylene.
3. Pipette,  $2.0 \pm 0.05 \text{ mL}$  and  $3.0 \pm 0.05 \text{ mL}$ .
4. Vortex stirring device.
5. Spectrophotometer, wavelength 420 nm or ICP-AES 249.678, 249.773 or 208.959 nm.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Buffer-masking solution: Dissolve 250 g of ammonium acetate (reagent grade  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ ), 25.0 g of disodium salt of ethylenedinitrilo-teraacetic acid ( $\text{Na}_2\text{-EDTA}$ ) in 400 mL of deionized water. Very slowly add 125 mL of glacial acetic acid, while stirring using a magnetic stirrer. Temporary acidic conditions may cause a slight precipitation of the EDTA salts. Continue to stir the solution until the EDTA dissolves. Do not heat the solution. Adjust the buffer to a pH of 5.4 to 5.6 with acetic acid or  $\text{NH}_4\text{OH}$  as necessary. Prepare fresh solution every two months.
3. Azomethine-H solution: Dissolve 0.9 g of azomethine-H, 2.0 g of L-ascorbic acid in 50 mL of deionized water prewarmed to  $60^\circ\text{C}$ . Dilute to 100 mL and store in refrigerator. Solution is stable for forty-eight (48) hours (see comments #3 and #4).
4. Standard Boron Calibration solutions. Prepare six boron calibration standards: concentration range  $0.10 - 4.0 \text{ mg L}^{-1}$ , prepared in deionized water from a standard  $1000 \text{ mg L}^{-1}$  solution.

### Procedure

1. Pipette a 2.0 mL aliquot of water into a 15 mL polypropylene tube followed by 3.0 mL of the Buffer-masking solution using a pipette and stir with vortex stirring device (See Comment #1 and #2).
2. Using a repipette add 2.0 mL of azomethine-H reagent and stir contents thoroughly. Allow the mixture to stand sixty (60) minutes.
3. Prepare standard curve following steps 4-5, substituting 2.0 mL of standard calibration solution for water. A method blank is prepared in the same manner using deionized water.
4. Adjust and operate spectrophotometer instrument according to manufacture's instructions. Calibrate instrument using standard calibration solutions. Determine boron concentration of a method blank and unknown water samples (See Comments #4 and #5). For laboratories utilizing ICP-AES instrumentation calibrate using the 249.773 nm or 249.678 nm wavelength and 0.05, 0.50, 1.0 and  $4.0 \text{ mg L}^{-1}$  calibration standards for boron determination (see Appendix A-1).

## Calculations

Calculate boron concentration of water sample from working standard curve. Report boron concentration to the nearest 0.01 mg L<sup>-1</sup>.

## Comments

1. Prepare all reagents and perform all analyses in polypropylene or Teflon labware. Do not use borosilicate glassware.
2. Check pipette dispensing volume, calibrate using an analytical balance.
3. EDTA chelate is added to eliminate chemical interferences from Al, Fe and Cu. Concentration of the chelate may have to be increased for soil extracts containing high concentrations of these elements.
4. The azomethine-H reagent should be added quickly so that color development is equal for all samples. A constant check must be maintained on linearity and drift of the standard curve when analyzing a large set of samples.
5. For water samples with a distinct coloration prepare a second solution and blank for step two of the procedure adding 1.0 mL of deionized water in place of azomethine-H solution and vortex well. The blank for this determination consists of 5.0 mL of 0.02 M CaCl<sub>2</sub> solution and 1.0 mL of buffer-masking solution.

## Literature

Gaines, T.P., and G.A. Mitchell. 1979. Boron determination in plant tissues by the azomethine H method. *Commun. Soil Sci. Plant Anal.* 10:1099-1108.

Horneck, D.A., J.M. Hart, K. Topper and B. Koespell. 1989. Methods of soil analysis used in the soil testing laboratory at Oregon State University. *Ag. Expt. Station SM 89:4.*

Watson, Maurice 1988. Recommended Soil Boron Test. p. 23-25. *In: W.C. Dahnke (ed.) Recommended chemical soil test procedures. North Dakota Agricultural Experiment Station Bulletin No. 499 (revised).*

Wolf, B. 1974. Improvements in the azomethine-H method for the determination of boron. *Comm. Soil Sci. Plant Anal.* 5:39-44.

U.S. Salinity Lab. Staff. 1954. Diagnosis and Improvement of saline and alkali soils. *Agr. Handbook 60, USDA, Washington, D.C.*

### Scope and Application

This method quantitatively determines the concentration (mmolc L<sup>-1</sup>, meq L<sup>-1</sup>) of dissolved Ca, Mg and Na in water using absorption spectrometry (AAS) or Inductively coupled plasma emission spectrometry ICP-AES. A chemical interference solution is used to minimize chemical matrix effects. The Sodium Absorption Ratio (SAR) of saturation paste extract is calculated from the concentration of these cations. The relationship between cation solution concentrations and exchangeable cations in the soil, is used to estimate exchangeable sodium percentage (ESP) from the SAR (Robbins, 1990). The method detection limit for these cations is approximately 0.02 mmolc L<sup>-1</sup> on a solution basis and it is generally reproducible within ± 7%. The unit mmolc L<sup>-1</sup> is the accepted scientific unit for reporting the concentration of anions and cations and is equivalent to meq L<sup>-1</sup>.

### Equipment

1. Analytical balance: 250 g capacity, resolution ± 0.01 g.
2. Atomic Absorption Spectrophotometer (AAS) instrument.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Chemical interference solution, 5000 mg L<sup>-1</sup> lanthanum oxide (La<sub>2</sub>O<sub>3</sub>) - 2000 mg L<sup>-1</sup>, cesium chloride (CsCl) solution. Dissolve: 4.691 g La<sub>2</sub>O<sub>3</sub> and 5.071 g CsCl in 1500 mL of deionized water and add 25.0 mL of HClO<sub>4</sub> and 25.0 mL of HNO<sub>3</sub> and dilute to 2000 mL.
3. Standard calibration solutions of Ca, Mg, and Na: Prepare six calibration solutions containing 0.05 - 1.3 mmolc L<sup>-1</sup> of Na, 0.05 - 3.5 mmolc L<sup>-1</sup> of Ca, and 0.02 - 1.6 mmolc L<sup>-1</sup> for Mg prepared from 1000 mg L<sup>-1</sup> standard reference solutions and dilute to volume with chemical interference solution.

### Procedure

1. For AAS instrumentation dilute an aliquot of the water sample 10:1 with chemical interference solution (See Comment #1 and #2).
2. Adjust AAS or ICP-AES instrument according to manufacturer's instructions. Calibrate instrument using calibration solutions and determine individually cation (Ca, Mg, and Na) concentrations, record as mg L<sup>-1</sup> of analyte.

### Calculations

$$[\text{Ca}] \text{ mmolc L}^{-1} = \frac{\text{Ca mg L}^{-1} \times 10}{20.0 \text{ mg mmolc}^{-1}}$$

$$[\text{Mg}] \text{ mmolc L}^{-1} = \frac{\text{Mg mg L}^{-1} \times 10}{12.15 \text{ mg mmolc}^{-1}}$$

$$[\text{Na}] \text{ mmolc L}^{-1} = \frac{\text{Na mg L}^{-1} \times 10}{23.0 \text{ mg mmolc}^{-1}}$$

$$\text{SAR} = \frac{[\text{Na}]}{([\text{Ca}] + [\text{Mg}])/2}^{1/2}$$

Report Ca, Mg, and Na concentrations to the nearest 0.1 mmolc L<sup>-1</sup> and SAR to the nearest 0.1 (See Comments #3 and #4).

$$\text{Adjusted SAR} = \frac{[\text{Na}]}{([\text{Mg}] + 0.215 \text{ Ca}_x (\text{P}_{\text{CO}_2})^{1/3})^{1/2}}$$

## Comments

1. Water samples containing greater than 750 mg L<sup>-1</sup> soluble salts ( $> 1.2 \text{ dS m}^{-1}$ , estimated from EC<sub>e</sub> Method S - 1.20) will require additional dilution.
2. Cations may also be determined by ion chromatography instrumentation.
3. For laboratories utilizing ICP-AES instrumentation calibrate use the 422. 673 nm wavelength for Ca, 285.213 nm for Mg, and 588.995 nm wavelength for sodium (see Appendix A-1) using the standards of the calibration ranges described above.
4. For samples that contain HCO<sub>3</sub> it may be necessary to calculate the Adjusted SAR. constitutes more than 25% of the anions it may be necessary to determine the adjusted SAR. See water method W - 1.60 to calculate.

## Literature

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U.S. Salinity Lab. Staff. 1954. Diagnosis and Improvement of saline and alkali soils. Agr. Handbook 60, USDA, Washington, D.C.

### Scope and Application

This method quantifies the concentration of sulfate ( $\text{SO}_4^{2-}$  mmolc L<sup>-1</sup> or meq L<sup>-1</sup>) in water. The unit mmolc L<sup>-1</sup> is the new accepted scientific unit for reporting the concentration of anions and cations and is equivalent to meq L<sup>-1</sup>. Sulfate may be determined using turbidimetric, ion chromatography, or ICP-AES instrument methods. This method outlines the turbidimetric analysis which closely follows that described in 1992 Standard Method of the Examination of Waste Water. Sulfate is determined to evaluate anion balance. It has a method detection limit is approximately 0.02 mmolc L<sup>-1</sup> and is generally reproducible within  $\pm 7\%$ .

### Equipment

1. Magnetic stirrer.
2. Repipette dispenser calibrated to  $2.0 \pm 0.05$  mL
3. Pipette 10.0 mL.
4. Magnetic stir plate and Teflon stir bar.
5. Nephelometer (preferred), Turbidimeter or Spectrophotometer 340 nm.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Turbidimetric solution. Dissolve 30.0 g of  $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ ; 5.0 g  $\text{CH}_3\text{COONa} \cdot 3\text{H}_2\text{O}$ ; 1.0 g  $\text{KNO}_3$ ; 20 mL acetic acid,  $\text{CH}_3\text{COOH}$  (99%) and 0.111 g  $\text{Na}_2\text{SO}_4$ , in 500 mL deionized water and add 5.0 g of powered gum acacia, or gelatin (See Comment #1) suspension agent. Dilute to 1000 mL final volume.
3. Barium chloride crystals. Parr turbidimetric grade,  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  crystals 20 - 30 mesh. Use high purity  $\text{BaCl}_2$ , as low purity may result in low recovery of  $\text{SO}_4^{2-}$  (See Comment #2).
4. Standard sulfate-sulfur calibration solutions. Prepare 5.0 mmolc L<sup>-1</sup>  $\text{SO}_4^{2-}$  calibration stock solution, dissolve 0.4353 g of oven dry  $\text{K}_2\text{SO}_4$  in 500 mL of deionized water and dilute to one 1000 mL. Prepare six 100 mL calibration solutions of: 0.05, 0.1, 0.2, 0.4, 0.8, and 1.0 mmolc L<sup>-1</sup>  $\text{SO}_4^{2-}$  from a 5.0 mmolc L<sup>-1</sup>  $\text{SO}_4^{2-}$  solution and bring to final volume with deionized water.

### Procedure

1. Dilute a 10.0 mL aliquot with 2.0 mL of deionized water. Repeat using sulfate standards and method blank.
2. Add 2.0 mL of turbidimetric solution using a repipette (See Comment #3 and #4). Add magnetic stir bar and beginning stirring.
3. While stirring add 0.2 g of  $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$  crystals with measuring spoon.
4. Stir for sixty ( $60 \pm 3$ ) seconds, then remove from stirrer and after five ( $5 \pm 0.5$ ) minutes read absorbance with nephelometer or spectrophotometer at 340 nm (See Comment #5 and #6). Repeat with sulfate calibration solutions and method blank. Using standard calibration solutions and determine sulfate concentration of water samples and method blank. Record as mmolc L<sup>-1</sup>  $\text{SO}_4^{2-}$  of analyte in extract solution to two significant digits.

### Calculations

Report water  $\text{SO}_4^{2-}$  concentration:

$$\text{mmolc L}^{-1} \text{SO}_4^{2-} = (\text{mmolc L}^{-1} \text{SO}_4^{2-} \text{ water} - \text{method blank}) \times (2)$$

$$(1.0 \text{ mmolc L}^{-1} \text{SO}_4^{2-} = 48.03 \text{ mg SO}_4^{2-} \text{ L}^{-1})$$

## Comments

1. A number of suspension agents have been reported in the literature which include: gum acacia, gelatin, glycerol, PVP-K30 (polyvinylpyrrolidone), and Tween 80 which have proven effective in turbidimetric analysis. Each of these will require experimentation and practice using  $\text{SO}_4\text{-S}$  spiking to fully refine the technique.
2. Use  $\text{BaCl}_2$  specifically designated for turbidimetric determination of sulfate-sulfur. Sources: J.T. Baker Cat. Parr Turbidimetric  $\text{BaCl}_2$ , JT0974-5; VWR JT0974-5; and GFS Chemicals, Reagent Grade ACS #602.
3. Care must be taken to clean all labware prior to analysis. Pre-rinse all extraction flasks, turbidimetric and spectrometer cuvette in hot water followed by 0.5 N HCl rinse with deionized water.
4. Check repipette volume, calibrate using an analytical balance.
5. Samples containing  $\text{SO}_4^{2-}$  concentrations greater than the highest standard will require dilution.
6. For laboratories utilizing ICP-AES instrumentation calibrate use the 182.669 nm wavelength and calibration standards of 0.05, 0.50, and 5.0 mmolc  $\text{L}^{-1}$   $\text{SO}_4^{2-}$  (see Appendix A-1).

## Literature

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- Dick, W.A. and M. A. Tabatabai. 1979. Ion chromatography determination of sulfate and nitrate in solis. *Soil Sci. Soc. Am. J.* 43:899-904.
- Rhoades, J.D. and M. Clark. 1978. Sampling procedures and chemical methods in use at the U.S. Salinity Laboratory for characterizing salt-affected soils and waters. U.S. Salinity Laboratory USDA, Riverside CA. pp 11-12.
- Rhoades, J.D. 1982. Soluble salts. p. 167-178. *In*: A. L. Page et al. (ed.) *Methods of soil analysis: Part 2. Agronomy Monogr.* 9. 2nd ed. ASA and SSSA, Madison, WI.
- Standard methods for the examination of waste water. 1992. pp.131 - 132. Arnold Greendery, Lenore S. Clescerl and Andrew D. Eaton (ed.) 18th ed. American Public Health Association, American Water Works Association, and Water Pollution Control Federation.

### Scope and Application

This method quantifies the concentration of nitrate (NO<sub>3</sub><sup>-</sup>) (mmolc L<sup>-1</sup> or meq L<sup>-1</sup>) in water. Nitrate may be determined using ion chromatography or cadmium reduction spectrophotometric methods. This method outlines the use of the cadmium reduction spectrophotometric method (automated) outlined by (Keeney, 1982). The method detection limit is approximately 0.04 mmolc L<sup>-1</sup> dependent on the method of analysis and is generally reproducible within ± 10%. Nitrate is determined to evaluate the content of irrigation water and animal water supply. The unit mmolc L<sup>-1</sup> is the accepted scientific unit for reporting the concentration of anions and cations and is equivalent to meq L<sup>-1</sup>.

### Equipment

1. Spectrophotometer, autoanalyzer, or flow injection analyzer (FIA) instrument.

### Reagents

1. Deionized water, ASTM Type I grade.
2. Standard calibration solutions of NO<sub>3</sub>-N. Prepare six calibration standards ranging from 0.05 to 1.5 mmolc L<sup>-1</sup> concentration, diluted in 0.05 N CaCl<sub>2</sub> solution prepared from 16.1 mmolc L<sup>-1</sup> (1000 mg L<sup>-1</sup>) NO<sub>3</sub><sup>-</sup> standard solution.

### Procedure

1. Prepare a water sample (See Comment #1).
2. Nitrate (NO<sub>3</sub><sup>-</sup>) content of the extract is determined using a spectrophotometer, automated flow analyzer (Technicon Method No. 329-74W/A) or FIA instrument. Calibrate using standard calibration solutions and operate instrument in accordance with manufacturer instructions. Determine nitrate concentration of water sample and method blank, unknown samples and record results as mg L<sup>-1</sup> of nitrate in extract solution (See Comment #2)

### Calculations

Report water:

$$\text{mmolc L}^{-1} \text{ NO}_3^- = (\text{mmolc L}^{-1} \text{ NO}_3^- \text{ water} - \text{method blank})$$

$$(1.0 \text{ mmolc L}^{-1} \text{ NO}_3^- = 62.0 \text{ mg L}^{-1} \text{ NO}_3^-)$$

$$\text{Nitrate maybe reported as NO}_3\text{-N mg L}^{-1}, \quad \text{NO}_3\text{-N mg L}^{-1} = \text{NO}_3^- \text{ mmolc L}^{-1} \times 0.238$$

### Comments

1. Care must be taken to clean all labware prior to analysis. Wash all labware with 0.1 N HCl and deionized water.
2. Samples containing nitrate concentrations greater than the highest standard will require dilution.



## Literature

Bremner, J. M. and D.R. Keeney. 1965. Determination and isotopic ratio analysis of different forms of nitrogen in soils: I. Apparatus and procedure for distillation for and determination of ammonium. *Soil Sci. Soc. Am. Proc.* 29:504-507.

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Gelderman, R.H. and P.E. Fixen. 1988. Recommended nitrate-N tests. p. 10-11. *In: W.C. Dahnke (ed.) Recommended chemical soil test procedures for the North Central Region.* North Dakota Agr. Expt. Sta. Bulletin No. 499 (revised).

Keeney, D.R. and D.W. Nelson. 1982. Nitrogen - inorganic forms. *In A.L. Page (eds.) Methods of soil analysis, part 2.* Agron. Monogr. 9, 2nd ed. ASA and SSSA, Madison, WI. p. 643-698.

Rhoades, J.D. 1982. Soluble salts. p. 167-178. *In: A. L. Page et al. (ed.) Methods of soil analysis: Part 2.* Agronomy Monogr. 9. 2nd ed. ASA and SSSA, Madison, WI.

## Scope and Application

The Kjeldahl method quantitatively determines the amount of nitrogen in water based on the wet oxidation of organic matter using sulfuric acid and digestion catalyst and conversion of nitrogen to ammonium (Issac and Johnson, 1976). Ammonium may be determined by distillation into boric acid and titration (Jones, 1989); spectrophotometric measurement (automated or manual); or diffusion-conductivity (Carlson, 1978). The method does not quantitatively recover nitrogen from heterocyclic rings (such as nicotinic acid) or from oxidized forms such as nitrate and nitrite. The Kjeldahl digest can be used for the determination of water total phosphorus. The method is used to assess nitrogen content. The method detection limit is approximately  $0.5 \text{ mg L}^{-1}$  and is generally reproducible within  $\pm 10\%$ .

## Equipment

1. Analytical balance: 100 g capacity, resolution  $\pm 0.1 \text{ mg}$ .
2. Acid fume hood and digestion heating block ( $400 \text{ }^\circ\text{C}$ ).
3. Volumetric digestion tubes, 75 mL.
4. Repipette dispenser, calibrated  $3.0 \pm 0.1 \text{ mL}$ .

## Reagents

1. Deionized water, ASTM Type I grade.
2. Digest catalyst accelerator: prepared by mixing (100:10:1) 100 g potassium sulfate ( $\text{K}_2\text{SO}_4$ ), 10 g anhydrous copper sulfate ( $\text{CuSO}_4$ ), and 1.0 g selenium (Se) metal powder. This can be purchased as a prepared material under the brand name Kjel-tab, distributed by various chemical suppliers.
3. Concentrated sulfuric acid ( $\text{H}_2\text{SO}_4$ ), reagent grade.
4. 30% hydrogen peroxide ( $\text{H}_2\text{O}_2$ ); use fresh, as this material rapidly decomposes.
5. Standard calibration solutions of  $\text{NH}_4\text{-N}$ . Prepare six calibration standards ranging from 0.2 to  $40.0 \text{ mg L}^{-1}$  concentration, diluted with 4% (v/v) sulfuric acid, prepared from  $1000 \text{ mg L}^{-1}$  ammonium nitrogen standard solution.

## Procedure

1. Place  $25.0 \pm 0.5 \text{ mL}$  of water sample (See Comment #1) into a 75 mL volumetric digestion tube (50 mL or 100 mL digestion tubes may be substituted). Include a method blank.
2. Add Kjel-tab and 6.0 mL of concentrated sulfuric acid (See Comments #2 and #3).
3. Mix on a vortex stirrer fifteen (15) seconds to thoroughly mix the sample with acid.
4. Place the digestion tube on a digestion block, preheated to  $90 \text{ }^\circ\text{C}$  for one hour.
5. Remove from the digestion block and carefully (*slowly*) add 2-5 mL of 30% hydrogen peroxide in 1 mL increments to each digestion tube until digests begin to clear. Because this reaction takes place very rapidly, slow additions should avoid excessive foaming.
6. Place the digestion tube back on the digestion block and maintained at  $95^\circ\text{C}$  for two (2) hours or until all water has been lost through volatilization. Proceed with heating to  $370 \text{ }^\circ\text{C}$  for two hours. At completion, a blue-green color may persist.
7. Remove samples from block and leave under fume hood for 5-10 minutes. Then add 10-20 mL of deionized water using a wash bottle to each tube to prevent hardening and crystal formation. Dilute digestion tubes to volume with deionized water, cap, and invert three times.

8. Sample digests can be analyzed for ammonium nitrogen by three standard methods: They are conventional ammonium distillation into boric acid and titration (Jones, 1989); spectrophotometric determination of ammonium (automated or manual); or diffusion-conductivity method of Carlson (1978). Determine ammonium concentration of a method blank, unknown samples and record results as  $\text{mg L}^{-1}$  of  $\text{NH}_4\text{-N}$  in the digest (See Comment #4 and #5).

### Calculations

Report total Kjeldahl nitrogen results to the nearest  $1.0 \text{ mg L}^{-1}$  :

$$\% \text{ N} = (\text{mg L}^{-1} \text{ NH}_4\text{-N in digest} - \text{method blank}) \times (0.075) \times (40)$$

### Comments

1. Use 50.0 mL of sample if nitrogen content is less than  $10 \text{ mg L}^{-1}$ .
2. Check repipette dispenser delivery volume, recalibrate using an analytical balance.
3. When adding reagent to vessels and handling digests always wear protective clothing (i.e. eye protection, lab coat, disposable gloves and shoes). Always handle reagents and opening of vessels in an acid hood capable of high air flow, 100 cfm.
4. Samples having  $\text{NH}_4\text{-N}$  concentrations exceeding the highest standard will require dilution and reanalysis.
5. Sulfuric acid digest containing selenium is classified as a hazardous waste and must be disposed of in a suitable manner.

### Literature

Carlson, R.M. 1978. Automated separation and conductometric determination of ammonia and dissolved carbon dioxide. *Anal. Chem.* 48:1528-1531.

Carlson, R.M., R.I. Cabrera, J.L. Paul, J. Quick, and R.Y. Evans. 1990. Rapid direct measurement of ammonium and nitrate in soil and plant tissue extracts. *Comm. in Soil Sci. Plant Anal.* 21:1519-1529.

Chapman, H.D. and P.F. Pratt. 1961. *Methods of analysis for soils, plants and waters*, Priced Publication 4034. Berkeley: University of California, Division of Agriculture Sciences.

Issac, R.A. and W.C. Johnson. 1976. Determination of total nitrogen in plant tissue. *J. of Assoc. of Off. Anal. Chem.* 59:98-100.

Jones, J.B. 1989. *Plant analysis techniques*. Athens, GA: Benton Laboratories, Inc. p. 16-18.

# TOTAL WATER PHOSPHORUS

## Open Vessel Digestion and Dissolution

W - 2.00

### Scope and Application

The method semi-quantitatively determines the concentration of phosphorus in water samples utilizing a nitric acid hydrogen peroxide extraction/dissolution in conjunction with microwave heating in closed teflon vessels. This method closely follows that outline in EPA method 3050A. Digest analyte concentrations spectrophotometric methods or inductively coupled plasma atomic emission spectrometry (ICP-AES). The method can also be used for the determination of trace-elements (K, Ca, Mg, Na, S, Mn, Fe, and Cu). It has a detection limit of 0.05 mg/L and is generally reproducible within  $\pm 7.0\%$ .

### Equipment

1. Analytical balance: 250 g capacity, resolution  $\pm 0.1$  mg.
2. Hot plate system.
3. Repipette dispensers, calibrated to  $0.5 \pm 0.05$  mL and  $2.0 \pm 0.08$  mL
4. Polypropylene or teflon digestion tube with cap, 50 mL graduated.
5. Atomic Absorption Spectrophotometer (AAS) and/or Inductively Coupled Plasma Atomic Emission Spectrometer (ICP-AES), vacuum or purged system.

### Reagent

1. Deionized water, ASTM Type II grade.
2. Concentrated nitric acid, trace metal grade, 12 N.
3. Concentrated hydrochloric acid.
4. Standard Calibration solutions of P ranging from 0.05 - 10.0 mg L<sup>-1</sup>. Dilute standard calibration solutions with 5 % nitric acid.

### Procedure

1. Place 10.0 mL  $\pm 0.2$  mL of water (See Comment #1, #2 and #3) in a 50 mL digestion polypropylene digestion vessel. Include a method blank.
2. Using repipettes add  $9.0 \pm 0.1$  mL of trace metal grade concentrated nitric acid and  $1.0 \pm 0.1$  mL of concentrated hydrochloric acid (See Comments #4 and #5).
3. Place digestion vessel in digestion block and heat to 90 °C for two (2) hours or until all water has volatilized. Continue heating at 120 °C for thirty (30) minutes.
4. Cool and dilute to final volume of 20 mL
5. Determine phosphorus content using Spectrophotometric analysis at 880 nm using method ( ) or by ICP-AES. Adjust and operate instruments in accordance with manufacturer's instructions. Calibrate instrument using calibration solutions. Determine the analyte concentrations of a method blank, unknown samples and record concentrations in mg L<sup>-1</sup>.

### Calculations

Report phosphorus to the nearest 2 significant digits as mg L<sup>-1</sup>:

$$\text{Phosphorus Content mg L}^{-1} = (\text{mg L}^{-1} - \text{method blank}) \times (0.5)$$

### Comments

1. Teflon PFA digestion vessel (liners) should be cleaned according to the following procedure: (1) soak liners in 1% solution of labware detergent for one hour; (2) rinse vessels in tap water; (3) rinse in solution of 0.5 N HCl; (4) three deionized water rinses (ASTM Type I grade); and (5) dry for one hour at 80 °C. Do not brush containers to clean.

2. Check repipette dispensing volume, calibrate using an analytical balance.
3. When adding reagent to vessels always wear protective clothing (i.e. eye protection, lab coat, disposable gloves and shoes). Always handle reagents and opening of vessels in an acid hood capable of high air flow, 100 cfm.
4. Samples having analyte concentrations exceeding the highest standard will require dilution and reanalysis.
5. Place 3.0 mL of concentrate Micro® clean detergent (Baxter Scientific) in digestion vessel and allow to stand 30 minutes, rinse out any particulate, and finish cleaning according to set vessel cleaning procedure.

## Literature

Kalra, Y. P., D. G. Maynard, and F. G. Radford. 1989. Microwave digestion of tree foliage for multi-element analysis. *Can. J. For. Res.* 19:981-985.

Soltanpour, P.N, G.W. Johnson, S.M. Workman, J.B. Jones and R.O. Miller. 1996. Inductively coupled plasma emission spectrometry and inductively coupled plasma-mass spectrometry, p. 91-139. *In: J. M. Bartels et al. (ed.) Methods of soil analysis: Part 3 Chemical methods. 3rd.ed. ASA and SSSA, Madison, WI. Book series no. 5.*

## APPENDIX A Emerging Methods

### ION SUPPLY RATES USING PRS™ PROBES

Laboratory method - Western Ag Innovations

#### Scope and Application

Plant Root Simulator (PRS™) probes (Western Ag Innovations, Saskatoon, SK) consist of ion-exchange resin membranes in rigid plastic housing that are used to measure ion fluxes in soil (Qian and Schoenau, 2002).

A laboratory method was developed based on Qian et al. (2008) to provide a standardized measure of soil ion supply rate for soils that have been dried and ground. The method involves three major activities: 1) regeneration of PRS™ probes to saturate resin membranes with desired counter ions, 2) burial of PRS™ probes under standard moisture and temperature conditions, and 3) analysis of ions adsorbed on PRS™ probes during exposure period. Method detection limits and reproducibility are provided in Table 1.

#### Equipment

##### Regeneration

1. Anion and cation PRS™ probes (Western Ag Innovations, Saskatoon, SK).
2. Large-volume vessel with capacity for constant mixing of contents (minimum 10 L capacity per 100 PRS probes).
3. Large Ziploc bags.

##### Placement

1. Plastic bags, approximately 0.3 L.
2. Syringe.
3. Regenerated anion and cation PRS™ probes.
4. Ziploc bags for cleaned PRS™ probes.

##### Analysis

1. Test tubes.
2. Instrument(s) for analysis of ions of interest, e.g., spectrophotometer, flow injection analyzer (FIA), inductively coupled plasma optical emissions spectrometer (ICP-OES) analyzer.

#### Reagents

##### Regeneration

1. Hydrochloric acid soaking solution (0.5 N HCl, 10 L per 100 PRS probes).
2. Sodium bicarbonate soaking solution (0.5 N NaHCO<sub>3</sub>, 10 L per 100 PRS probes).
3. Purified water (deionized or reverse osmosis, ASTM Type I grade).
4. Ethylenediaminetetraacetate acid soaking solution (0.01 M EDTA, 10 L per 100 PRS probes).

##### Placement

1. Purified water (deionized or reverse osmosis, ASTM Type I grade).

##### Analysis

1. Hydrochloric acid (0.5 N HCl, see comment #1 for specifications).
2. Standard calibration solutions, e.g.,
  - a. NO<sub>3</sub>-N: 0.2 to 40 ug ml<sup>-1</sup>
  - b. NH<sub>4</sub>-N: 0.2 to 5 ug ml<sup>-1</sup>
  - c. P: 0.02 to 5 ug ml<sup>-1</sup>

## Procedure

### Regeneration

1. Submerge clean PRS™ probes in HCl acid soaking solution and mix for 1 hr (see comment #2).
2. Rinse with purified water (replace soaking solution with purified water and mix for 10 min).
3. Replace purified water with 0.5 N NaHCO<sub>3</sub> and mix for 4 hr. Repeat this process three times (see comment #2).
4. Rinse with purified water.
5. Anion PRS™ probes only: after rinse, replace purified water with 0.01 M EDTA and mix for 30 min (see comment #3).
6. Rinse with purified water.
7. Analyze a minimum of 3 PRS™ probe pairs to ensure that initial ion levels are sufficiently low for objectives (see comment #1).
8. Store PRS™ probes in sealed Ziploc bags.

### Placement (saturated paste)

1. Place 100 g of soil in plastic bag. Add water gradually and uniformly from bottom to top with a syringe until saturated (see comment #4).
2. After allowing the sample to equilibrate for 1 hr ( $\pm 0.2$  hr), insert one anion probe and one cation PRS™ probe into the bag. Ensure that entire membrane surface area is uniformly covered.
3. Allow samples to stand for 3 hr ( $\pm 0.1$  hr) at room temperature (20 to 25 °C). Record temperature (within 1 °C).
4. Remove probes from soil and immediately wash with purified water until no soil remains on any part of the probes (critical if total concentration of each ion is measured, e.g., ICP-OES analysis).
5. Store cleaned probes in Ziploc bags until analysis.

### Analysis

1. Add 35.0 ml of HCl to Ziploc bags containing one pair of PRS™ probes (17.5 ml of 0.5 N HCl per probe, i.e., 1.0 ml cm<sup>-2</sup>). Ensure that there are no air pockets between probes and the acid by moving the probes and releasing the air so that all of the ion exchange membrane is covered by the acid. Gently force any air out of the bag and seal. Allow the probes to remain submerged in the acid for 1 hr.
2. Include samples with no probes (acid blanks) and samples with non-exposed probes (probe blanks).
3. Pour the eluant into a labeled test tube.
4. Determine ion concentrations in eluant using appropriate methods. Calibrate using standard calibration solutions and operate instrument in accordance with manufacturer instructions.

## Calculation

Calculations should only be conducted for ions with contaminant levels below the criteria provided in comment #1. For routine measurements with equal numbers of anion and cation probes per sample and 1.0 ml eluant cm<sup>-2</sup> of membrane:

Ion supply rate (ug 10 cm<sup>-2</sup> burial period<sup>-1</sup>) = 20 x ion concentration in eluant (ug ml<sup>-1</sup>)

Table 1. Within-lab and among-lab relative standard deviations (RSD) for soil ion supply rate measurements using the PRS™ probe laboratory method (3 hr burial in saturated soil sample).

Ion	Within-lab RSD (%) <sup>†</sup>	Among-lab RSD (%) <sup>†</sup>
NO <sub>3</sub> -N	10	15
NH <sub>4</sub> -N	15	20
P	10	15
K	10	15
S	10	20
Ca	10	15
Mg	10	20
Fe	25	25
Mn	15	15
Cu	30	30
Zn	20	20

<sup>†</sup>Based on testing of three reference soils by six different labs.

#### Comments

1. Maximum acceptable contamination in 0.5 N HCl acid used for elution or probe blanks<sup>†</sup>.

Ion	Maximum in 0.5 N HCl acid (ppm)	Maximum probe blank (ug 10 cm <sup>-2</sup> )
NO <sub>3</sub> -N	0.2	5
NH <sub>4</sub> -N	0.2	5
P	0.01	0.4
K	0.2	10
S	0.2	10
Ca	0.5	100
Mg	0.2	10
Fe	0.01	1.0
Mn	0.01	0.2
Cu	0.01	0.06
Zn	0.01	0.3

<sup>†</sup>These criteria are guidelines and may require adjustment based on soils and objectives. High values for probe blanks may be due to insufficient regeneration (see comment #2) or contaminants in wash water.

2. During regeneration, PRS probes are first soaked in 0.5 M HCl to saturate exchange sites with H<sup>+</sup> and Cl<sup>-</sup>, then repeatedly soaked in NaHCO<sub>3</sub> to saturate exchange sites with Na<sup>+</sup> and HCO<sub>3</sub><sup>-</sup>. Each soaking replaces approximately 95% of the ions on the membrane with ions from the soaking solution. Minimum soaking durations are provided. Mixing does not have to be rapid, but should be continuous. Mixing can be accomplished by circulating, stirring or agitating the soaking solution. It may be necessary to increase the number of HCl and NaHCO<sub>3</sub> soakings if initial levels on resin membrane are high or soil samples have low supply rates of ions of interest.
3. Anion PRS™ probes are treated with EDTA to increase adsorption of polyvalent ions such as P, Fe, Mn, Cu, and Zn (Liang and Schoenau, 1995).



4. Nutrient supply rates can be determined at a range of soil moisture contents (e.g., field capacity), but care is required to ensure that soil samples are consistently brought to desired soil moisture content and that desired aeration conditions are maintained. Adding water till saturation and limiting the burial period is convenient for routine measurements with dry, ground soil samples. For 100 g of air-dry soil, addition of 20 to 30 ml of water is typically required, but lower amounts may be sufficient for soils that are high in sand, while higher amounts may be required for soils high in clay or organic matter. The feel and appearance of the sample can be used to determine when sufficient water has been added:
  - a. Consistency of pudding.
  - b. No free standing water on the surface.
  - c. Indentation of sample is removed by gentle shaking (may not apply to soils with >70% sand).
  - d. Soil surface glistens as it reflects light.

### **Literature**

Liang, J., and J.J. Schoenau. 1995. Development of resin membranes as a sensitive indicator of heavy metal toxicity in the soil environment. *International Journal of Environmental Analytical Chemistry* 59:265-275.

Qian, P., and J.J. Schoenau. 2002. Practical applications of ion exchange resins in agricultural and environmental soil research. *Canadian Journal of Soil Science* 82:9–21.

Qian, P., J.J. Schoenau, and N. Ziadi. 2008. Ion supply rates using ion-exchange resins, p. 135-140, *In* M. R. Carter and E. C. Gregorich (eds.) *Soil sampling and methods of analysis*, second edition. CRC Press, Boca Raton, FL.

## Appendix B

Table A-1. Suggested wavelengths and estimated detection limits for elements by ICP-AES.

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<u>Element</u>	<u>Wavelength (nm)</u>	<u>Detection Limit ug/L</u>
Al	309.271	20
	396.152	30
As	193.696	50@
	197.197	80@
B	249.773	5
	249.678	6
	208.959	10
Ba	455.403	1
	493.408	2
Ca	396.847	0.5
	422.673	10
Cd	317.933	10
	214.438	3
	226.502	3
Co	238.892	6
	237.862	10
Cr	205.552	6
Cu	324.754	5
	224.700	8
Fe	238.204	5
	259.940	6
K	404.721	43@
Mg	279.553	0.2
	285.213	2
Mn	257.610	1
Mo	202.030	8
	203.844	10
Na	588.995	30
Ni	221.647	10
P	213.618	80
	178.292	20*
Pb	220.353	40
S	180.669	90*
	181.979	90*
Se	196.026	80@
Zn	213.856	2
	202.548	4

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@ Sensitive to specific instrument and operation conditions.

\* Requires vacuum or purged spectrometer.

Soltanpour, P.N, G.W. Johnson, S.M. Workman, J.B. Jones and R.O. Miller. 1996. Inductively coupled plasma emission spectrometry and inductively coupled plasma-mass spectrometry. p. 91-139. *In*: J. M. Bartels et al. (ed.) *Methods of soil analysis: Part 3 Chemical methods*. 3rd.ed. ASA and SSSA, Madison, WI. Book series no. 5.

## Appendix C

Table D-1. US Standard Testing Sieve sizes and openings.

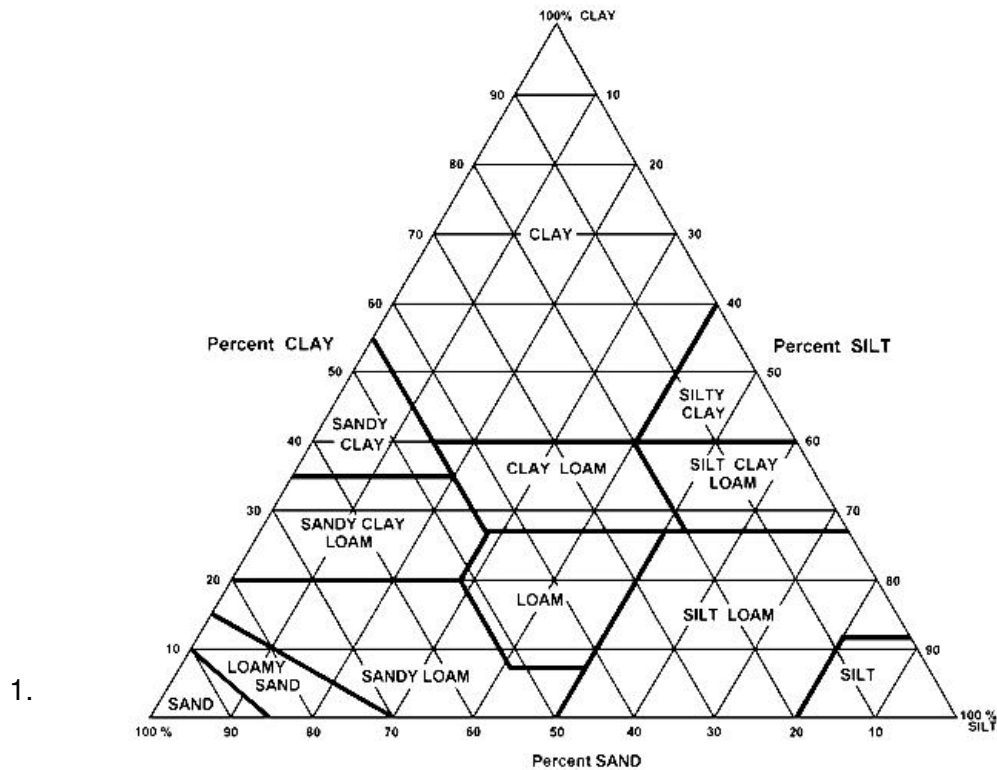
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Sieve Wire Mesh Size	Sieve Opening mm
4	4.75
5	4.00
6	3.35
7	2.80
8	2.36
10	2.00
12	1.70
14	1.40
16	1.18
18	1.00
20	0.85
25	0.710
30	0.600
35	0.500
40	0.425
45	0.355
50	0.300
60	0.250
70	0.212
80	0.180
100	0.150
120	0.125
140	0.106
200	0.075
230	0.630
270	0.530
325	0.045

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## Appendix D

Figure E-1. USDA Soil textural classification.



For more information on the relating soil texture to bulk density a calculator is available at the pedosphere web site at: [http://www.pedosphere.com/resources/bulkdensity/worktable\\_us.cfm](http://www.pedosphere.com/resources/bulkdensity/worktable_us.cfm)

Saxton, K.E., W.J. Rawls, J.S. Romberger, and R.I. Papendick. 1986. Estimating generalized soil-water characteristics from texture. *Soil Sci. Soc. Am. J.* 50(4):1031-1036.